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## Article

# Synthesis, X-ray Structure, Conformational Analysis, and DFT Studies of a Giant *s*-Triazine *bis*-Schiff Base

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**Abstract:** The current work involves the synthesis of 2,2'-(6-(piperidin-1-yl)-1,3,5-triazine-2,4-diyl)bis(hydrazin-2-yl-1-ylidene)bis(methanylylidene)diphenol **4**, characterization, and the DFT studies of the reported compound. The crystal unit cell parameters of **4** are  $a = 8.1139(2) \text{ \AA}$ ,  $b = 11.2637(2) \text{ \AA}$ ,  $c = 45.7836(8) \text{ \AA}$ . The unit cell volume is  $4184.28(15) \text{ \AA}^3$  and  $Z = 4$ . It crystallized in the orthorhombic crystal system and *Pbca* space group. The O ... H, N ... H, C ... H, H ... H and C ... C intermolecular contacts which affect the crystal stability were quantitatively analyzed using Hirshfeld calculations. Their percentages were calculated to be 9.8, 15.8, 23.7, 46.4, and 1.6% from the whole contacts occurred in the crystal, respectively. Conformational analysis was performed using DFT calculations for 17 suggested conformers and the most stable conformer was found to be the one which is stabilized by two intramolecular O-H ... N hydrogen bonding interactions. This conclusion was further revealed by natural bond orbital calculations.

**Keywords:** *s*-triazine; *bis*-Schiff base; conformational analysis; Hirshfeld analysis; NBO



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## 1. Introduction

Given the effectual reactivity of TCT (cyanuric chloride, 2,4,6-trichlorotriazine) with a diversity of nucleophiles, it is regularly used in organic synthesis as template to access numerous molecular systems [1,2]. TCT is an important moiety due to its lower price, marketable availability, and the three chlorine atoms can be replaced in stepwise manner [1]. TCT derivatives were reported to have a broad range of biological activities [3]. Some new *s*-triazine based chalcones and their derivatives were found to work as potent antimicrobial, anti-cancer and anti-malarial agents [4–9].

On the other hand, Schiff bases of triazine derivatives are considered as extraordinary class of compounds. The two connected nitrogen atoms (-CH=N-N-) are of distinct nature and the C=N that is conjugated with a lone pair of the second nitrogen are responsible for their properties [10–16]. This family of compounds are regularly used as polydentate chelating agents that form a diversity of complexes with a range of different metals [17–19]. Enormous hydrazones derivatives and their complexes have been used in several applications, such as metal-ions extraction and microdetermination of metal-ion [20]. Moreover, their biological activities have been reported in literatures [21–25]. Many synthesized *s*-triazine hydrazone derivatives were reported and explored with superior interest in

supramolecular and coordination chemistry [26,27], material chemistry [28], and complexation with several metal-ions [29–33].

It is well known that the *s*-triazine *bis*-Schiff base compounds could exist in two main structures which are shown in Figure 1. In view of the interesting importance and the diverse structural features of this class of *s*-triazine based compounds, the present work presents the synthesis, characterization, X-ray single crystal structure investigations combined with Hirshfeld and DFT calculations of a novel giant *s*-triazine based Schiff base (4). Relative stability of 17 suggested conformers of 4 were also investigated.

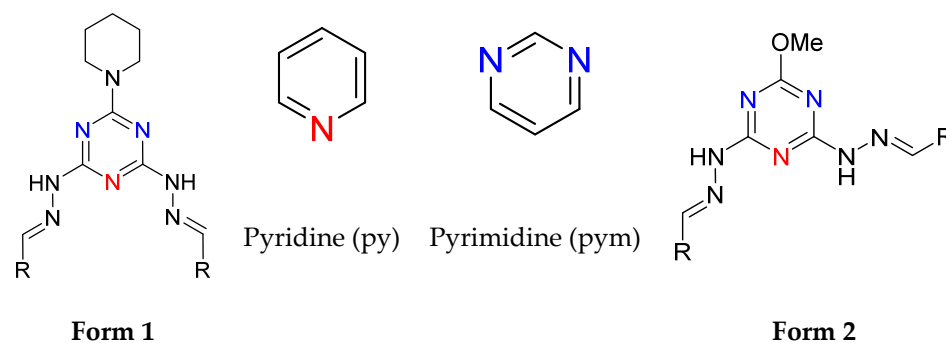


Figure 1. The general structure of *s*-triazine *bis*-Schiff base.

## 2. Materials and Methods

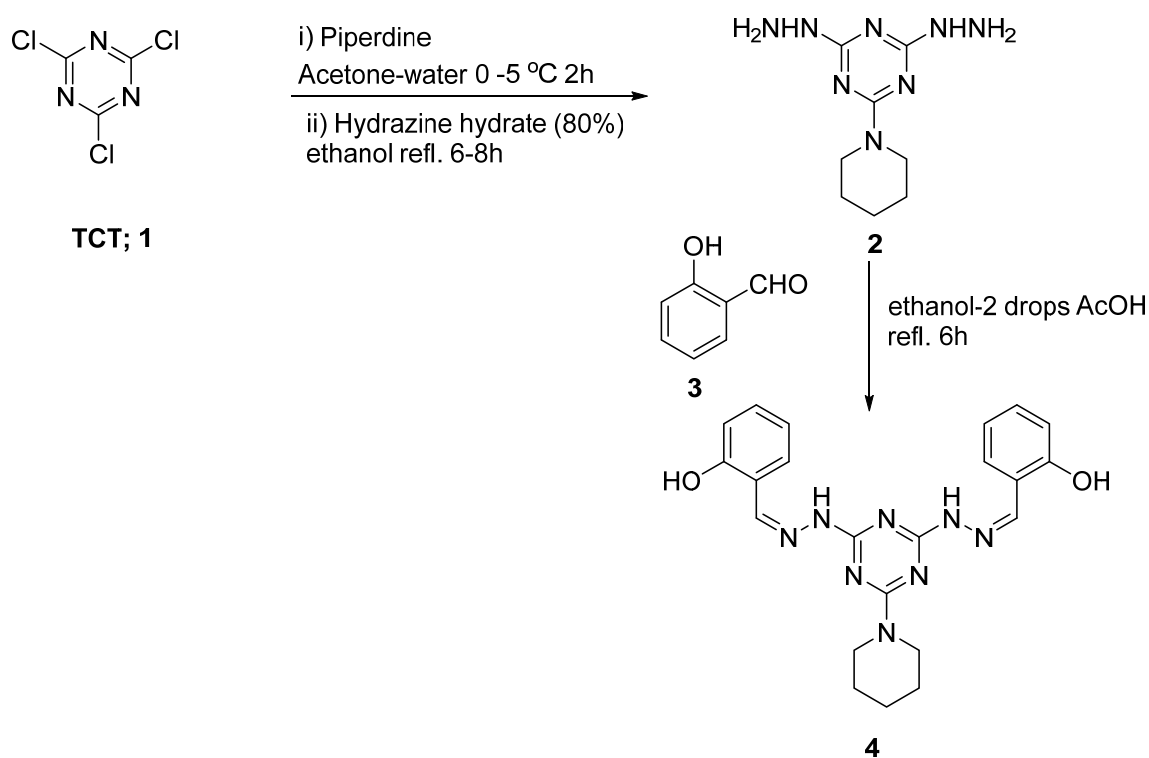
Chemicals were purchased from Sigma-Aldrich Company (Chemie GmbH, 82024 Taufkirchen, Germany). Perkin-Elmer 2400 instrument (PerkinElmer, Inc., 940 Winter Street, Waltham, MA, USA) was used for CHN analyses.

### 2.1. Synthesis of 2,2'-(6-(piperidin-1-yl)-1,3,5-triazine-2,4-diyl)bis(hydrazin-2-yl-1-ylidene))bis(methanylylidene)diphenol; 4

2,4-Dihydrazinyl-6-(piperidin-1-yl)-1,3,5-triazine (10 mmol) was added portionwise to a hot solution of 2-hydroxybenzaldehyde (10 mmol) in ethanol (100 mL) containing acetic acid (2–3 drops; Scheme 1). After complete addition, the reaction was heated under reflux for 6 h and then was left cooling at room temperature and then filter the pure pale-yellow solid (the complete reaction was followed by TLC (ethyl acetate-hexane 2:1)). The product was left to dry and then recrystallized from ethanol. Yield 95%; mp 263–265 °C;  $^1\text{H}$  NMR (DMSO- $d_6$ ):  $\delta$  = 1.52 (brs, 4H, 2CH<sub>2</sub>), 1.63 (brs, 2H, CH<sub>2</sub>), 3.80 (brs, 2H, CH<sub>2</sub>), 6.87–6.95 (m, 4H, Ar), 7.37–7.41 (m, 2H, Ar), 7.21–7.26 (m, 2H, Ar), 8.29 (s, 2H, CH=N), 11.20 (brs, 2H, NH), 12.15 (brs, 2H, OH) ppm;  $^{13}\text{C}$  NMR (DMSO- $d_6$ ):  $\delta$  = 24.3, 25.6, 43.6, 116.1, 118.9, 130.3, 143.9, 157.4, 162.8, 164.2 ppm; Anal. Calc. for C<sub>22</sub>H<sub>24</sub>N<sub>8</sub>O<sub>2</sub> (432.49 g/mol): C, 61.10; H, 5.59; N, 25.91. Found: C, 61.23; H, 6.61; N, 25.74.

### 2.2. X-ray Structure Determinations

The crystal of 4 was immersed in cryo-oil, mounted in a loop at 120 K, and the data were collected on a Rigaku Oxford Diffraction Supernova diffractometer using Cu K $\alpha$  radiation. The *CrysAlisPro*(v.1.171.40.67a) [34] software package was used for cell refinement and data reduction. A gaussian absorption correction (*CrysAlisPro* [34]) was applied to the intensities before structure solution. The structure was solved by intrinsic phasing (*SHELXT* [35]) method. Structural refinement was carried out using *SHELXL*(2018/3) [36] software with *SHELXLE* [37] graphical user interface. The NH and OH hydrogen atoms were located from the difference Fourier map and refined isotropically. All other hydrogen atoms were positioned geometrically and constrained to ride on their parent atoms, with C-H = 0.95–0.99 Å and  $U_{\text{iso}} = 1.2 \cdot U_{\text{eq}}$ (parent atom). The CIF data are given in details in Supplementary data and crystal structure measurement and refinements details are summarized in Table 1. Crystal Explorer 17.5 program [38] was used for Hirshfeld surface analysis.



Scheme 1. Synthesis of 4.

Table 1. Crystal Data of 4.

CCDC	2118553
empirical formula	C <sub>22</sub> H <sub>24</sub> N <sub>8</sub> O <sub>2</sub>
Fw	432.49
temp (K)	120(2)
$\lambda$ (Å)	1.54184
Crystal system	Orthorhombic
space group	Pbca
a (Å)	8.1139(2)
b (Å)	11.2637(2)
c (Å)	45.7836(8)
V (Å <sup>3</sup> )	4184.28(15)
Z	8
$\rho_{\text{calc}}$ (Mg/m <sup>3</sup> )	1.373
$\mu$ (Mo K $\alpha$ ) (mm <sup>-1</sup> )	0.762
No. reflns.	29,199
Unique reflns.	4404
Completeness to $\theta = 67.684^\circ$	99.9%
GOOF (F <sup>2</sup> )	1.047
R <sub>int</sub>	0.0345
R1 <sup>a</sup> (I $\geq$ 2 $\sigma$ )	0.0374
wR2 <sup>b</sup> (I $\geq$ 2 $\sigma$ )	0.0911

<sup>a</sup>  $R1 = \frac{\sum ||F_o| - |F_c||}{\sum |F_o|}$ . <sup>b</sup>  $wR2 = \frac{[\sum [w(F_o^2 - F_c^2)^2]]^{1/2}}{[\sum w(F_o^2)]^{1/2}}$ .

### 3. Computational Details

Density functional calculations (DFT) at the B3LYP method and using 6-31G(d,p) basis sets were performed to optimize the molecular structure of the suggested 17 conformers of the studied *s*-triazine *bis*-Schiff base (Figure 2). For this task, Gaussian 09 package was used [39,40]. All optimized structures are local minimum as indicated from the absence of any imaginary vibrational mode. The calculated energies and thermodynamic

parameters of the suggested conformers were used to predict the most stable form. NBO 3.1 program [41] was used for natural bond orbital analysis.

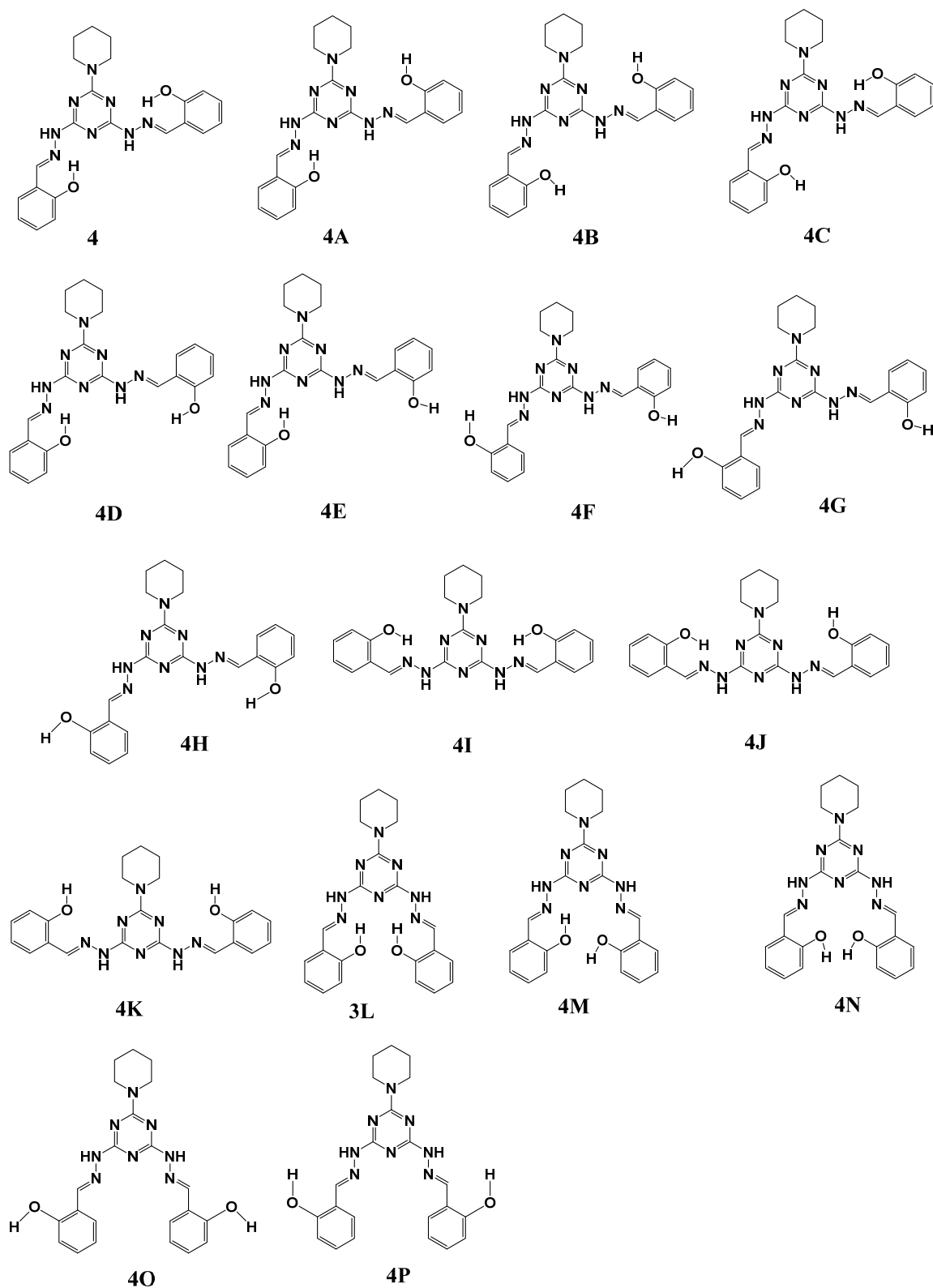


Figure 2. The suggested conformers of the *s*-triazine *bis*-Schiff base; 4.

## 4. Results and Discussion

### 4.1. Crystal Structure Description

The X-ray structure of **4** is shown in Figure 3. The compound crystallized in the orthorhombic crystal system and *Pbca* space group with unit cell parameters of  $a = 8.1139(2) \text{ \AA}$ ,  $b = 11.2637(2) \text{ \AA}$ ,  $c = 45.7836(8) \text{ \AA}$ , and unit cell volume of  $4184.28(15) \text{ \AA}^3$  while  $Z = 4$ . Some bond distances and angles are depicted in Table 2.

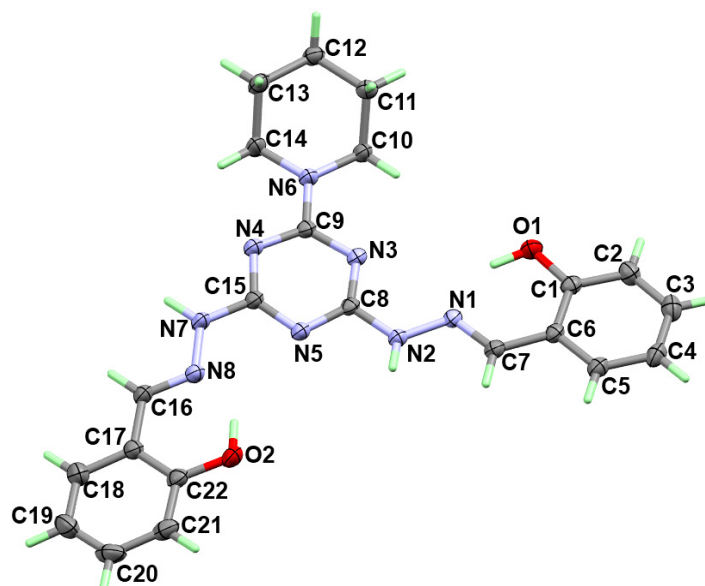


Figure 3. X-ray structure with thermal ellipsoids at 50% probability level for **4**.

Table 2. Selected bond lengths [ $\text{\AA}$ ] and angles [ $^\circ$ ] for **4**.

Atoms	Distance	Atoms	Distance
O1-C1	1.3578(14)	N4-C9	1.3477(15)
O2-C22	1.3624(16)	N5-C15	1.3379(15)
N1-C7	1.2882(15)	N5-C8	1.3445(15)
N1-N2	1.3697(13)	N6-C9	1.3493(15)
N2-C8	1.3691(15)	N6-C10	1.4657(15)
N3-C8	1.3291(15)	N6-C14	1.4707(15)
N3-C9	1.3622(15)	N7-C15	1.3565(15)
N4-C15	1.3394(15)	N7-N8	1.3622(14)
Atoms	Angle	Atoms	Angle
C7-N1-N2	116.50(10)	C10-N6-C14	114.56(10)
C8-N2-N1	121.50(10)	C15-N7-N8	120.60(10)
C8-N3-C9	113.29(10)	C16-N8-N7	117.41(10)
C15-N4-C9	113.93(10)	O1-C1-C2	118.15(11)
C15-N5-C8	112.03(10)	O1-C1-C6	121.81(11)
C9-N6-C10	122.81(10)	C2-C1-C6	120.03(11)
C9-N6-C14	121.91(10)	C3-C2-C1	120.06(12)

It is well known that the *s*-triazine *bis*-Schiff base compounds could exist in two main structures which are shown in Figure 1. The solid-state structure of **4** indicated that the conformation of the two hydrazine arms are quite different and exists in Form 2 rather than Form 1. A main reason for the extrastability of Form 2 is the less steric among the salicylidene moieties. In Form 1, strong repulsion between the two salicylidene arms makes this option is not sterically favored situation.

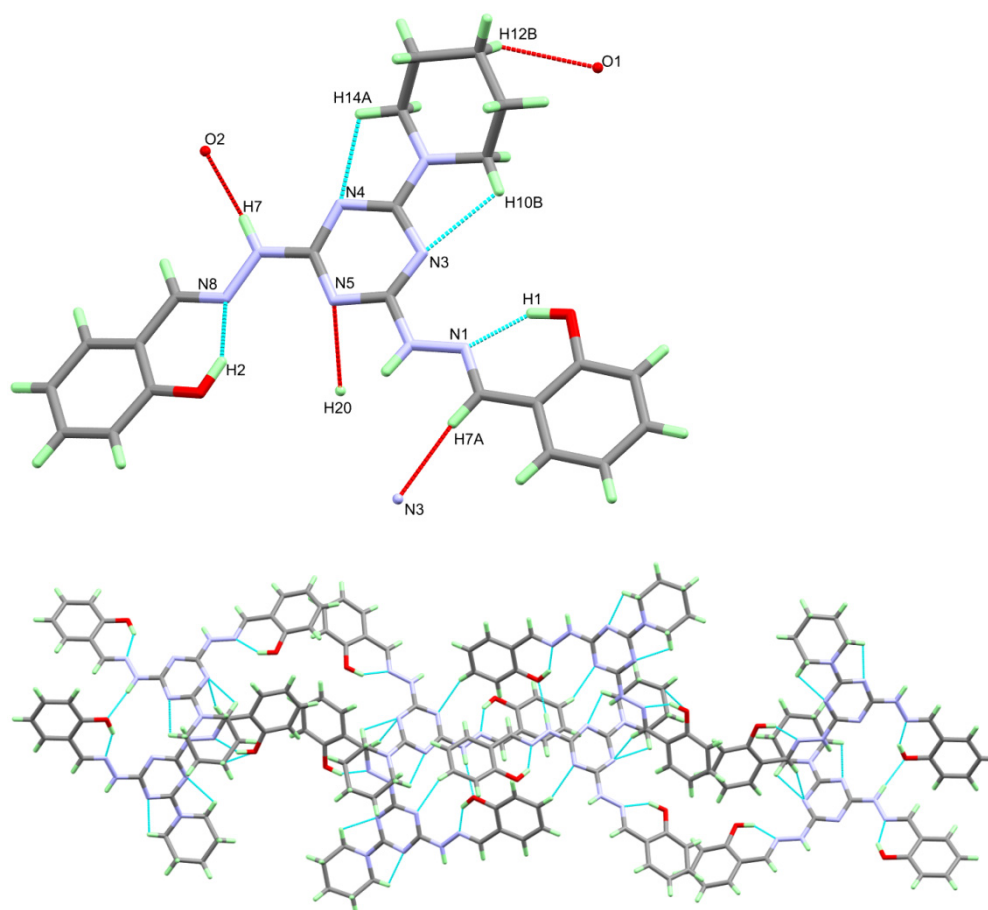
In addition, this molecular conformation of the *bis*-Schiff base *s*-triazine is found stabilized by two strong intramolecular O-H...N hydrogen bonding interactions between the OH group of the salicylidene moiety acting as hydrogen bond donor with the nitrogen

atom of the azomethine group as a hydrogen bond acceptor. The donor-acceptor distances are 2.648(3) and 2.608(1) Å for the O1-H1...N1 and O2-H2...N8 intramolecular O-H...N hydrogen bonds, respectively. Additionally, there are two weak intramolecular C-H...N hydrogen bonds which are C10-H10B...N3 and C14-H14A...N4 with donor-acceptor distances of 2.785(2) and 2.746(2) Å, respectively. More details regarding the hydrogen bond parameters are listed in Table 3. The intramolecular hydrogen bonding interactions are presented as turquoise dotted lines in the upper part of Figure 4, while the red dotted lines represent the intermolecular hydrogen bonding interactions.

**Table 3.** Hydrogen bonds for 4 [Å and °].

D-H...A	d(D-H)	d(H...A)	d(D...A)	<(DHA)
O1-H1...N1	0.958(2)	1.789(2)	2.648(1)	147.6(2)
O2-H2...N8	0.92(2)	1.80(2)	2.608(1)	145.0(2)
N7-H7...O2 <sup>i</sup>	0.927(2)	1.996(2)	2.918(1)	173.9(2)
C7-H7A...N3 <sup>ii</sup>	0.95	2.61	3.494(1)	155.0
C10-H10B...N3	0.99	2.35	2.7846(15)	106.0
C12-H12B...O1 <sup>iii</sup>	0.99	2.57	3.2842(15)	129.0
C14-H14A...N4	0.99	2.29	2.7459(16)	107.0
C20-H20...N5 <sup>iv</sup>	0.95	2.52	3.4165(18)	158.0

<sup>i</sup>  $1/2 - x, 1/2 + y, z$ ; <sup>ii</sup>  $1/2 - x, -1/2 + y, z$ ; <sup>iii</sup>  $3/2 - x, 1/2 + y, z$  and <sup>iv</sup>  $-1/2 + x, 1/2 - y, 1 - z$ .

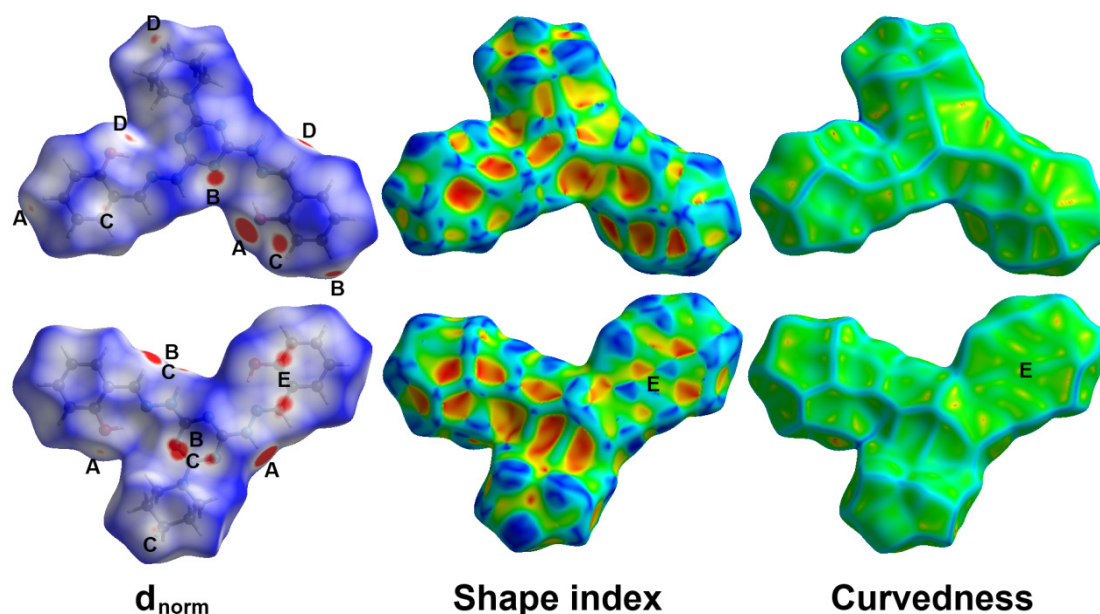


**Figure 4.** Hydrogen bond contacts (**upper**) and molecular packing (**lower**) of 4.

The supramolecular structure of 4 is controlled by strong N7-H7...O2 hydrogen bonding interaction with a hydrogen-acceptor distance of 1.996(2) Å and donor-acceptor distance of 2.918(1) Å. In addition, the molecules are further connected by weak C-H...O interactions (Table 3). Presentation of the molecular packing is given in Figure 4 (lower part).

#### 4.2. Analysis of Molecular Packing

The Hirshfeld surfaces of **4** are shown in Figure 5. The most important contacts having interaction distances shorter than the van der Waals (vdW) radii sum of the two interacting atoms are labeled A to E in the  $d_{\text{norm}}$  map. The O ... H, N ... H, C ... H, H ... H, and C ... C contacts are the most important in the crystal stability. The interaction distances obtained from the Hirshfeld calculations are depicted in Table 4.



**Figure 5.** The  $d_{\text{norm}}$  maps for the important interactions in **4**. A: O ... H, B: N ... H; C: H ... C; D: H ... H, and E: C ... C contacts.

**Table 4.** Intermolecular interactions in **4**.

Contact	Distance	Contact	Distance
H12A ... C5	2.747	H2A ... N4	2.522
H7A ... C9	2.499	H7A ... N3	2.488
H7A ... C9	2.494	H20 ... N5	2.394
H16 ... C22	2.756	O2 ... H7	1.914
H15 ... C21	2.509	O1 ... H3	2.533
C16 ... C21	3.348	H1 ... H12B	2.155
C16 ... C22	3.235	H21 ... H16	2.294 <sup>a</sup>

<sup>a</sup> longer distances compared to the vdWs radii sum.

The percentage contributions for all possible interactions in the crystal are presented graphically in Figure 6. The O ... H, N ... H, C ... H, H ... H and C ... C contributed by 9.8, 15.8, 23.7, 46.4, and 1.6% from the whole fingerprint area, respectively. As can be seen from Figure 5, all these interactions appeared as red colour regions in the  $d_{\text{norm}}$  where the contact distances are shorter than vdW sum of the interacting atoms.

In addition, the decomposed fingerprint plot gave good indication on the importance of these contacts (Figure 7). For example, the O ... H interactions appeared as very sharp spikes indicating strong interactions. The presence of some short C ... C contacts as red spots in the  $d_{\text{norm}}$  with red/blue triangle in the shape index and flat green area in curvedness map revealed very well the presence of  $\pi$ - $\pi$  stacking interactions. The shortest C ... C contacts are C16 ... C21 (3.348 Å) and C16 ... C22 (3.235 Å).



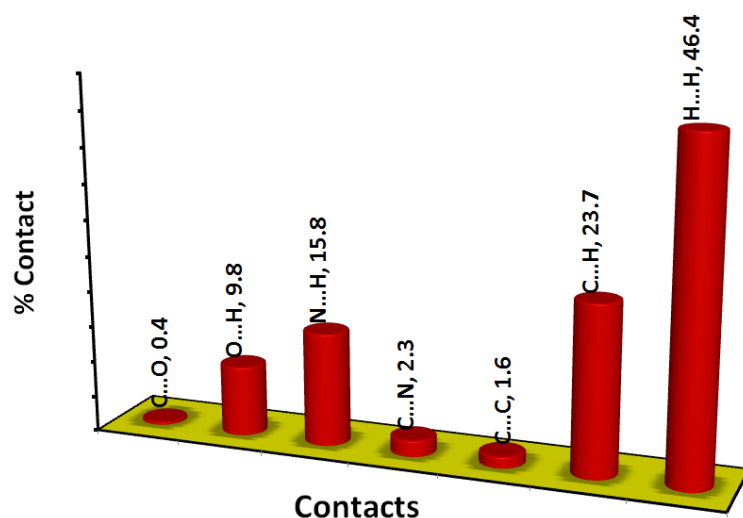


Figure 6. Intermolecular interactions and their percentages in 4.

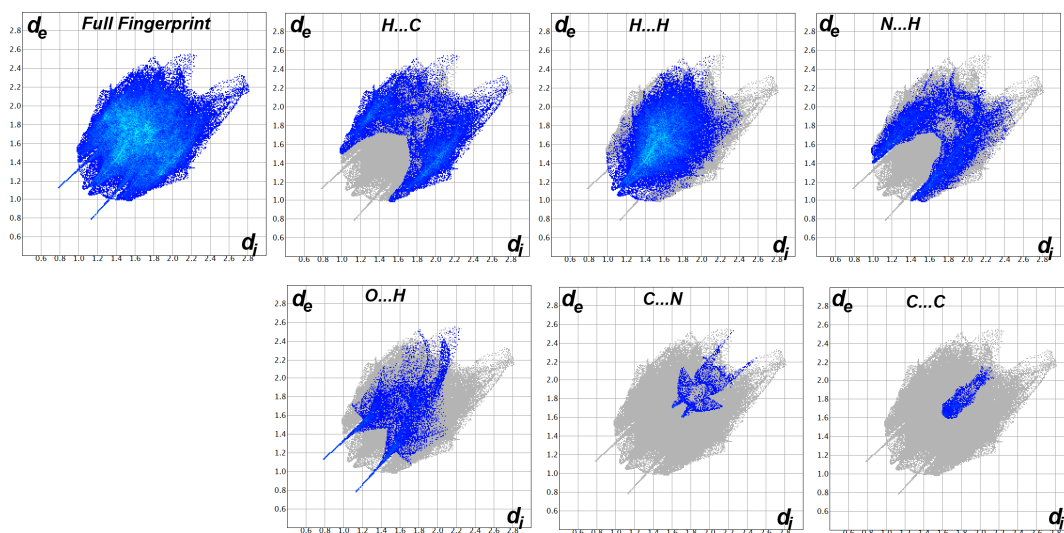
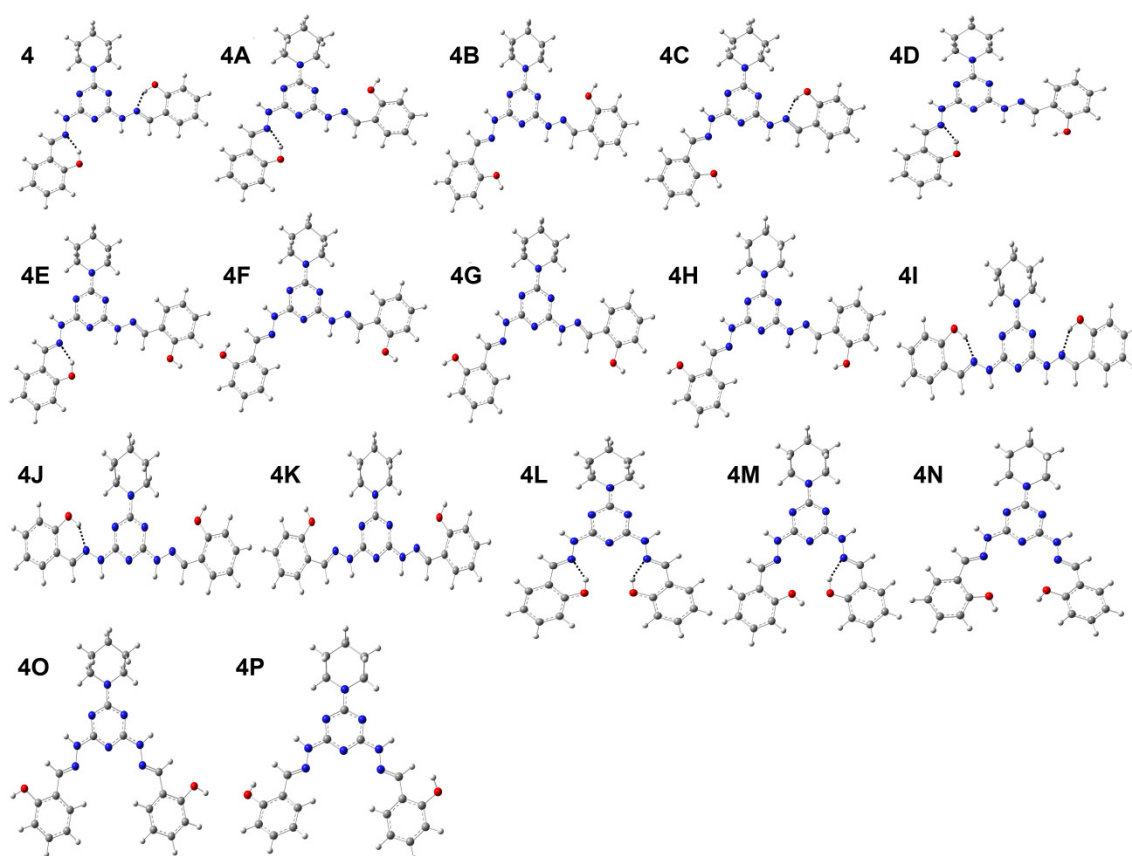


Figure 7. Full and decomposed fingerprint plots for the important interactions in 4.

#### 4.3. Conformational Analysis

The structure of the 17 suggested conformers were calculated and the optimized geometries are shown in Figure 8. The total energies and thermodynamic parameters of the of the studied *s*-triazine *bis*-Schiff base are depicted in Table 5. The results indicated that conformer 4 is the most stable thermodynamically as this conformer has the lowest energy among the studied conformers which is found in accord with the reported X-ray structure of this compound. The second and third most stable conformers are 4I and 4L, respectively. These conformers are energetically higher than the most stable one by only 0.2951 and 2.2345 kcal/mol, respectively. A clear common reason for the extrastability of these three conformers is the presence of two intramolecular O-H . . . N hydrogen bonds in the three structures (Figure 8). On the other hand, some of the other conformers showed one intramolecular O-H . . . N hydrogen while the others did not show any intramolecular O-H . . . N hydrogen bond. The conformers 4B, 4K, and 4N were the highest energetically as these conformers did not comprise any intramolecular O-H . . . N hydrogen bond. The relative energies of the studied conformers are presented graphically in Figure 9.



**Figure 8.** The optimized geometries of the suggested conformers of **4**. The intramolecular O-H ... N hydrogen bonds are presented by black dotted line.

**Table 5.** The calculated energies and thermodynamic properties of the studied conformers.

Param.	<b>4</b>	<b>4A</b>	<b>4B</b>	<b>4C</b>	<b>4D</b>	<b>4E</b>	<b>4F</b>	<b>4G</b>	<b>4H</b>
E <sup>a</sup>	−1441.3046	−1441.2844	−1441.2635	−1441.2841	−1441.286847	−1441.290592	−1441.272653	−1441.276378	−1441.272801
ZPVE <sup>a</sup>	0.4535	0.4526	0.4517	0.4525	0.452338	0.452615	0.451841	0.452084	0.451863
E <sub>tot</sub> <sup>a</sup>	−1440.8511	−1440.8318	−1440.8118	−1440.8316	−1440.8345	−1440.8380	−1440.8208	−1440.8243	−1440.8209
ΔE <sup>b</sup>	0.0000	12.0882	24.6525	12.2360	10.4044	8.2278	18.9990	16.8140	18.9203
H <sup>a</sup>	−1440.8229	−1440.8031	−1440.7827	−1440.8029	−1440.805625	−1440.80919	−1440.791531	−1440.795082	−1440.791662
ΔH <sup>b</sup>	0.0000	12.3787	25.2167	12.5274	10.8118	8.5748	19.6558	17.4276	19.5736
G <sup>a</sup>	−1440.9125	−1440.8946	−1440.8756	−1440.8942	−1440.897779	−1440.900977	−1440.884482	−1440.887956	−1440.884582
ΔG <sup>b</sup>	0.0000	11.2197	23.1635	11.4368	9.2211	7.2144	17.5650	15.3850	17.5022
S <sup>c</sup>	188.6180	192.5080	195.5080	192.2760	193.953	193.183	195.632	195.469	195.568
	<b>4I</b>	<b>4J</b>	<b>4K</b>	<b>4L</b>	<b>4M</b>	<b>4N</b>	<b>4O</b>	<b>4P</b>	
E <sup>a</sup>	−1441.3044	−1441.2844	−1441.2638	−1441.3010	−1441.2878	−1441.2597	−1441.2747	−1441.2672	
ZPVE <sup>a</sup>	0.4538	0.4526	0.4518	0.4534	0.4526	0.4515	0.4521	0.4516	
E <sub>tot</sub> <sup>a</sup>	−1440.8506	−1440.8318	−1440.8120	−1440.8475	−1440.8353	−1440.8081	−1440.8227	−1440.8156	
ΔE <sup>b</sup>	0.2951	12.1058	24.5161	2.2345	9.9103	26.9448	17.8395	22.2519	
H <sup>a</sup>	−1440.8228	−1440.8031	−1440.7830	−1440.8192	−1440.8069	−1440.7790	−1440.7935	−1440.7863	
ΔH <sup>b</sup>	0.0201	12.3743	25.0285	2.2916	10.0362	27.5096	18.4359	22.9514	
G <sup>a</sup>	−1440.9125	−1440.8939	−1440.8750	−1440.9093	−1440.8955	−1440.8715	−1440.8856	−1440.8790	
ΔG <sup>b</sup>	0.0094	11.6414	23.5444	1.9785	10.6518	25.7024	16.8559	21.0106	
S <sup>c</sup>	188.6540	191.0780	193.5960	189.6690	186.5540	194.6810	193.9180	195.1290	

<sup>a</sup> A.U. <sup>b</sup> kcal/mol <sup>c</sup> Cal/Mol. K.

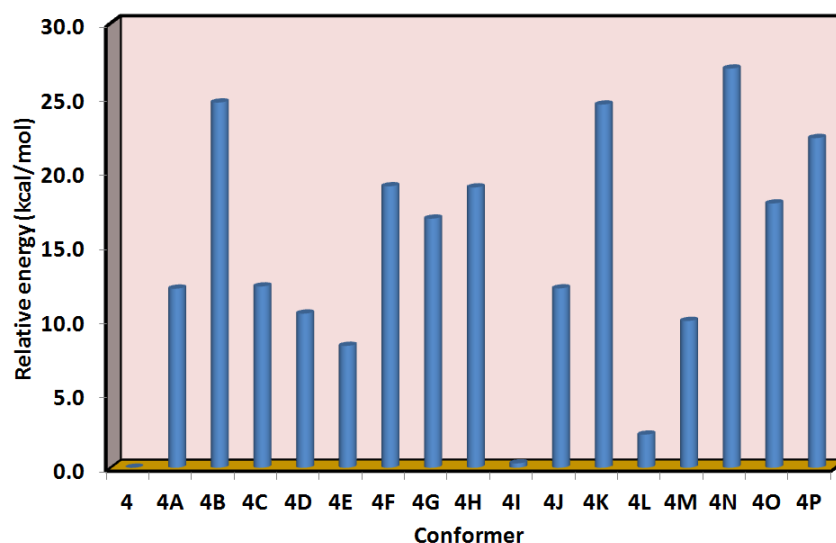


Figure 9. Relative energy of the suggested conformers of 4 compared to the most stable one.

#### 4.4. Optimized Geometry

The calculated structure of 4 is shown in Figure 10. Few differences between the experimental and optimized structures were detected where the deviations in bond distances not exceed 0.02 Å (Table S1; Supplementary Data). These deviations could be attributed to the crystal packing effects. Generally, there are good correlations between the calculated bond angles and distances with the experimental measurements (Figure 11).

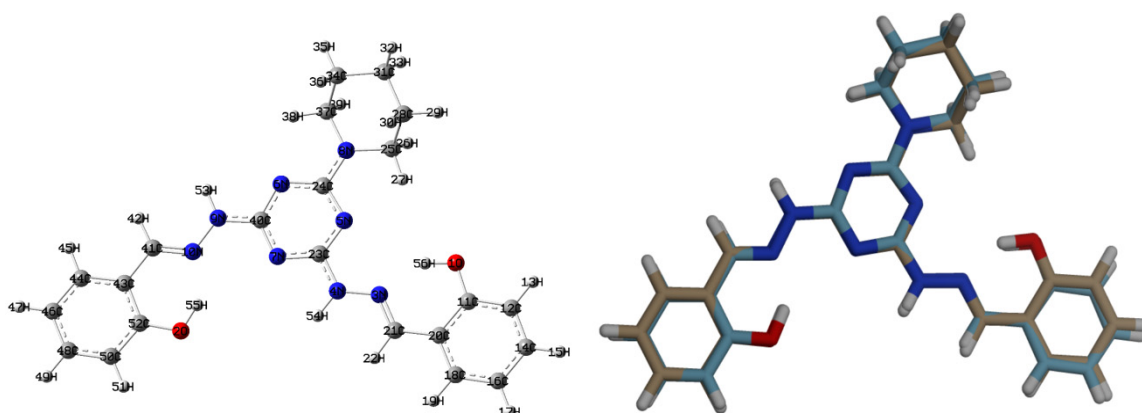


Figure 10. The calculated structure (left) of 4 and its overlay with the experimental one (right).

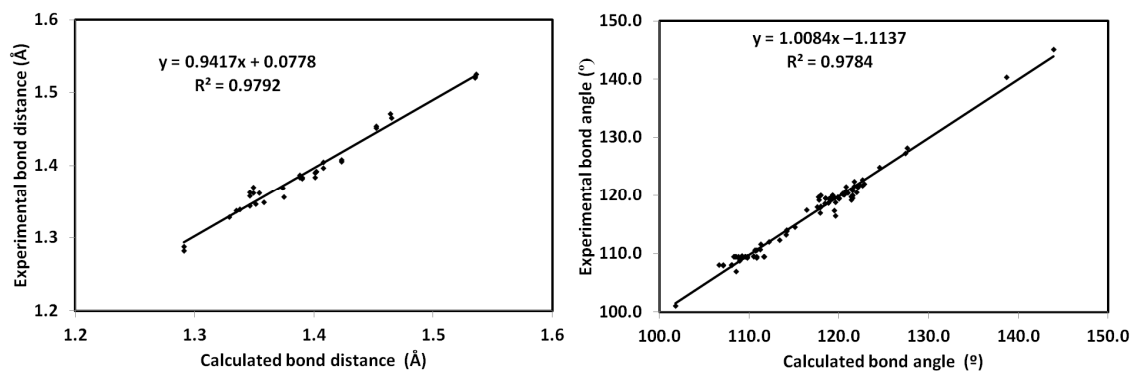


Figure 11. Correlations between the calculated and experimental geometric parameters.

The studied system comprised CHNO skeleton where the oxygen and nitrogen atoms as well as the majority of carbon atoms are negatively charged (Figure 12). The morpholine oxygen atoms and the *s*-triazine N-atoms have the highest negative charge. The studied compound is polar molecule (2.7591 Debye) and the dipole moment vector is shown in Figure 13 (left part). The red regions in molecular electrostatic potential reveal high electron density related to hydroxyl oxygen atoms and the aryl moieties. In contrast, the blue regions are related to the atomic sites with lowest electron density which is close to the NH protons.

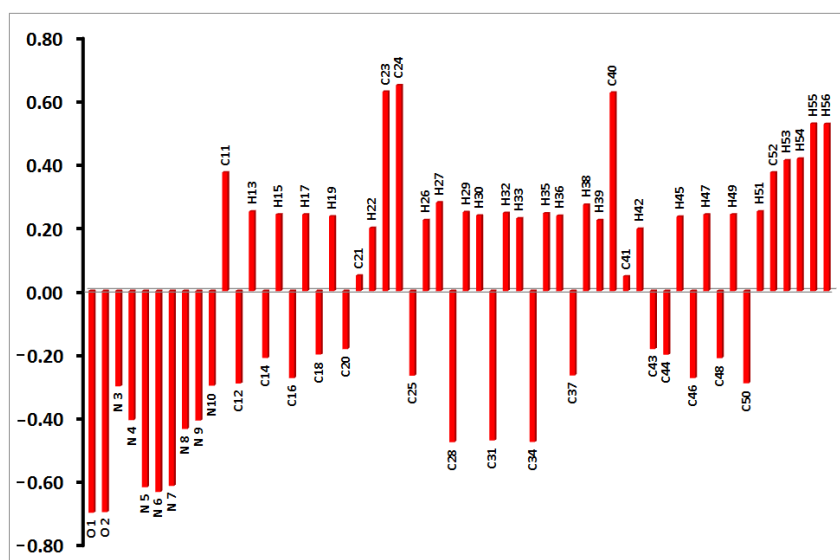


Figure 12. The natural atomic charges in 4.

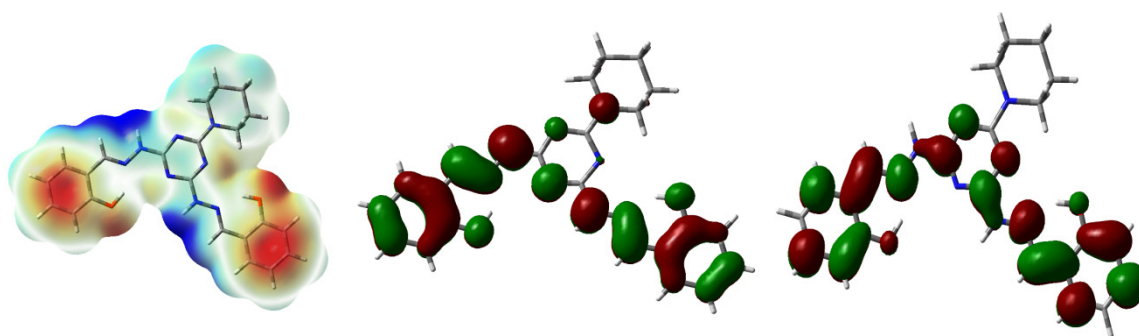


Figure 13. The MEP, HOMO, and LUMO of 4.

It is clear that the HOMO and LUMO levels are mainly distributed over the  $\pi$ -system of 4 (Figure 13). As a result, the HOMO $\rightarrow$ LUMO excitation is mainly a  $\pi$ - $\pi^*$  transition. The HOMO $\rightarrow$ LUMO excitation energy is 4.1536 eV. In addition, the reactivity indices include ionization potential ( $I = -E_{\text{HOMO}}$ ), electron affinity ( $A = -E_{\text{LUMO}}$ ), chemical potential ( $\mu = -(I + A)/2$ ), hardness ( $\eta = (I - A)/2$ ), as well as electrophilicity index ( $\omega = \mu^2/2\eta$ ) [42–47]. The calculated values of these descriptors are 5.5391, 1.3856,  $-3.4624$ , 4.1536, and 1.4431, respectively.

#### 4.5. NBO Analysis

The intramolecular charge transfer (IMCT) plays very important rule in the stability of compound. In this regard, the different IMCT processes ( $\sigma$ - $\sigma^*$ ,  $\pi$ - $\pi^*$ ,  $n$ - $\sigma^*$  and  $n$ - $\pi^*$ ) in 4 were calculated [48,49] and their stabilization energies ( $E^{(2)}$ ) are listed in Table 6. The maximum interaction energy due to the  $\sigma$ - $\sigma^*$  intramolecular charge transfer (IMCT) is 5.99 kcal/mol for the BD(1)N7-C23 $\rightarrow$ BD\*(1)N9-C40. The  $\pi$ - $\pi^*$ ,  $n$ - $\pi^*$  and

$n \rightarrow \sigma^*$  IMCT processes have higher interaction energies with maximum  $E^{(2)}$  values of 48.26, 71.52, and 24.48 kcal/mol for the  $BD(2)N5-C24 \rightarrow BD^*(2)N7-C23$ ,  $LP(1)N8 \rightarrow BD^*(2)N5-C24$ , and  $LP(1)N3 \rightarrow BD^*(1)O1-H56$  IMCT, respectively. Interestingly, the presence of  $LP(1)N3 \rightarrow BD^*(1)O1-H56$  and  $LP(1)N10 \rightarrow BD^*(1)O2-H55$  IMCT processes with high  $E^{(2)}$  values of 24.48 and 23.08 kcal/mol, respectively confirmed the stability of conformer **4** via intramolecular O-H...N hydrogen bonding interactions.

**Table 6.** The IMCT processes and  $E^{(2)}$  values calculated using NBO method.

NBO <sub>i</sub>	NBO <sub>j</sub>	E(2)	NBO <sub>i</sub>	NBO <sub>j</sub>	E(2)
$\sigma \rightarrow \sigma^*$			$\pi \rightarrow \pi^*$		
BD(1)O1-H56	BD*(1)C11-C12	5.06	BD(2)N3-C21	BD*(2)C18-C20	7.87
BD(1)O2-H55	BD*(1)C50-C52	5.02	BD(2)N5-C24	BD*(2)N 7-C23	48.26
BD(1)N5-C24	BD*(1)N 4-C23	5.63	BD(2)N6-C40	BD*(2)N 5-C24	39.11
BD(1)N7-C23	BD*(1)N 9-C40	5.99	BD(2)N7-C23	BD*(2)N 6-C40	45.32
BD(1)C20-C21	BD*(1)N 3-N 4	5.22	BD(2)N10-C41	BD*(2)C43-C44	7.88
BD(1)C41-C43	BD*(1)N 9-N10	5.20	BD(2)C11-C12	BD*(2)C14-C16	25.09
			BD(2)C11-C12	BD*(2)C18-C20	17.00
			BD(2)C14-C16	BD*(2)C11-C12	17.55
			BD(2)C14-C16	BD*(2)C18-C20	24.64
			BD(2)C18-C20	BD*(2)N 3-C21	24.97
			BD(2)C18-C20	BD*(2)C11-C12	21.73
$n \rightarrow \sigma^*$			$n \rightarrow \pi^*$		
LP(1)O 1	BD*(1)C11-C20	7.96	BD(2)C18-C20	BD*(2)C14-C16	17.03
LP(1)O 2	BD*(1)C43-C52	8.02	BD(2)C43-C44	BD*(2)N10-C41	24.95
LP(1)N 3	BD*(1)O 1-H56	24.48	BD(2)C43-C44	BD*(2)C46-C48	17.02
LP(1)N 3	BD*(1)N 4-H54	8.74	BD(2)C43-C44	BD*(2)C50-C52	21.68
LP(1)N 3	BD*(1)C21-H22	9.92	BD(2)C46-C48	BD*(2)C43-C44	24.65
LP(1)N 5	BD*(1)N 7-C23	13.54	BD(2)C46-C48	BD*(2)C50-C52	17.57
LP(1)N 5	BD*(1)N 6-C24	12.09	BD(2)C50-C52	BD*(2)C43-C44	17.07
LP(1)N 6	BD*(1)N 5-C24	11.84	BD(2)C50-C52	BD*(2)C46-C48	25.10
LP(1)N 6	BD*(1)N 7-C40	13.29	LP(2)O 1	BD*(2)C11-C12	35.52
LP(1)N 7	BD*(1)N 5-C23	12.74	LP(2)O 2	BD*(2)C50-C52	35.53
LP(1)N 7	BD*(1)N 6-C40	12.66	LP(1)N 4	BD*(2)N 3-C21	31.38
LP(1)N10	BD*(1)O 2-H55	23.08	LP(1)N 4	BD*(2)N 7-C23	53.44
LP(1)N10	BD*(1)C41-H42	10.00	LP(1)N 8	BD*(2)N 5-C24	71.52
LP(1)N10	BD*(1)N 9-H53	8.82	LP(1)N 9	BD*(2)N 6-C40	52.75
			LP(1)N 9	BD*(2)N10-C41	31.48

## 5. Conclusions

The synthesis and X-ray structure of the new 2,2'-(6-(piperidin-1-yl)-1,3,5-triazine-2,4-diyl)bis(hydrazin-2-yl-1-ylidene))bis(methanylylidene)diphenol giant *s*-triazine *bis*-Schiff base were presented. Among the suggested 17 conformers, **4** was found to be the most stable one in agreement with the reported X-ray structure. Generally, conformers with larger number of intramolecular O-H...N hydrogen bonding interactions are the most stable. The presence of intramolecular O-H...N hydrogen bonds was further revealed by NBO calculations. Using Hirshfeld calculations, the O...H, N...H, C...H, H...H, and C...C intermolecular contacts are the most important in the crystal stability. Their percentages were calculated to be 9.8, 15.8, 23.7, 46.4, and 1.6% from the whole contacts occurred in the crystal, respectively. MEP, HOMO, LUMO, dipole moment, and natural charges were also presented.

**Supplementary Materials:** The following are available online at <https://www.mdpi.com/article/10.3390/cryst11111418/s1>, Table S1: The calculated and experimental bond distances and angles, Table S2: Bond lengths [Å] and angles [°] for **4**, Table S3: Anisotropic displacement parameters ( $\text{\AA}^2 \times 10^3$ ) for **4**. The anisotropic displacement factor exponent takes the form:  $-2p^2[h^2 a^{*2}U^{11} + \dots$

+ 2 h k a\* b\* U<sup>12</sup>], Table S4: Hydrogen coordinates ( $\times 10^4$ ) and iso-tropic displacement parameters ( $\text{\AA}^2 \times 10^3$ ) for **4**, Table S5: Torsion angles [ $^\circ$ ] for **4**, Figure S1: <sup>1</sup>H and <sup>13</sup>C NMR spectra of **4**.

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