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## Article

# Synthesis, and Molecular Structure Investigations of a New *s*-Triazine Derivatives Incorporating Pyrazole/Piperidine/Aniline Moieties

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**Abstract:** In this work, we synthesized two new *s*-triazine incorporates pyrazole/piperidine/aniline moieties. Molecular structure investigations in the light of X-ray crystallography combined with Hirshfeld and DFT calculations were presented. Intermolecular interactions controlling the molecular packing of 4-(3,5-dimethyl-1*H*-pyrazol-1-yl)-*N*-phenyl-6-(piperidin-1-yl)-1,3,5-triazin-2-amine; **5a** and *N*-(4-bromophenyl)-4-(3,5-dimethyl-1*H*-pyrazol-1-yl)-6-(piperidin-1-yl)-1,3,5-triazin-2-amine; **5b** were analyzed using Hirshfeld calculations. The most dominant interactions are the H...H, N...H and H...C contacts in both compounds. The N...H and H...C interactions in **5a** and the N...H, Br...H and H...H interactions in **5b** are the most important. In addition, DFT calculations were used to compute the molecular structures of **5a** and **5b**; then, their electronic properties, as well as the <sup>1</sup>H- and <sup>13</sup>C-NMR spectra, were predicted. Both compounds are polar where **5a** (1.018 Debye) has lower dipole moment than **5b** (4.249 Debye). The NMR chemical shifts were calculated and very good correlations between the calculated and experimental data were obtained ( $R^2 = 0.938\text{--}0.997$ ).

**Keywords:** *s*-triazine; pyrazole; hydrazino-*s*-triazine; Hirshfeld calculations



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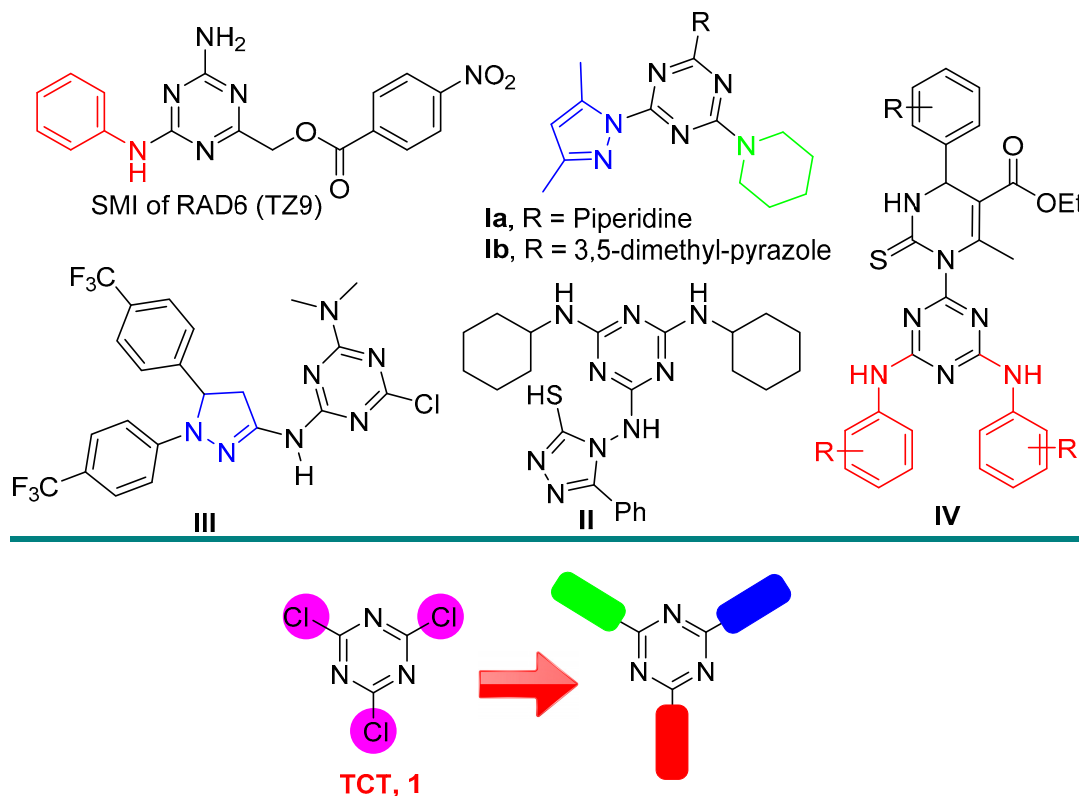
## 1. Introduction

*s*-Triazine scaffold has been extensively studied in the literature due to its wide applications in different fields, including coordination chemistry [1] and drug discovery development, such as Enasidenib (Idhifa), as a representative example, having the *s*-triazine privileged structure, which has been approved by FDA in 2017 for the IDH2-positive acute leukemia treatment [2,3]. Other pharmaceutical applications of the substituted *s*-triazine core structure were reported as fungicidal [4], antiprotozoal [5], antibacterial [6], sedative and anti-inflammatory [7], cytotoxic [8], antiviral [9], analgesic [10], and herbicide activity [11].

A combination of the *s*-triazine privileged structure with other *N*-heterocycles such as pyrazoles [12–14] and quinazoline have raised a great of attention in cancer chemotherapeutics targeting enzyme inhibitors such as TK and EGFR inhibitors [15,16].

In addition, *s*-triazine core structure has gained attention in the pharmaceutical industry but has also had a great impact in the polymer industrial applications due to ease of functionalization and mechanical and thermal stability [17]. Indeed, this fascinating structure is exploited in the metal–organic framework (MOFs) [18] as well as the storage and capture of carbon dioxide chemistry [19–21].

Several examples that have been reported in the literature have showed that replacement of the three chlorine atoms in the 2,4,6-trichlorotriazine (TCT) with different *N*-heterocycles or amines are an important task and exhibited several applications [22–24] (Figure 1). For example, El-Faham et al. have been designed and synthesized in a series of *s*-triazine hybrids comprising mon-/di-pyrazolyl along with piperidine moiety, which were subsequently assessed in vitro towards cancer activity and also in vivo toxicity in zebrafish embryos [25] (Compound I, Figure 1).



**Figure 1.** Some reported hybrids-based *s*-triazine (I–IV) and our target product from 2,4,6-trichlorotriazine (TCT, 1) with different moieties.

Matthew A. Sanders et al., have reported a tri-substituted *s*-triazine as a novel and lead compound TZ9 (Figure 1) for selective RAD6B inhibitory activity and cellular anti-cancer activity [26]. El-Hamamsy and co-workers have been designed and synthesized a new scaffold-based, *s*-triazine-targeting hDHFR for lung cancer treatment. Compound II was found to have significant activity compared to methotrexate against the A549 cancer cell line [27]. Prashanth et al., reported novel lead compounds based on *s*-triazine tethered pyrazole moiety as a potential anti-cancer agent targeting EGFR kinase inhibitors with  $IC_{50} = 229.4$  nm (Compound III, Figure 1) [28]. Additionally, the Singh research team have reported a new series of hybrids derived from monastrol-1,3,5-triazine and were studied in vitro and in vivo for their anti-cancer activity. The results indicated the potential efficacy of compound IV against cancer via inhibition of EGFR-TK (Figure 1) [29].

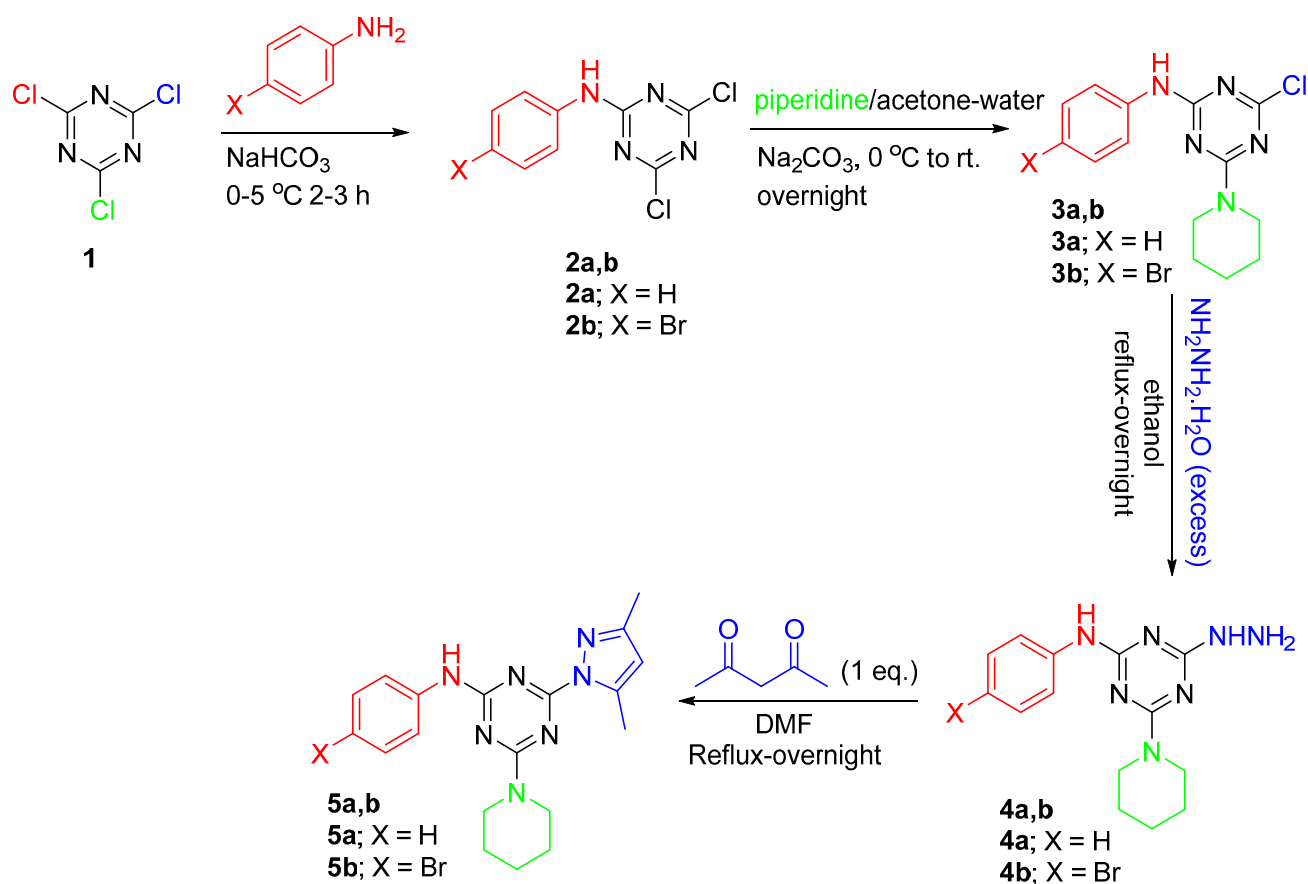
Therefore, the synthesis of new *s*-triazine derivatives connected to different amines such as aniline, piperidine, or heterocycles, such as pyrazole rings, could be useful for the drug discovery or as a new material for different applications.

In continuation of this research program towards the synthesis of *s*-triazine core structure with different moieties for biological activity applications. We reported here, the synthesis and characterization of two new compounds based on an *s*-triazine privileged structure having three different moieties such pyrazole, piperidine, and an aniline derivative. The chemical insights of the new compounds were explored via combined experimental and computational investigations.

## 2. Materials and Methods

Chemicals were purchased from Sigma-Aldrich Company (Chemie GmbH, Taufkirchen, Germany).  $^1\text{H}$ - and  $^{13}\text{C}$ -NMR spectra were recorded in  $\text{CDCl}_3$  and  $\text{DMSO-}d_6$  on a Jeol Spectrometer (Jeol, Tokyo, Japan) (400 MHz and 500 MHz). Infrared spectra were recorded on a Thermo Scientific Nicolet iS10 FT-IR spectrometer (Thermo Fisher Scientific, Waltham, MA, USA). X-ray diffraction data were collected on a Rigaku Oxford Diffraction Supernova diffractometer and processed with CrysAlisPro software v. 1.171.41.93a (Rigaku Oxford Diffraction, Yarnton, UK, 2020) using  $\text{Cu K}\alpha$  radiation.

Compounds **2a,b** and **3a,b**, which were required for this study, were prepared according to reported literature [25]. The synthetic pathway for the desired compounds are depicted in Scheme 1.

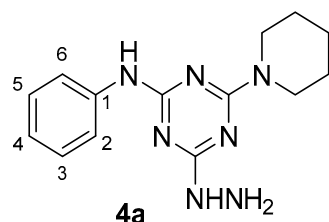


**Scheme 1.** Synthesis of a new hits tethered pyrazole/piperidine/aniline-based *s*-triazine privileged structure.

### General procedure (GP1) for the synthesis of compounds **4a,b**

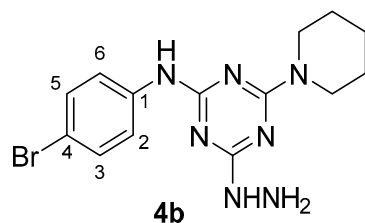
For the synthesis of monohydrazino *s*-triazine derivatives, **4a,b** excess hydrazine hydrate-80% (5.0 equiv.) was added dropwise to a solution of disubstituted *s*-triazine derivative **3a,b** (1.0 equiv.) in refluxing ethanol overnight, the evolution of the reaction was monitored by TLC (ethyl acetate/hexane 3:7). After completion of the reaction, the solution was cooled, and the precipitate was filtered and washed with diethyl ether to afford monohydrazino *s*-triazine derivatives as white solids in a very good yield and purity. Spectroscopic data fully agreed with the proposed structures. All monohydrazino *s*-triazine derivatives were used directly without further purification in the next step to synthesize the corresponding pyrazolo *s*-triazene derivatives.

#### 4-Hydrazinyl-N-phenyl-6-(piperidin-1-yl)-1,3,5-triazin-2-amine 4a



The reaction between 4-chloro-*N*-phenyl-6-(piperidin-1-yl)-1,3,5-triazin-2-amine **3a** (1.45 g, 5 mmol) and excess hydrazine hydrate (1.26 mL, 25.0 mmol) was carried out in refluxing ethanol (30 mL) for 12 h according to the general procedure (GP1). TLC analysis (ethyl acetate/hexane 3:7) showed a single spot which indicated completion of the reaction and formation of the product **4a** as a pure compound. Thereafter, the resulting white solid was filtered and used in the next step without further purification (1.10 g, 3.86 mmol, 77%), m.p. 136–140 °C; IR (KBr,  $\text{cm}^{-1}$ ): 3288, 3160 (2NH), 1586 (C=N), 1506–1440 (aromatic C-C);  $^1\text{H-NMR}$  (400 MHz,  $\text{DMSO-}d_6$ )  $\delta$  9.00 (s, 1H, -NHNH<sub>2</sub>), 7.95 (s, 1H, -NH-), 7.74 (d,  $J = 8.0$  Hz, 2H, H-2, H-6), 7.23 (t,  $J = 7.7$  Hz, 2H, H-3, H-5), 6.90 (t,  $J = 7.3$  Hz, 1H, H-4), 3.72 (s, 4H, 2NCH<sub>2</sub>-), 1.61 (q,  $J = 6.2$  Hz, 2H, -CH<sub>2</sub>-), 1.52–1.45 (m, 4H, 2-CH<sub>2</sub>-);  $^{13}\text{C-NMR}$  (126 MHz,  $\text{DMSO-}d_6$ )  $\delta$  168.19, 164.62, 164.40 (3C=N), 141.13 (C1), 128.85 (C3,5), 121.71 (C4), 119.91 (C2,6), 44.11 (2CH<sub>2</sub>N), 25.99 (2-CH<sub>2</sub>-), 24.92 (-CH<sub>2</sub>-); Chemical Formula: C<sub>14</sub>H<sub>19</sub>N<sub>7</sub>.

#### *N*-(4-Bromophenyl)-4-hydrazinyl-6-(piperidin-1-yl)-1,3,5-triazin-2-amine



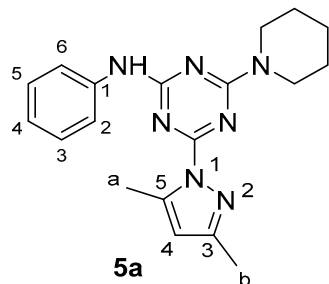
The reaction of *N*-(4-bromophenyl)-4-chloro-6-(piperidin-1-yl)-1,3,5-triazin-2-amine **3b** (1.843 g, 5.00 mmol) and excess hydrazine hydrate (1.26 mL, 25.0 mmol) in refluxing ethanol (35 mL) was carried out according to the general procedure (GP1). After 12 h, the TLC analysis (ethyl acetate/hexane 3:7) provided evidence that the desired hydrazine derivative **4b** was produced in its pure form. Ethanol and excess hydrazine were evaporated by vacuo as a part of work up procedure followed by an addition of 15 mL of *n*-hexane to the product as a white precipitate, which was collected by filtration (1.55 g, 4.75 mmol, 95%); 198–201 °C; IR (KBr,  $\text{cm}^{-1}$ ): 3224, 3230 (2NH), 1581 (C=N), 1506–1440 (aromatic C-C);  $^1\text{H-NMR}$  (400 MHz,  $\text{DMSO-}d_6$ )  $\delta$  9.16 (s, 1H, -NHNH<sub>2</sub>), 7.97 (s, 1H, -NH-), 7.74 (d,  $J = 8.5$  Hz, 2H, H-2, H-6), 7.38 (d,  $J = 8.6$  Hz, 2H, H-3, H-5), 4.19 (s, 1H, -NH), 3.71 (s, 4H, 2NCH<sub>2</sub>-), 1.60 (q,  $J = 5.8$  Hz, 2H, -CH<sub>2</sub>-), 1.51–1.44 (m, 4H, 4H, 2-CH<sub>2</sub>-);  $^{13}\text{C-NMR}$  (101 MHz,  $\text{DMSO-}d_6$ )  $\delta$  168.35, 164.64 (3C=N), 140.71 (C1), 131.57 (C4), 121.80 (C3,5), 113.12 (C2,6), 44.18 (2CH<sub>2</sub>N-), 26.04 (2-CH<sub>2</sub>-), 24.96 (-CH<sub>2</sub>-); Chemical Formula: C<sub>14</sub>H<sub>18</sub>BrN<sub>7</sub>.

#### General procedure (GP2) for the synthesis of compounds 5a,b

For the synthesis of mono-pyrazole *s*-triazine derivatives **5a,b**, a solution of acetyl acetone (1.2 equiv.) in DMF was added to a stirring solution of monohydrazino *s*-triazine derivative **4a,b** (1 equivalent) in DMF at rt. The reaction mixture was treated with triethylamine (0.8 equiv.) and was allowed to reflux while stirring overnight. It was noticed that the reactivity of bromo-hydrazino *s*-triazine derivative **5b** was relatively lower than that of the other hydrazino *s*-triazine derivatives, which required starting with excess amounts of acetylacetone (1.8 equiv.) and triethylamine (1.2 equiv.) for each equivalent of **4b**. After completion of the reaction as identified by TLC analysis (ethyl acetate/hexane 5:5), it was allowed to cool down to rt., and then a small amount of ice-cold water was added with continuous stirring for 30 min. Thereafter, the reaction mixture was placed into an ice

bath for 1 h and the precipitate of the product was separated by filtration and washed with several portions of ice-cold water and then dried overnight under vacuum to afford the product a very good yield. Some of the final products needed further purification by column chromatography on silica gel (gradient 20–60% EtOAc in hexane).

#### 4-(3,5-Dimethyl-1H-pyrazol-1-yl)-N-phenyl-6-(piperidin-1-yl)-1,3,5-triazin-2-amine

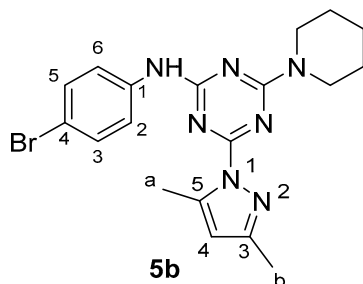


According to the general procedure (GP2), 4-hydrazinyl-*N*-phenyl-6-(piperidin-1-yl)-1,3,5-triazin-2-amine **4a** (428 mg, 1.50 mmol) was reacted with acetylacetone (180 mg, 1.80 mmol) in the presence of triethylamine (121 mg, 1.20 mmol) in refluxing DMF (10 mL). The reaction progress was monitored by TLC (ethyl acetate/hexane 5:5), which indicated the completion of the reaction after 8 h. Subsequently, the reaction mixture was stirred with ice water and placed in an ice bath for about 30 min to create the white precipitate of compound **5a**, which was filtered, washed with water, and dried under vacuum (503 mg, 1.44 mmol, 96%); m.p. 195–197 °C; IR (KBr,  $\text{cm}^{-1}$ ): 3440 (NH), 1588 (C=N), 1512–1440 (aromatic C-C);  $^1\text{H-NMR}$  (400 MHz,  $\text{CDCl}_3$ )  $\delta$  7.57 (s, 2H, H-2, H-6), 7.33 (s, 2H, H-3, H-5), 7.05 (s, 1H, H-4), 6.00 (s, 1H, H-4; Pyrazole), 3.87–3.78 (m, 4H, 2NCH<sub>2</sub>-), 2.66 (s, 3H, H-b), 2.32 (s, 3H, H-a), 1.71 (s, 2H, -CH<sub>2</sub>-), 1.66 (s, 4H, 2-CH<sub>2</sub>-);  $^{13}\text{C-NMR}$  (101 MHz,  $\text{CDCl}_3$ )  $\delta$  164.75, 164.19, 162.63 (3C=N), 151.58 (C5; Pyraz.), 143.37 (C3; Pyraz.), 138.82 (C1, Ph), 128.94 (C3,5; Ph), 123.13 (C4, Ph), 119.99 (C2,6; Ph), 110.84 (C4, Pyraz.), 45.09 (2CH<sub>2</sub>N-), 25.83 (2-CH<sub>2</sub>-), 24.75 (-CH<sub>2</sub>-), 16.21, 14.10 (2CH<sub>3</sub>; Pyraz.); Chemical Formula: C<sub>19</sub>H<sub>23</sub>N<sub>7</sub>.

#### Compound 5a was recrystallized from ethanol:

About 15 mL of absolute ethanol was added to compound **5a** and the mixture was heated gently until the entire amount dissolved completely. Thereafter, the solution was allowed to cool at rt. overnight to afford the product its crystalline form.

#### *N*-(4-Bromophenyl)-4-(3,5-dimethyl-1H-pyrazol-1-yl)-6-(piperidin-1-yl)-1,3,5-triazin-2-amine



The reaction of *N*-(4-bromophenyl)-4-hydrazinyl-6-(piperidin-1-yl)-1,3,5-triazin-2-amine **4b** (728 mg, 2.00 mmol) and acetylacetone (360 mg, 3.6 mmol) in the presence of triethylamine (243 mg, 2.40 mmol) was carried out in refluxing DMF (15 mL) while stirring overnight. Once the reaction was complete based on TLC analysis (ethyl acetate/hexane 5:5), it was cooled in an ice bath while stirring with 15 mL of ice water to afford the product **5b** as a white precipitate, which was separated by filtration, washed with water, and dried under vacuum (668 mg, 1.57 mmol, 79%); m.p. 198–201 °C; IR (KBr,  $\text{cm}^{-1}$ ): 3444 (NH), 1585 (C=N), 1530–1480 (aromatic C-C);  $^1\text{H-NMR}$  (400 MHz,  $\text{CDCl}_3$ )  $\delta$  7.44–7.46 (d,  $J = 13.3$  Hz, 4H, H-2, H-3, H-5, H-6), 6.00 (s, 1H, H-4; Pyrazole), 3.81 (s, 4H, 2NCH<sub>2</sub>-), 2.65 (s, 3H, H-b),

2.31 (s, 3H, H-a), 1.71 (s, 2H, -CH<sub>2</sub>-), 1.64 (s, 4H, 2-CH<sub>2</sub>-); <sup>13</sup>C-NMR (126 MHz, CDCl<sub>3</sub>) δ 160.70–168.00 (3C=N), 151.02 (C5; Pyraz.), 142.37 (C3; Pyraz.), 136.56 (C1, Ph), 130.78 (C3,5; Ph), 120.65 (C2,6; Ph), 114.82 (C4; Ph), 110.10 (C4; Pyraz.), 44.22 (2CH<sub>2</sub>N-), 24.73 (2-CH<sub>2</sub>-), 23.58 (-CH<sub>2</sub>-), 15.06, 13.02 (2CH<sub>3</sub>; Pyraz.); Chemical Formula: C<sub>19</sub>H<sub>22</sub>BrN<sub>7</sub>.

#### Compound 5b was recrystallized from DCM/Diethyl ether:

Compound 5b was dissolved in about 10 mL of DCM with gentle heating until the dissolution was complete, then a layer of diethyl ether (50 mL) was allowed to slide through the flask. The crystals of compound 5b were collected after 48 h.

#### Crystal structure determination

The technical experimental method for the mounting data and solved the chemical structures [30–33] are amended in the Supplementary Materials (SI). The crystallographic details are summarized in Table 1.

**Table 1.** Crystal Data of compounds 5a and 5b.

	5a	5b
CCDC	2113908	2113909
empirical formula	C <sub>19</sub> H <sub>23</sub> N <sub>7</sub>	C <sub>19</sub> H <sub>22</sub> BrN <sub>7</sub>
fw	349.44	428.34
temp (K)	120(2)	120(2)
λ (Å)	1.54184	1.54184
cryst syst	Orthorhombic	Orthorhombic
space group	<i>Pbca</i>	<i>Pbcn</i>
<i>a</i> (Å)	11.49940(10)	37.4570(3)
<i>b</i> (Å)	14.91530(10)	8.62930(10)
<i>c</i> (Å)	21.10870(10)	24.8053(2)
<i>V</i> (Å <sup>3</sup> )	3620.50(4)	8017.76(13)
<i>Z</i>	8	16
ρ <sub>calc</sub> (Mg/m <sup>3</sup> )	1.282	1.419
μ (Mo Kα) (mm <sup>-1</sup> )	0.648	2.940
No. reflns.	56200	63091
Unique reflns.	3819	8430
Completeness to θ = 67.684°	%100	%100
GOOF ( <i>F</i> <sup>2</sup> )	1.028	1.038
<i>R</i> <sub>int</sub>	0.0304	0.0301
<i>R</i> <sub>1</sub> <sup>a</sup> ( <i>I</i> ≥ 2σ)	0.0347	0.0321
<i>wR</i> <sub>2</sub> <sup>b</sup> ( <i>I</i> ≥ 2σ)	0.0893	0.0822

$$^a R_1 = \frac{\sum ||F_o| - |F_c||}{\sum |F_o|}, \quad ^b wR_2 = \frac{[\sum [w(F_o^2 - F_c^2)^2] / \sum [w(F_o^2)^2]]^{1/2}}{\sum [w(F_o^2)^2]^{1/2}}$$

#### Computational Study

Computational investigation software [34–40] employed in this study are provided in the Supplementary Materials (SI).

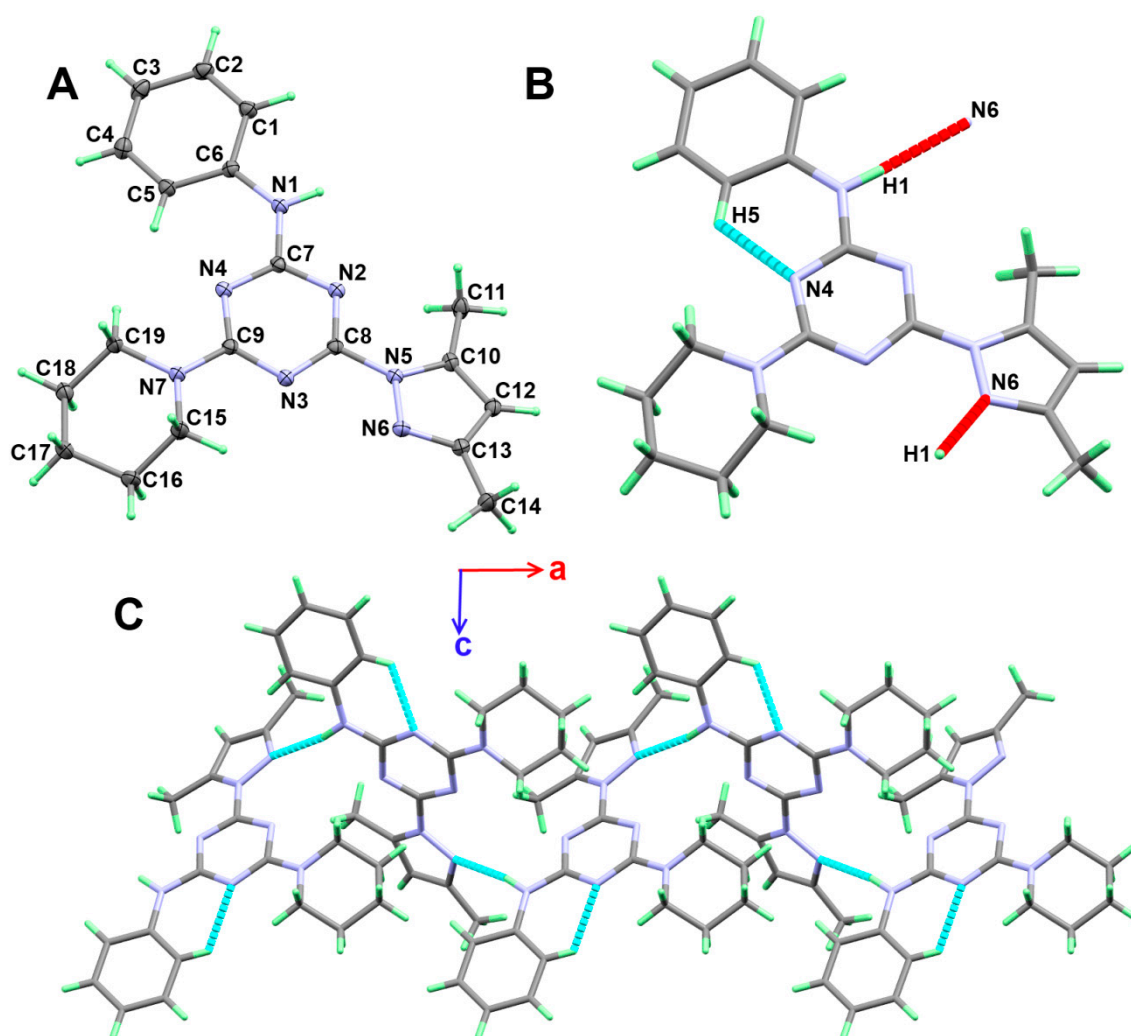
### 3. Results and Discussion

#### 3.1. Chemistry

The target *s*-triazine hybridized structures with different moieties have been synthesized, as shown in Scheme 1. Insertion of the aniline derivative was carried out initially via reaction of 2,4,6-trichlorotriazine (TCT) with the substituted aniline at 0–5 °C for 2–3 h. Subsequently, the second reactive chlorine atom was replaced by the piperidine moiety in acetone-water as a solvent at 0 °C to rt. overnight. Next, the least reactive chlorine atom was replaced with hydrazine. Finally, the hydrazine derivative was cyclized with acetyl acetone to afford the last pyrazole moiety in DMF under reflux condition. The chemical features of the synthesized *s*-triazine having different three moieties were assigned and confirmed via NMR, IR, and a single crystal X-ray diffraction analysis technique.

### 3.2. Crystal Structure Description

The X-ray structure of **5a** showing atom numbering and thermal ellipsoids drawn at 30% probability level is shown in Figure 2A. The compound crystallized in the orthorhombic system and *Pbca* space group with unit cell parameters of  $a = 11.49940(10)$  Å,  $b = 14.91530(10)$  Å, and  $c = 21.10870(10)$  Å (Table 1). The unit cell volume is  $3620.50(4)$  Å<sup>3</sup> with  $Z = 8$ , and one molecule as asymmetric unit. The structure comprised three rings connected with the *s*-triazine core, which are the aminophenyl, pyrazolyl, and morpholine moieties. The morpholine moiety occurred in the chair form and the mean plane passing through the *s*-triazine core makes angles of 17.27 and 14.71° with the phenyl and pyrazolyl moieties, respectively. Selected bond lengths and angles are listed in Table 2.



**Figure 2.** (A) Atom numbering and thermal ellipsoids at 30% probability level (B) hydrogen bond contacts (C) and molecular packing along ac plane for **5a**.

The molecular structure of **5a** is stabilized by C-H...N intramolecular hydrogen bond between the one C-H bond from the phenyl moiety as hydrogen bond donor and the neighbouring nitrogen atom from the *s*-triazine core. The donor–acceptor distance of the C-H...N intramolecular hydrogen bond is 2.9366(13) Å. The intramolecular C-H...N and the intermolecular N-H...N hydrogen bond contacts are shown in Figure 2B as turquoise and red dotted lines, respectively. Viewing of the supramolecular structure of **5a** is shown in Figure 2C, while the hydrogen bond parameters are depicted in Table 3.



**Table 2.** Selected bond lengths [Å] and angles [°] for **5a** and **5b**.

Atoms	Distance	Atoms	Distance
<b>5a</b>		<b>5b</b>	
N1-C7	1.3611(12)	Br1-C3	1.9072(17)
N1-C6	1.4072(13)	N1-C7	1.355(2)
N2-C8	1.3230(13)	N1-C6	1.411(2)
N2-C7	1.3508(13)	N2-C7	1.328(2)
N3-C8	1.3242(13)	N2-C8	1.348(2)
N3-C9	1.3552(13)	N3-C8	1.341(2)
N4-C7	1.3310(12)	N3-C13	1.464(2)
N4-C9	1.3551(12)	N3-C9	1.464(2)
N5-N6	1.3770(11)	N4-C14	1.324(2)
N5-C10	1.3841(12)	N4-C8	1.360(2)
N5-C8	1.4082(12)	N5-C14	1.328(2)
N6-C13	1.3218(13)	N5-C7	1.360(2)
N7-C9	1.3483(13)	N6-N7	1.3772(19)
N7-C19	1.4653(13)	N6-C15	1.388(2)
N7-C15	1.4660(13)	N6-C14	1.404(2)
Atoms	Angle	Atoms	Angle
C7-N1-C6	130.29(8)	C7-N1-C6	127.70(14)
C8-N2-C7	113.12(8)	C7-N2-C8	114.69(14)
C8-N3-C9	112.94(8)	C8-N3-C13	122.58(14)
C7-N4-C9	114.25(8)	C8-N3-C9	122.28(14)
N6-N5-C10	111.63(8)	C13-N3-C9	114.33(13)
N6-N5-C8	119.98(8)	C14-N4-C8	113.38(14)
C10-N5-C8	128.25(8)	C14-N5-C7	112.51(13)
C13-N6-N5	104.77(8)	N7-N6-C15	111.75(13)
C9-N7-C19	122.30(8)	N7-N6-C14	118.27(13)
C9-N7-C15	121.50(8)	C15-N6-C14	129.96(14)
C19-N7-C15	115.29(8)	C18-N7-N6	104.93(13)

**Table 3.** Hydrogen bonds for **5a** and **5b** [Å and °].

D-H ... A	d(D-H)	d(H ... A)	d(D ... A)	<(DHA)
<b>5a</b>				
N1-H1 ... N6#1	0.894(15)	2.219(15)	3.0995(12)	167.9(13)
C5-H5 ... N4	0.95	2.33	2.9366(13)	121.3
#1 $x - 1/2, y, -z + 3/2$				
<b>5b</b>				
C5-H5 ... N2	0.95	2.39	2.887(2)	111.9
C5b-H5b ... N2B	0.95	2.25	2.870(2)	121.9
C9B-H9BA ... N6B#1	0.99	2.59	3.355(2)	134.4
C9B-H9BA ... N7B#1	0.99	2.54	3.501(2)	165
C10B-H10C ... Br1B#2	0.99	2.97	3.7507(17)	136.7
N1B-H1B ... N5	0.85(2)	2.21(2)	3.0621(19)	176(2)
N1-H1 ... N4B	0.92(2)	2.25(2)	3.1452(18)	165(2)
N1-H1 ... N7B	0.92(2)	2.43(2)	3.040(2)	123.9(19)
#1 $-x + 1, -y + 1, -z + 1$ #2 $-x + 1, y, -z + 1/2$				

On the other hand, compound **5b** in the orthorhombic system and *Pbcn* space group with unit cell parameters of  $a = 37.4570(3)$  Å,  $b = 8.62930(10)$  Å, and  $c = 24.8053(2)$  Å (Table 1). The unit cell volume is  $8017.76(13)$  Å<sup>3</sup> with  $Z = 16$  and two molecules as an asymmetric unit. The structure of the two molecular units of **5b** is shown in Figure 3A,B.



contacts which appeared as red regions, indicating shorter distances than the van der Waals (vdWs) radii sum of the interacting atoms.

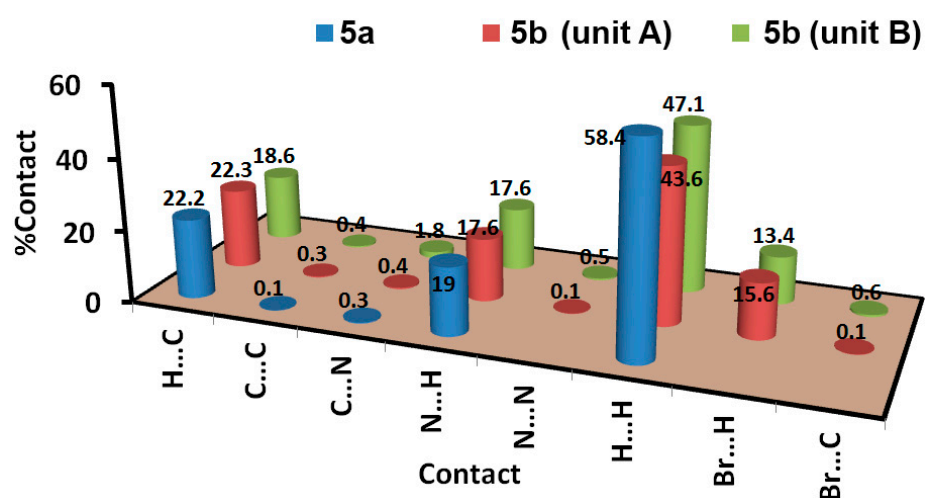


Figure 4. The percentages of all possible intermolecular interactions in the studied systems.

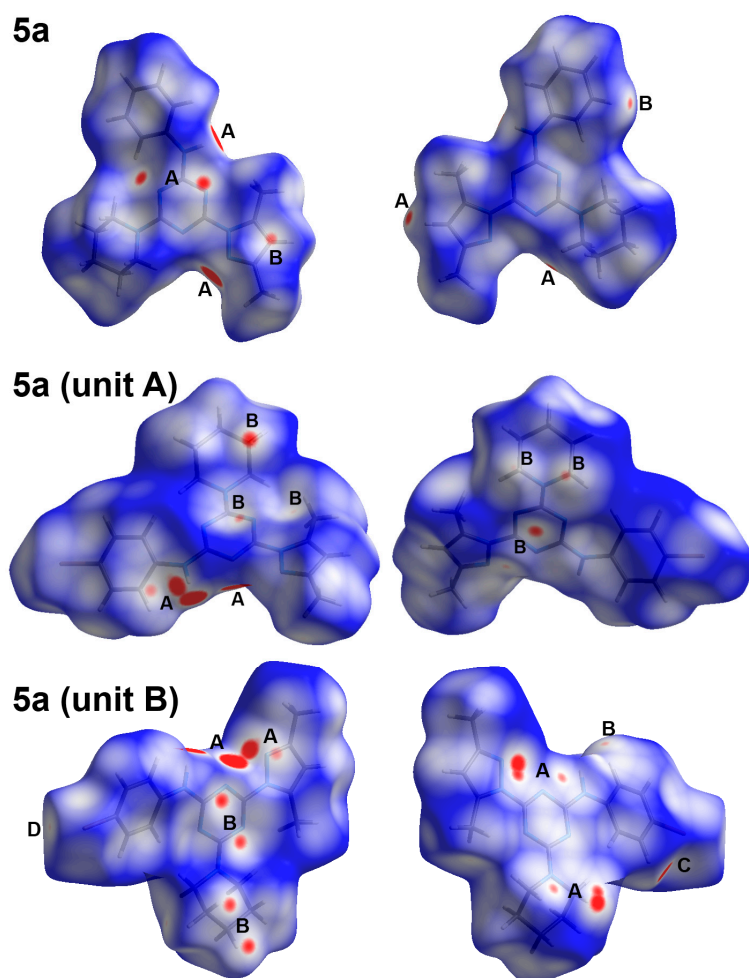


Figure 5. The  $d_{\text{norm}}$  maps of the studied systems showing the most important interactions: (A) N... H, (B) H... C, (C) H... H and (D) Br... H.

Another piece of evidence which confirms the importance of these contacts is the sharp spikes that appeared in the decomposed fingerprint plots of the N... H and H... C

contacts. The sharp spikes indicate significantly short interactions (Figure 6). Summary of the short N ... H and H ... C interactions having distances smaller than the vdWs radii sum of the interacting atoms are listed in Table 4.

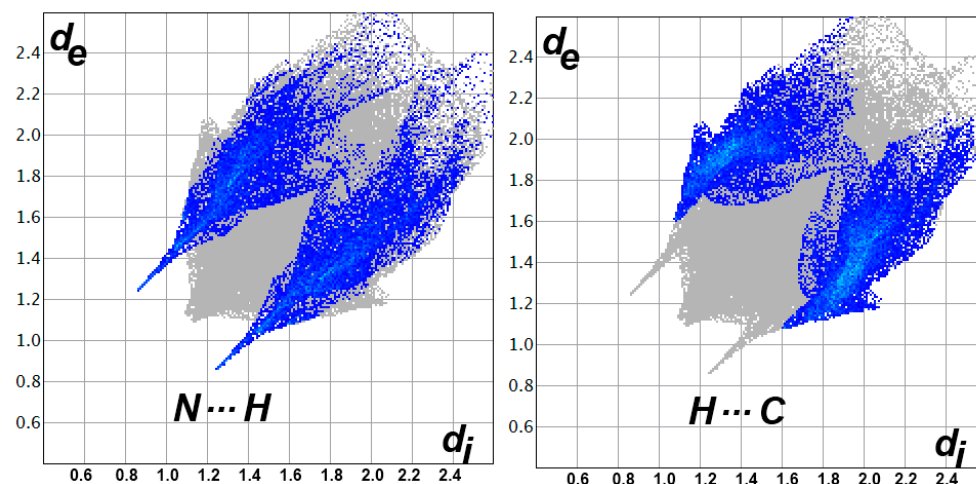


Figure 6. Decomposed fingerprint plots for the short intermolecular interactions in **5a**.

Table 4. All short intermolecular interactions in the studied systems.

Contact	Distance	Contact	Distance	Contact	Distance
<b>5a</b>			<b>5b</b>		
N6 ... H1	2.108	H13A ... C15	2.763	N7 ... H1B	2.627
N4 ... H12	2.476	H9B ... C14	2.715	N6B ... H9BA	2.523
N2 ... H15b	2.492	H1AB ... C14	2.740	N7B ... H9BA	2.446
H4 ... C12	2.687	H12B ... C8	2.632	N4B ... H13D	2.561
		N7B ... H1	2.380	H4B ... H4B	1.795
		N4B ... H1	2.162	Br1B ... H10C	2.902
		N5 ... H1B	2.055		

The structure of **5b** comprised an extra Br atom attached at the 4-position of the phenyl ring. Hence, more intermolecular interactions were detected in this compound (Figure 4). Since there are two different molecular units per asymmetric formula, there are two sets of results for this molecular system. Based on Hirshfeld calculations, and similar to **5a**, the H ... H, N ... H and H ... C are the most dominant contacts. In addition, there is significant amount of Br ... H contacts. The H12B ... C8 (2.632 Å), N5 ... H1B (2.055 Å), H4B ... H4B (1.795 Å), and Br1B ... H10C (2.902 Å) are the shortest contacts, respectively. All four of these contacts appeared as a red region in the  $d_{\text{norm}}$  map (Figure 5). An important difference between the two molecular units of **5b** is that the Br ... H and H ... H interactions are more important in the molecular unit with the letter B in the atom numbering than the other unit. Moreover, the percentages of the H ... H, Br ... H, and H ... C contacts are also different in both units (Figure 7).

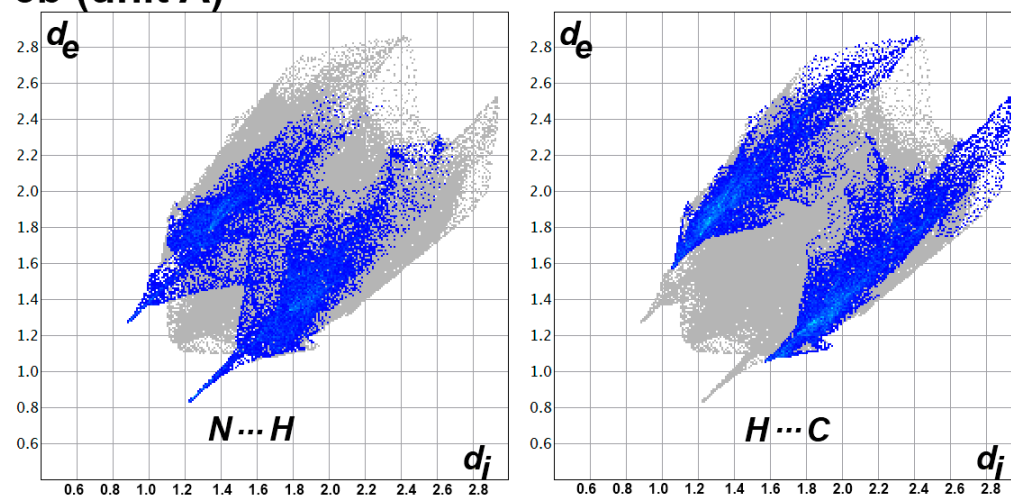
## DFT studies

### Optimized Geometry

The calculated optimized structures of **5a** and **5b** agreed very well with the reported X-ray structures, as indicated from the excellent matching between them (Figure 8) as well as the excellent correlations between the calculated and experimental bond distances and angles (Figure 9). In addition, the differences between the calculated and experimental bond distances are found to not exceed 0.01 Å (Table S1; Supplementary data). Moreover, the correlation coefficients ( $R^2$ ) are calculated to be very close to unity (Figure 9). Some higher deviations occurred in the bond angles, which could be attributed to the relaxation of the molecular structure as a consequence of the geometry optimization in the absence of the

crystal packing effect. Generally, the differences between the calculated and experimental bond angles did not exceed  $2.8^\circ$  (Table S1; Supplementary data).

### 5b (unit A)



### 5b (unit B)

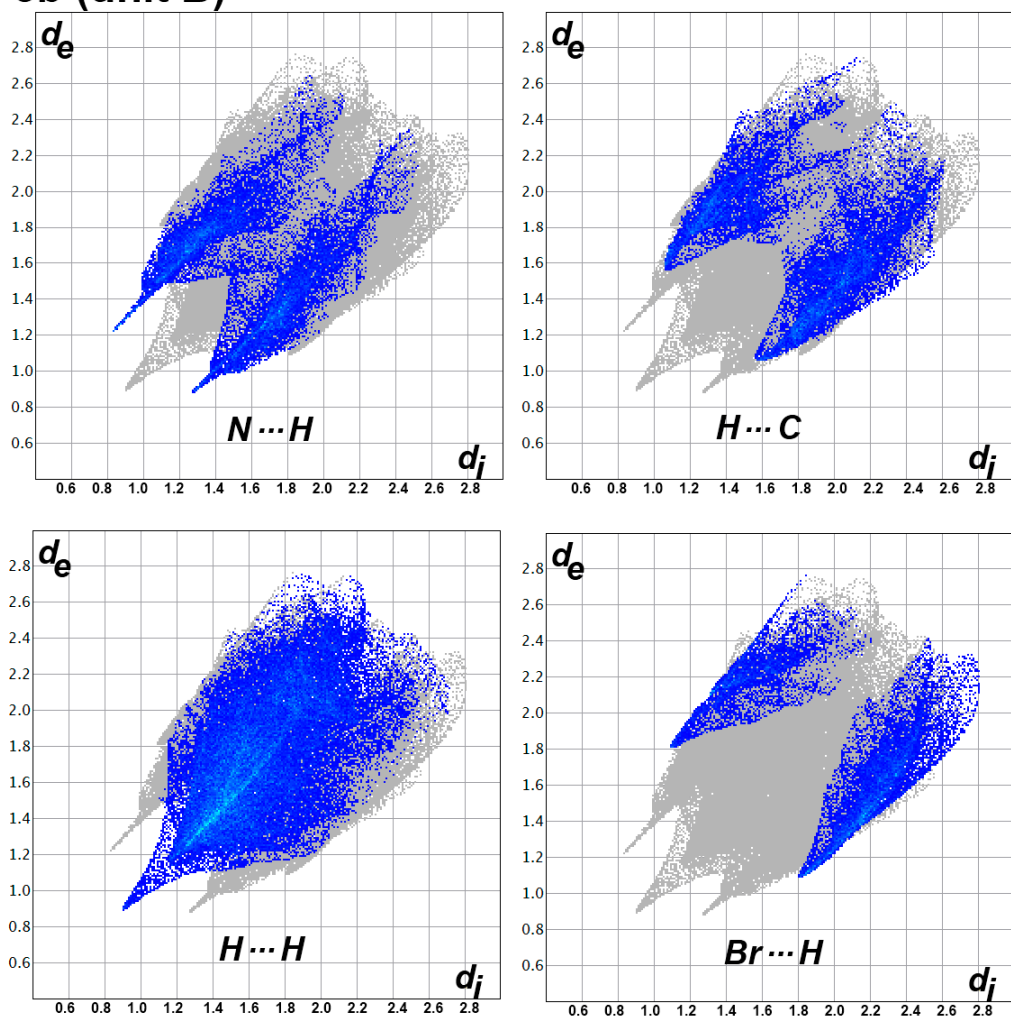
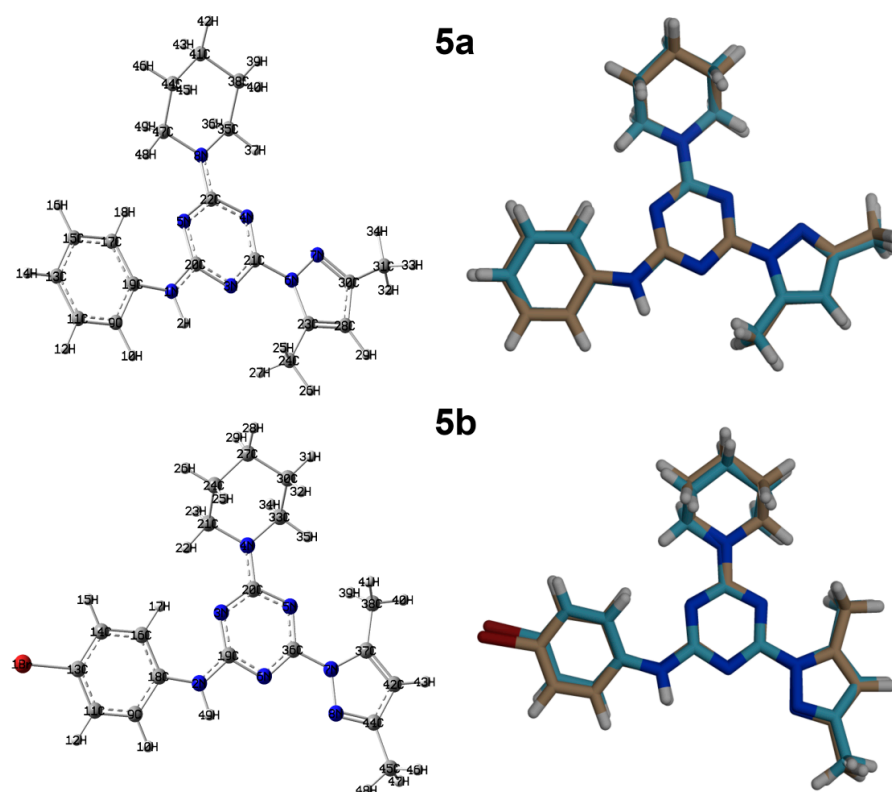
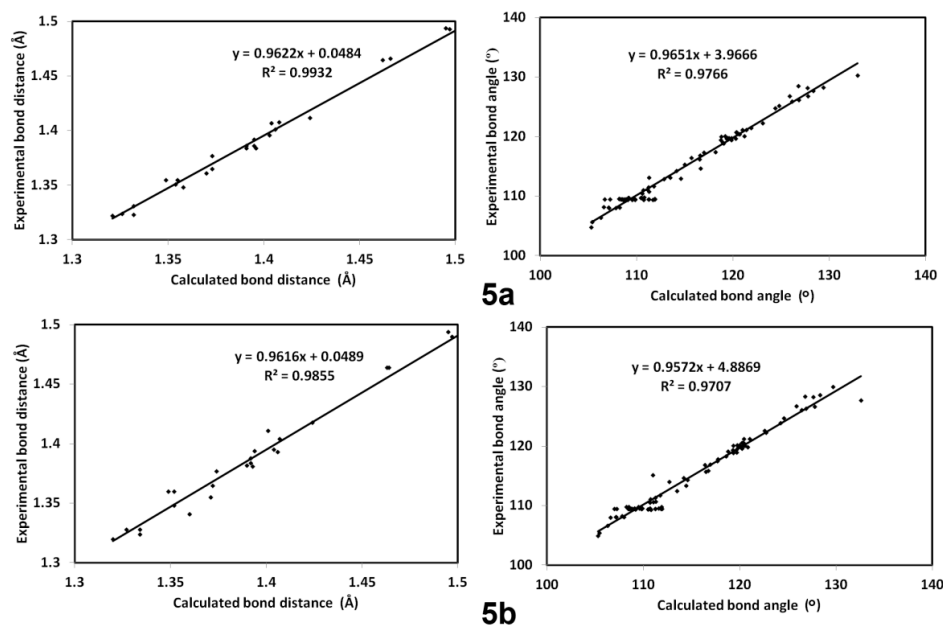


Figure 7. Decomposed fingerprint plots for the short intermolecular interactions in 5b.





**Figure 8.** Optimized structures (left) and overlay of the optimized with experimental structures (right) for the studied compounds.



**Figure 9.** The straight line correlations between the calculated and experimental geometric parameters (Table S1; Supplementary data).

Both of the studied systems comprised CHN backbone and differ only in the substituent attached to the phenyl moiety. Graphical presentations of the natural charges at the different atomic sites are shown in Figure 10. The N atoms of the triazine core and the aniline moiety are more electronegative than the N-atoms of pyrazolyl and piperidinyl moieties in both compounds (Table S2; Supplementary data). The C-atoms of the CH<sub>3</sub> groups are the most electronegative C-sites. In contrast, all C-atoms bonded to nitrogen

have high positive charge. Of C-atoms, the *s*-triazine C-atoms are the most electropositive, as these carbons found located between two nitrogen sites in the *s*-triazine moiety. Moreover, all hydrogen sites have a positive charge where the NH protons are the most positively charged. The natural charge values are calculated to be 0.434 and 0.443 e for **5a** and **5b**, respectively. In the right part of Figure 11, the molecular electrostatic potential (MEP) maps of both compounds are shown. The red regions are related to the atomic sites with high electron density, while the blue regions are related to the atomic sites with lowest electron density, which represents the most possible sites as hydrogen bond acceptor and donors, respectively.

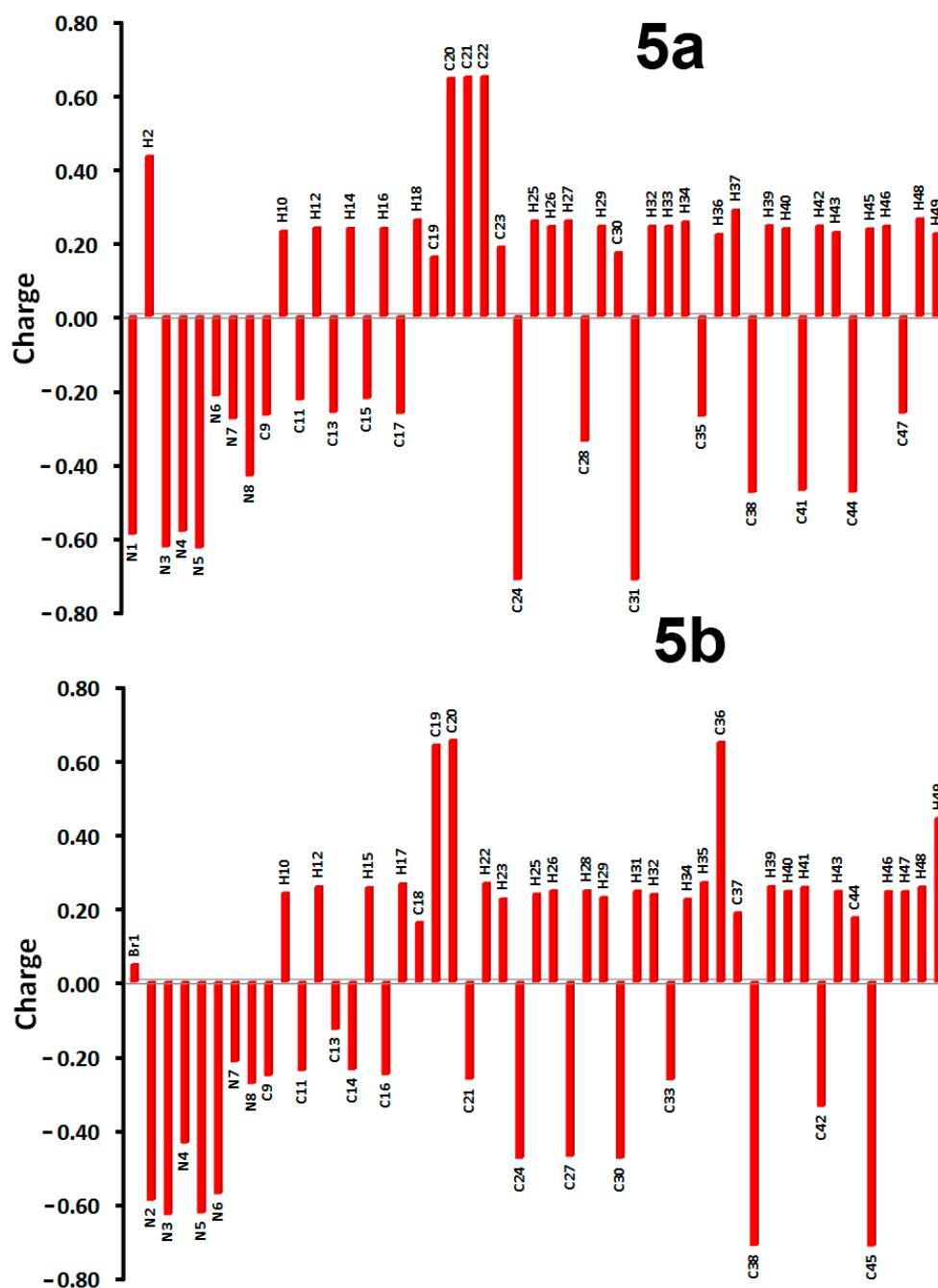
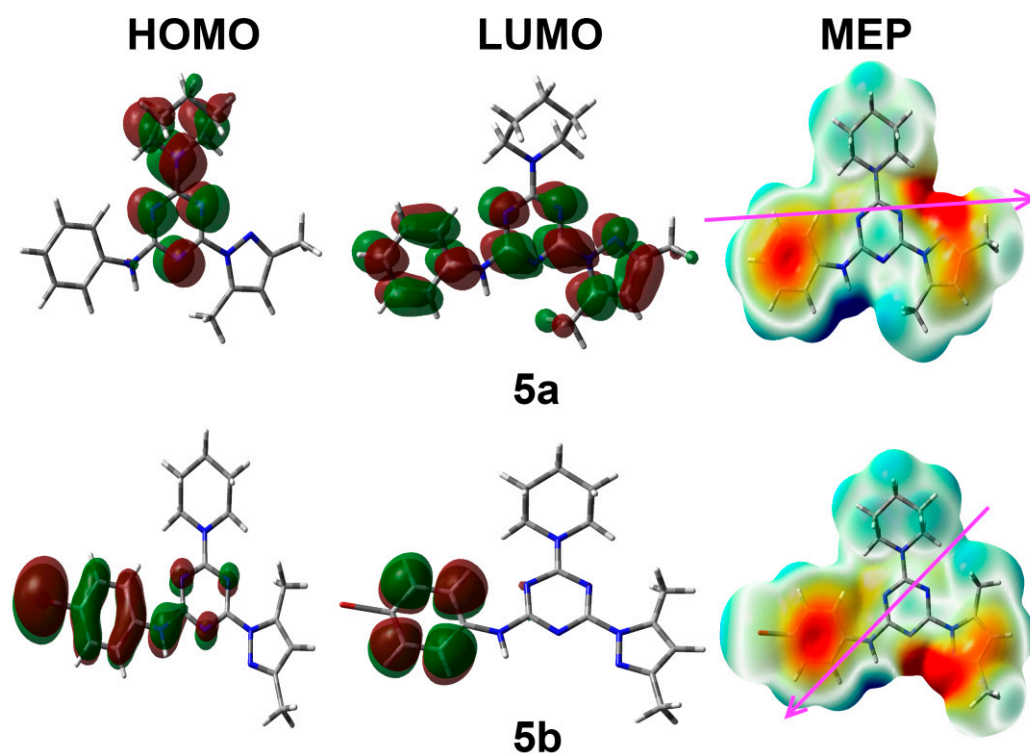


Figure 10. Natural charge populations at the different atomic sites for **5a** and **5b**.



**Figure 11.** The MEP, HOMO, and LUMO for the studied compounds. In MEP, the red and blue colors indicate the most negative and most positive regions, respectively.

Both compounds are polar molecules with net dipole moment of 1.018 and 4.249 Debye for **5a** and **5b**, respectively. The compound with bromo substituent at the phenyl ring (**5b**) has a higher polarity than the one with an unsubstituted phenyl ring (**5a**). The substituent has also a great effect on the direction of the dipole moment vector, as can be seen from Figure 11 (left part).

Moreover, the substituent at the phenyl moiety has a direct impact on the distribution of the HOMO and LUMO levels on the studied molecular systems. In case of **5a**, the HOMO level is mainly located over the *s*-triazine core and the piperidiny fragment, while the LUMO is distributed over the  $\pi$ -system of the *s*-triazine core, phenyl, and pyrazolyl moieties. In this case, the HOMO-LUMO energy gap is calculated to be 4.917 eV. On the other hand, the HOMO and LUMO are mainly located over the phenyl moiety in case of **5b**. The HOMO has some distribution over the *s*-triazine core, where the HOMO-LUMO energy gap is slightly less in **5b** (4.804 eV) compared to **5a**.

### 3.4. Electronic Reactivity Parameters

The HOMO and LUMO energies are presented in Table 5. The molecular reactivity parameters such as electrophilicity index ( $\omega = \mu^2/2\eta$ ), electron affinity ( $A = -E_{\text{LUMO}}$ ), ionization potential ( $I = -E_{\text{HOMO}}$ ), hardness ( $\eta = (I - A)/2$ ), as well as chemical potential ( $\mu = -(I + A)/2$ ) reactivity indices, were calculated based on the HOMO and LUMO energies [41–44], and the results are also listed in Table 5. The ionization potential and electron affinity values of **5b** are lower than **5a**. In contrast, the chemical hardness, chemical potential, and electrophilicity index have higher values for **5b** than **5a**.



**Table 5.** Reactivity parameters in eV of the studied compounds.

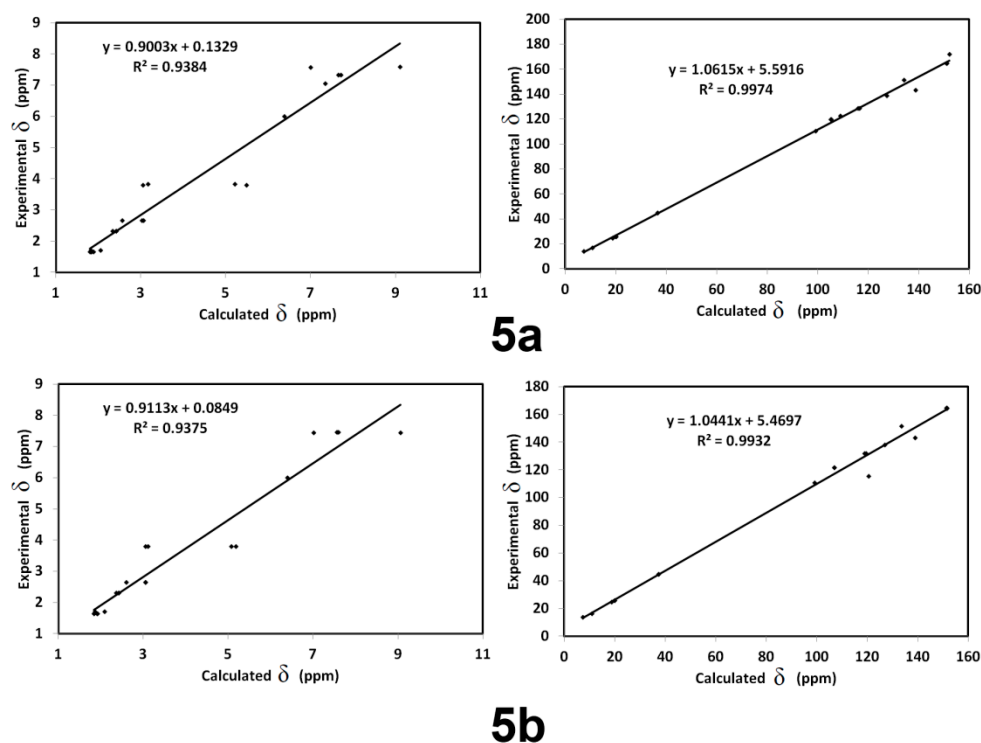
Parameter	5a	5b
$E_{\text{HOMO}}$	−5.6856	−5.7166
$E_{\text{LUMO}}$	−0.7690	−0.9129
I	5.6856	5.7166
A	0.7690	0.9129
$\eta$	4.9166	4.8037
$\mu$	−3.2273	−3.3148
$\omega$	1.0592	1.1437

### 3.5. NBO Analysis

The electron conjugation process has a great impact on the stability of molecular system [45,46]. In this regard, the different electron delocalization processes and their stabilization energies ( $E^{(2)}$ ) are summarized in Table 6. The  $\sigma \rightarrow \sigma^*$ ,  $\pi \rightarrow \pi^*$ ,  $n \rightarrow \pi^*$  and  $n \rightarrow \sigma^*$  intramolecular charge transfer (IMCT) processes stabilized the system up to 6.37, 47.40, 72.58, and 13.20 kcal/mol for the BD(1)C28-C30  $\rightarrow$  BD\*(1)C23-C24, BD(2)N5-C22  $\rightarrow$  BD\*(2)N3-C20, LP(1)N8  $\rightarrow$  BD\*(2)N5-C22 and LP(1)N5  $\rightarrow$  BD\*(1)N3-C20 in **5a**, respectively. The corresponding values for **5b** are 6.39, 44.78, 70.72, and 12.99 kcal/mol for the BD(1)C42-C44  $\rightarrow$  BD\*(1)C37-C38, BD(2)N5-C20  $\rightarrow$  BD\*(2)N6-C36, LP(1)N4  $\rightarrow$  BD\*(2)N5-C20, and LP(1)N6  $\rightarrow$  BD\*(1)N3-C19, respectively.

### 3.6. NMR Spectra

The calculated HNMR and CNMR chemical shifts by employing GIAO method and final data are listed in Supplementary Materials in Tables S3 and S4. Figure 12 presented the correlations between computed NMR chemical shifts with experimental data. As we observed from the presented graphs, the correlation coefficients ( $R^2$ ) are high and close to 1. The  $R^2$  values range from 0.93 to 0.94, indicating the close agreement between the computed and experimental results.

**Figure 12.**  $^1\text{H}$  and  $^{13}\text{C}$  NMR correlations between the calculated and experimental data.

**Table 6.** The  $E^{(2)}$  (kcal/mol) values for the charge transfer interactions in **5a** and **5b**<sup>a</sup>.

Donor NBO	Acceptor NBO	$E^{(2)}$	Donor NBO	Acceptor NBO	$E^{(2)}$
5a			5b		
$\sigma \rightarrow \sigma^*$					
BD(1)N3-C20	BD*(1)N6-C21	5.72	BD(1)N3-C19	BD*(1)N4-C20	4.75
BD(1)N4-C21	BD*(1)N8-C22	4.60	BD(1)N3-C20	BD*(1)N2-C19	5.55
BD(1)N5-C20	BD*(1)N8-C22	4.71	BD(1)N5-C20	BD*(1)N7-C36	5.24
BD(1)N5-C22	BD*(1)N1-C20	5.68	BD(1)N5-C36	BD*(1)N4-C20	4.75
BD(1)N6-N7	BD*(1)C30-C31	4.93	BD(1)N6-C19	BD*(1)N7-C36	6.05
BD(1)C15-C17	BD*(1)N1-C19	4.76	BD(1)N6-C36	BD*(1)N2-C19	4.37
BD(1)C23-C28	BD*(1)N6-C21	4.89	BD(1)N7-N8	BD*(1)C44-C45	4.93
BD(1)C23-C28	BD*(1)C30-C31	5.08	BD(1)C9-C11	BD*(1)Br1-C13	4.83
BD(1)C28-C30	BD*(1)C23-C24	6.37	BD(1)C14-C16	BD*(1)Br1-C13	4.87
			BD(1)C14-C16	BD*(1)N2-C18	4.49
			BD(1)C37-C42	BD*(1)C44-C45	5.06
			BD(1)C42-C44	BD*(1)C37-C38	6.39
$\pi \rightarrow \pi^*$					
BD(2)N3-C20	BD*(2)N4-C21	45.28	BD(2)N3-C19	BD*(2)N5-C20	38.69
BD(2)N3-C20	BD*(2)N5-C22	4.65	BD(2)N3-C19	BD*(2)N6-C36	4.66
BD(2)N4-C21	BD*(2)N3-C20	4.93	BD(2)N5-C20	BD*(2)N3-C19	4.60
BD(2)N4-C21	BD*(2)N5-C22	38.58	BD(2)N5-C20	BD*(2)N6-C36	44.78
BD(2)N4-C21	BD*(1)N6-C21	5.66	BD(2)N6-C36	BD*(2)N3-C19	38.96
BD(2)N5-C22	BD*(2)N3-C20	47.40	BD(2)N6-C36	BD*(2)N5-C20	5.07
BD(2)N5-C22	BD*(2)N4-C21	4.88	BD(2)N8-C44	BD*(2)C37-C42	11.51
BD(2)N7-C30	BD*(2)C23-C28	11.48	BD(2)C9-C11	BD*(2)C13-C14	17.56
BD(2)C9-C11	BD*(2)C13-C15	17.18	BD(2)C9-C11	BD*(2)C16-C18	19.86
BD(2)C9-C11	BD*(2)C17-C19	20.4	BD(2)C13-C14	BD*(2)C9-C11	21.19
BD(2)C13-C15	BD*(2)C9-C11	22.41	BD(2)C13-C14	BD*(2)C16-C18	16.81
BD(2)C13-C15	BD*(2)C17-C19	18.66	BD(2)C16-C18	BD*(2)C9-C11	18.55
BD(2)C17-C19	BD*(2)C9-C11	18.26	BD(2)C16-C18	BD*(2)C13-C14	23.75
BD(2)C17-C19	BD*(2)C13-C15	21.9	BD(2)C37-C42	BD*(2)N8-C44	27.09
BD(2)C23-C28	BD*(2)N7-C30	27.29			
$n \rightarrow \pi^*$					
LP(1)N1	BD*(2)N3-C20	59.13	LP(1)N4	BD*(2)N5-C20	70.72
LP(1)N1	BD*(2)C17-C19	36.38	LP(1)N7	BD*(2)N6-C36	51.52
LP(1)N6	BD*(2)N4-C21	51.37	LP(1)N7	BD*(2)N8-C44	23.23
LP(1)N6	BD*(2)N7-C30	23.44	LP(1)N7	BD*(2)C37-C42	34.67
LP(1)N6	BD*(2)C23-C28	34.89			
LP(1)N8	BD*(2)N5-C22	72.58			
$n \rightarrow \sigma^*$					
LP(1)N3	BD*(1)N4-C21	12.25	LP(1)N3	BD*(1)N5-C20	12.07
LP(1)N3	BD*(1)N5-C20	12.16	LP(1)N3	BD*(1)N6-C19	12.90
LP(1)N4	BD*(1)N3-C21	12.99	LP(1)N4	BD*(1)C21-H23	4.79
LP(1)N4	BD*(1)N5-C22	12.74	LP(1)N5	BD*(1)N3-C20	12.09
LP(1)N5	BD*(1)N3-C20	13.20	LP(1)N5	BD*(1)N6-C36	12.63
LP(1)N5	BD*(1)N4-C22	11.88	LP(1)N6	BD*(1)N3-C19	12.99
LP(1)N7	BD*(1)N6-C23	6.89	LP(1)N6	BD*(1)N5-C36	12.60
LP(1)N7	BD*(1)C28-C30	6.08	LP(1)N8	BD*(1)N7-C37	6.87
LP(1)N8	BD*(1)C44-C47	4.48			

<sup>a</sup> Atom numbering refer to Figure 5.

#### 4. Conclusions

Two new hybrids tethered pyrazole/piperidine/aniline moieties-based *s*-triazine privileged structure were synthesized in a high chemical yield using a very simple and efficient method. The supramolecular structures of **5a** and **5b** were quantitatively determined using Hirshfeld analysis. The calculated structures using DFT B3LYP/6-31G(d,p) agreed very well with the experimental X-ray geometries. Moreover, calculated NMR spectra are found in very good correlations with the experimental data. These compounds

well be further explored by our groups with other derivatives in the near future for their biological applications.

**Supplementary Materials:** The following are available online at <https://www.mdpi.com/article/10.3390/cryst11121500/s1>, Crystal structure determination; protocol and softwares utilized in this study for computational investigation; Figures S1–S12: <sup>1</sup>HNMR, <sup>13</sup>CNMR, and IR spectrum for the synthesized compounds; Table S1–S4: Computational study investigation.

**Author Contributions:** Conceptualization, A.E.-F. and A.B.; methodology, I.S.; software, S.M.S. and M.H.; formal analysis, I.S. and M.H.; investigation, I.S.; resources, A.E.-F. and A.B.; data curation, I.S., S.M.S., M.H. and A.D.; writing—original draft preparation, S.M.S., M.H. and A.B.; writing—review and editing, S.M.S., A.B. and A.E.-F.; visualization, A.B., and A.E.-F.; supervision, A.D., A.B. and A.E.-F.; project administration, A.E.-F. and A.B.; funding acquisition, A.B. All authors have read and agreed to the published version of the manuscript.

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**Conflicts of Interest:** The authors declare no conflict of interest.

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