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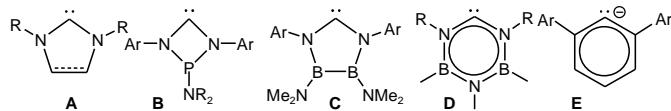
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A σ -Donor with a Planar Six- π -Electron $B_2N_2C_2$ Framework: Anionic N-heterocyclic Carbene or Heterocyclic Terphenyl Anion? **

Taryn D. Forster, Kelly E. Krahulic, Heikki M. Tuononen, Robert McDonald, Masood Parvez, Roland Roesler*

Dedicated to Professor Tristram Chivers

The remarkable success of N-heterocyclic carbenes^[1] (NHCs) as ligands in main group and especially in transition metal chemistry and catalysis^[2] has prompted the synthesis and characterization of numerous representatives of this class of compounds. Derivatives with a five-membered framework (**A**) have been investigated most extensively. However four-, six- and seven-membered carbenes were also reported. Emerging trends in carbene ligand design involve the replacement of one or both intraannular nitrogen atoms with elements such as carbon, phosphorus, and sulphur,^[3] and the replacement of the carbon backbone with heteroelements (**B – D**).^[4] In this way, the steric and electronic properties of the ligand can be effectively tuned. Although NHCs are traditionally neutral ligands, a few anionic derivatives have been prepared in which the charge is localized on an adjoining borate or cyclopentadienyl moiety.^[5] Herein we report the synthesis of an anionic NHC (**3**) having a six π -electron system with the charge localized on the ring framework. This compound has the potential to coordinate in either σ (η^1) or π (η^6) fashion.



The reaction of the corresponding trimethylsilyl formamidine^[4c] with 1,1-bis(methylchloroboryl)ethane^[6] in the presence of trimethylsilyl triflate (Scheme 1) yielded the zwitterionic ring compound **1**. Crystallographic analysis (Figure 1) revealed an iminium borate structure with the triflate coordinated to

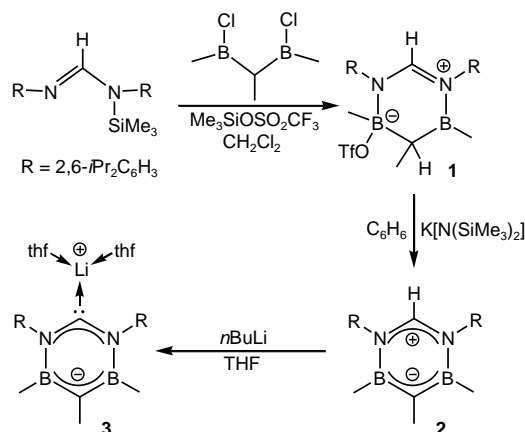
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one of the boron atoms.^[7] The $B_2N_2C_2$ ring in **1** has a half-boat conformation with the sp^3 hybridized carbon C(2) situated outside of the slightly



Scheme 1. Synthesis of derivatives **1 – 3**.

twisted B_2N_2C (1) plane. The B-C and B-N bonds involving the borate center B(1) are longer than the corresponding bonds to borane center B(2), and the two N-C(1) bonds within the ring are not equal. The ¹¹B NMR spectrum of **1** in benzene displayed only one signal at $\delta = 40.7$ ppm, which corresponds to a three coordinated boron and indicates that the triflate dissociates in solution to form an ion pair. The ¹H and ¹³C spectra are in agreement with C_s symmetry.

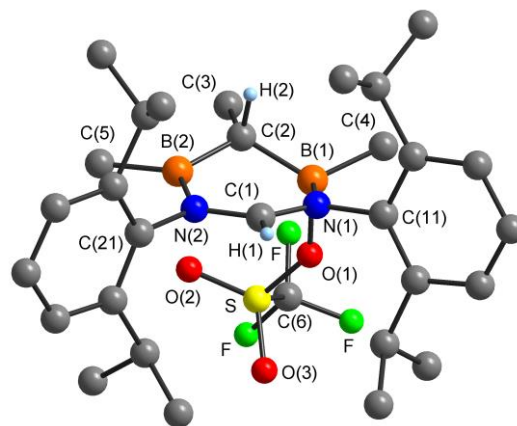


Figure 1. Molecular structure of **1**. Hydrogen atoms on the organic substituents have been omitted for clarity. Selected bond lengths [Å] and angles [°]: N(1)-C(1) 1.298(2), N(2)-C(1) 1.345(2), N(1)-B(1) 1.595(3), N(2)-B(2) 1.475(3), C(2)-B(1) 1.603(3), C(2)-B(2) 1.550(3), O(1)-B(1) 1.626(3), N(1)-C(1)-N(2) 125.74(17), C(1)-N(1)-B(1) 121.48(16), C(1)-N(2)-B(2) 122.23(16), B(1)-C(2)-B(2) 113.66(16).

Deprotonation of **1** with an equimolar amount of $K[N(SiMe_3)_2]$ cleanly yielded the 1,5-diaza-2,4-diborane **2**. Few carbaborazines have been described, and a derivative with this framework has not been reported; however its stability and properties have been the object of theoretical calculations.^[8] At -40 °C, the ¹³C NMR spectrum featured signals corresponding to the ring carbon atoms adjacent to nitrogen and boron at $\delta = 146.8$ and 119.2 ppm, respectively, in the typical range for benzene derivatives.

The X-ray structural determination of **2** (Figure 2) confirmed the proposed structure, which contains a planar B₂N₂C₂ ring with equivalent pairs of B-N, B-C and N-C bond lengths.^[7] The endocyclic N-C bonds lengths (ca. 1.33 Å) are intermediate in length between the two values observed in **1**, and practically identical to the bonds observed in the cationic precursors to carbenes **B** – **D**.^[4] The B-N bond lengths in **2** (1.50 Å) are similar to the those observed in **C** and significantly longer than the those in borazines (1.42 – 1.44 Å).^[9] The intraannular B-C and N-C bonds (1.48 and 1.33 Å, respectively) are shorter than the corresponding bonds observed in the regioisomers 1,3-diaza-2,4-diborane and 1,4-diaza-2,3-diborane (ca. 1.53 and 1.38–1.40 Å, respectively).^[10] The B-N bonds in the latter derivatives are equal with those in borazines. As proposed by Bertrand and co-workers for the cationic precursor to **D**,^[4d] the six π electrons in **2** appear to be distributed onto two allyl-like fragments (B₂C⁻ and N₂C⁺) to form a zwitterionic structure, rather than delocalized over the entire ring framework.

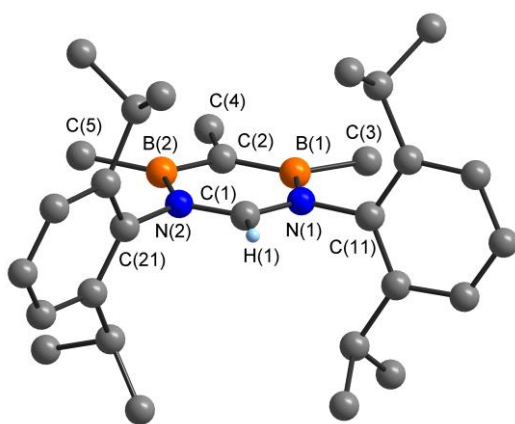


Figure 2. Molecular structure of **2**. Hydrogen atoms on the organic substituents have been omitted for clarity. Selected bond lengths [Å] and angles [°]: N(1)-C(1) 1.3264(17), N(2)-C(1) 1.3274(17), N(1)-B(1) 1.5027(19), N(2)-B(2) 1.4980(19), C(2)-B(1) 1.477(2), C(2)-B(2) 1.479(2), N(1)-C(1)-N(2) 122.83(12).

Deprotonation of **2** with *n*BuLi produced the lithium salt **3**, which was stable in solid state under an inert atmosphere but decomposed slowly in tetrahydrofuran (thf) through deprotonation of the solvent and reformation of the starting material. The crystallographic determination (Figure 3) revealed that **3** contains a planar B₂N₂C₂ ligand σ -coordinated to the lithium ion.^[7] The distorted trigonal planar coordination environment of lithium is completed by two thf molecules, and the ring carbon atoms and the lithium atom lie on a crystallographic C₂ axis. Consistent with other NHCs, deprotonation of **2** results in lengthening of the intraannular N-C bonds by approximately 0.03 Å to a value of 1.36 Å. The N(1)-C(1)-N(1') angle in **3** (114.0°) is more acute than in **2** (122.8°) and **1** (125.7°). This is a result of a lateral compression of the ring skeleton, causing an increase of the C(1)⋯C(2) distance of 0.13 Å in **3** versus **2**. The overall geometry of **3** strongly resembles the structure of isoelectronic lithium terphenyl derivatives.^[11] The planar six- π -electron σ -donating anion in **3** could therefore be seen as either an anionic NHC or a heterocyclic terphenyl anion (**E**).

The Li-C bond length and the chemical shift of the carbon atom connected to lithium can be considered when comparing **3** to a lithium NHC complex and a lithium terphenyl. The Li(1)-C(1) bond measures 2.152(6) Å, which is typical for lithium NHC complexes.

In representatives of this class, the substitution of the lithium center or the charge of the NHC ligand have little influence on the length of the Li-C bond (2.124(4) – 2.197(4) Å).^[12] Conversely, the length of the Li-C bond in trigonal-planar, three-coordinate terphenyl lithium derivatives falls within a narrow range (2.074(16) – 2.128(4) Å).^[13] The ¹³C NMR signal corresponding to the carbene carbon could not be located in the spectrum of **3**. However, the potassium analog, obtained by deprotonation of **2** with benzyl potassium, displayed a signal for the carbene carbon at 239.1 ppm, which is significantly downfield shifted in comparison to the *ipso* carbon in terphenyl lithium derivatives (175 – 201 ppm)^[13] and the carbene carbon atom in NHC-Li (189 – 198 ppm)^[12] and NHC-K complexes^[14] (199 – 208 ppm). It is uncertain if this signal belongs to the free anionic carbene or the potassium-coordinated ligand. To our knowledge, potassium terphenyl derivatives have not been reported.

A density functional theory treatment of the simplified structures **I** – **III** showed an increase in the energy of both the highest occupied molecular orbitals (HOMOs; σ) and the lowest unoccupied molecular orbitals (LUMOs; π) in the order of **I** to **II** to **III**, which indicates that the σ -donating ability decreases in order **III** > **II** >> **I** and the π -accepting ability decreases in reverse order (Figure 4). NHCs are generally considered to be excellent σ -donors and poor π -acceptors. Analysis of the electron localization function (ELF)^[15] for the compounds in question gives occupancies of 2.28, 2.35, and 2.45 electrons for the carbon lone pair basins in **III**, **II** and **I**, respectively. Hence, the ELF analysis suggests that NHC **I** is a better σ -donor than is predicted by the orbital energies alone. The same is implied by the trends in Mulliken atomic charges, which show that the carbene carbon atom in NHC **I** bears slightly more negative charge (-0.20) than that in **II** (-0.15) or **III** (-0.14). To fully resolve the issue of relative σ -donating ability of the studied systems, we performed charge decomposition analyses (CDA)^[16] for complexes of **I–III** with Li⁺ and CuCl. The results show that the σ -donating ability of ligands **I–III** follows the trend in orbital energies for the Li⁺ complexes (total donation of σ electrons is 0.484, 0.566, and 0.622 for **I**, **II**, and **III**, respectively), but within the CuCl series ligand **II** has similar σ -donating ability to **I** (total donation of 0.236 and 0.233 electrons, respectively) and **III** has slightly less (0.206 electrons). Hence, the computational data indicate that the new anionic carbene **II** has similar or even greater σ -donating ability, depending on the metal fragment, than the traditional ligand **I**.

In summary, experimental and theoretical evidence indicates that the anionic carbene in **3** bridges the gap between two classical systems that have quite different properties: NHCs and terphenyl anions. Its facile synthesis from the formally zwitterionic diazadiborane **2** renders the novel system a promising anionic ligand, with which both σ (η^1) and π (η^3 or η^6) coordination modes are possible. Its coordinating ability towards transition metals is currently being investigated.

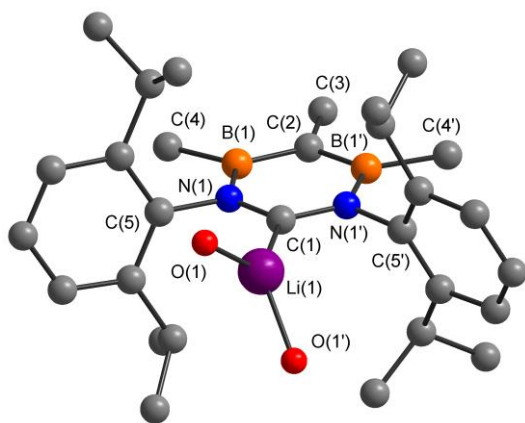


Figure 3. Molecular structure of **3**. Hydrogen atoms on the organic substituents and the carbon atoms in the thf molecules have been omitted for clarity. Selected bond lengths [Å] and angles [°]: C(1)-Li(1) 2.152(6), N(1)-C(1) 1.3627(19), N(1)-B(1) 1.495(3), C(2)-B(1) 1.475(3), N(1)-C(5) 1.450(2), C(4)-B(1) 1.599(3), C(2)-C(3) 1.518(4), O(1)-Li(1) 1.946(4), N(1)-C(1)-N(1') 114.0(2), N(1)-C(1)-Li(1) 123.01(10), O(1)-Li(1)-C(1) 130.48(13), O(1)-Li(1)-O(1') 99.0(3).

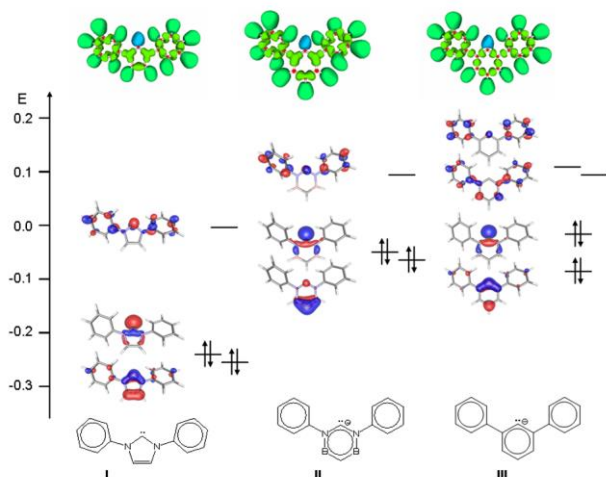


Figure 4. Frontier molecular orbitals and ELF diagrams for systems I – III.

Experimental Section

All operations were carried out with exclusion of air and moisture. The NMR spectra were run on a Bruker Avance DRX-400 spectrometer. All chemicals were prepared according to reported procedures or purchased from commercial sources. Computational details are presented in Supporting Information.

1: 1,1-bis(chloromethylboryl)ethane (345 mg, 2.3 mmol) was slowly added to a solution of N,N'-bis(2,6-diisopropylphenyl)trimethylsilylformamide (1.0 g, 2.3 mmol) in dichloromethane (20 mL) at room temperature. After stirring for one hour, trimethylsilyltriflate (0.5 g, 2.2 mmol) was added and the solution was stirred for an additional 30 minutes. Volatiles were removed *in vacuo* and the oily residue was recrystallized from hexane, yielding **1** as a colorless, crystalline solid (0.82 g, 1.14 mmol, 61 %). ¹H NMR (400 MHz, CD₂Cl₂, 25 °C): δ = 0.21 (s, 6H, BCH₃), 0.87 (q, 1H, B₂CHCH₃, ³J_{HH} = 7.0 Hz), 1.14, 1.15 (d, 6H, CH(CH₃)₂, ³J_{HH} = 6.9 Hz), 1.24, 1.25 (d, 6H, CH(CH₃)₂, ³J_{HH} = 6.8 Hz), 1.32 (d, 3H, B₂CHCH₃, ³J_{HH} = 7.0 Hz), 2.94 (sept, 2H, CH(CH₃)₂, ³J_{HH} = 6.9 Hz), 3.24 (sept, 2H, CH(CH₃)₂, ³J_{HH} = 6.8 Hz),

7.23, 7.27 (dd, 2H, *m*-C₆H₃, ³J_{HH} = 7.8 Hz, ⁴J_{HH} = 1.5 Hz), 7.37 (t, 2H, *p*-C₆H₃, ³J_{HH} = 7.8 Hz), 7.57 (s, 1H, N₂C(H)); ¹³C NMR (100 MHz, CD₂Cl₂, -50 °C): δ = 0.8 (B₂CHCH₃), 4.5 (BCH₃), 9.9 (B₂CHCH₃), 24.1, 24.2, 25.6, 25.8 (CH(CH₃)₂), 28.5, 29.4 (CH(CH₃)₂), 124.2, 124.6 (*m*-C₆H₃), 128.9 (*p*-C₆H₃), 136.6 (*i*-C₆H₃), 144.0, 144.8 (*o*-C₆H₃), 160.9 (N₂CH); ¹¹B NMR (128 MHz, C₆D₆, 25 °C): δ = 40.7 ppm; MS (EI, 70 eV): *m/z* = 443(8) [M⁺ - O₃SCF₃], 427(100) [M⁺ - O₃SCF₃ - CH₃]; HRMS (EI, 70 eV): found *m/z* = 443.3727 [M⁺ - O₃SCF₃], calcd. for H₄₅C₂₉N₂¹¹B₂ m = 443.3769.

2: A mixture of **1** (1.0 g, 1.69 mmol) and K[N(SiMe₃)₂] (0.36 g, 1.8 mmol) was stirred in toluene (30 mL) for 1 h at room temperature and then filtered. Volatiles were removed *in vacuo*, leaving behind **2** as a crystalline, colorless solid (0.64 g, 1.45 mmol, 86%). ¹H NMR (400 MHz, CD₂Cl₂, 25 °C): δ = 0.26 (s, 6H, BCH₃), 1.12, 1.18 (d, 12H, CH(CH₃)₂, ³J_{HH} = 6.8 Hz), 2.06 (s, 3H, B₂CHCH₃), 2.75 (sept, 4H, CH(CH₃)₂, ³J_{HH} = 6.8 Hz), 7.24 (d, 4H, *m*-C₆H₃, ³J_{HH} = 7.6 Hz), 7.36 (t, 2H, *p*-C₆H₃, ³J_{HH} = 7.6 Hz), 7.57 (s, 1H, N₂CH); ¹³C NMR (100 MHz, CD₂Cl₂, -40 °C): δ = 0.1 (br, BCH₃), 16.9 (B₂CCH₃), 23.6, 24.8 (CH(CH₃)₂), 27.9 (CH(CH₃)₂), 119.2 (br, B₂CCH₃), 123.8 (*m*-C₆H₃), 128.2 (*p*-C₆H₃), 139.1 (*i*-C₆H₃), 144.1 (*o*-C₆H₃), 146.8 (N₂CH); ¹¹B NMR (128 MHz, CD₂Cl₂, 25 °C): δ = 38.4 ppm; MS (EI, 70 eV): *m/z* = 442(24) [M⁺], 427(80) [M⁺ - CH₃]; HRMS (EI, 70 eV): found *m/z* = 442.3656, calcd. for H₄₄C₂₉N₂¹¹B₂ m = 442.3691.

3: Solid *n*BuLi (5 mg, 0.9 μmol) was added at room temperature to a solution of **2** (20 mg, 0.45 μmol) in THF (1 mL). The solvent was removed *in vacuo* and the colorless product (19 mg, 0.32 μmol, 71 %) was recrystallized from toluene, yielding colorless crystals. ¹H NMR (400 MHz, THF-*d*₈, 25 °C): δ = 0.02 (s, 6H, BCH₃), 1.09, 1.12 (d, 12H, CH(CH₃)₂, ³J_{HH} = 6.9 Hz), 1.89 (s, 3H, B₂CHCH₃), 3.23 (sept, 4H, CH(CH₃)₂, ³J_{HH} = 6.9 Hz), 7.15 (s, 6H, C₆H₃); ¹³C NMR (100 MHz, THF-*d*₈, 25 °C): δ = 2.3 (s, br, BCH₃), 17.8 (B₂CCH₃), 25.2, 25.4 (CH(CH₃)₂), 28.5 (CH(CH₃)₂), 124.1 (*m*-C₆H₃), 126.4 (*p*-C₆H₃), 146.2 (*o*-C₆H₃), 149.7 (*i*-C₆H₃); ¹¹B NMR (128 MHz, THF-*d*₈, 25 °C): δ = 36.3 ppm; ⁷Li NMR (155 MHz, THF-*d*₈, 25 °C): δ = 0.15 ppm.

- [1] a) D. Bourissou, O. Guerret, F. P. Gabbaï, G. Bertrand, *Chem. Rev.* **2000**, *100*, 39 – 91; b) *Carbene Chemistry*, G. Bertrand, Ed., Marcel Dekker, New York, **2002**; c) R. W. Alder, M. E. Blake, L. Chaker, J. N. Harvey, F. Paolini, J. Schütz, *Angew. Chem. Int. Ed.* **2004**, *43*, 5896 – 5911; d) V. Nair, S. Bindu, V. Sreekumar, *Angew. Chem. Int. Ed.* **2004**, *43*, 5130 – 5135; e) D. Enders, T. Balensiefer, *Acc. Chem. Res.* **2004**, *37*, 534 – 541.
- [2] a) L. Jafarpour, S. P. Nolan, *J. Organomet. Chem.* **2001**, *617* – 618, 17 – 27; b) D. Enders, H. Gielen, *J. Organomet. Chem.* **2001**, *617* – 618, 70 – 80; c) W. A. Herrmann, *Angew. Chem. Int. Ed.* **2002**, *41*, 1290 – 1309; d) A. C. Hillier, G. A. Grasa, M. S. Viciu, H. M. Lee, C. Yang, S. P. Nolan, *J. Organomet. Chem.* **2002**, *653*, 69 – 82; e) M. C. Perry, K. Burgess, *Tetrahedron: Asymmetry* **2003**, *14*, 951 – 961; f) E. Peris, R. H. Crabtree, *Coord. Chem. Rev.* **2004**, *248*, 2239 – 2246; g) C. M. Crudden, D. P. Allen, *Coord. Chem. Rev.* **2004**, *248*, 2247 – 2273; h) V. César, S. Bellemin-Laponnaz, L. H. Gade, *Chem. Soc. Rev.* **2004**, *33*, 619 – 636; i) N. M. Scott, S. P. Nolan, *Eur. J. Inorg. Chem.* **2005**, 1815 – 1828.
- [3] F. E. Hahn, *Angew. Chem. Int. Ed.* **2006**, *45*, in press; b) H. Nikawa, T. Nakahodo, T. Tsuchiya, T. Wakahara, G. M. A. Rahman, T. Akasaka, Y. Maeda, M. T. H. Liu, A. Meguro, S. Kyushin, H. Matsumoto, N. Mizorogi, S. Nagase, *Angew. Chem. Int. Ed.* **2005**, *44*, 7567 – 7570; c) V. Lavallo, Y. Canac, C. Präsang, B. Donnadiou, G. Bertrand, *Angew. Chem. Int. Ed.* **2005**, *44*, 5705 – 5709; d) D. Martin, A. Baceiredo, H. Gornitzka, W. W. Schoeller, G. Bertrand, *Angew. Chem. Int. Ed.* **2005**, *44*, 1700 – 1703.
- [4] a) E. Despagnet-Ayoub, R. H. Grubbs, *J. Am. Chem. Soc.* **2004**, *126*, 10198 – 10199; b) E. Despagnet-Ayoub, R. H. Grubbs, *Organometallics* **2005**, *24*, 338 – 340; c) K. E. Krahulic, G. D. Enright, M. Parvez, R. Roesler, *J. Am. Chem. Soc.* **2005**, *127*, 4142 – 4143; d) C. Präsang, B. Donnadiou, G. Bertrand, *J. Am. Chem. Soc.* **2005**, *127*, 10182 – 10183.
- [5] a) R. Fränkel, C. Birg, U. Kernbach, T. Habereeder, H. Nöth, W. P. Fehlhammer, *Angew. Chem. Int. Ed.* **2001**, *40*, 1907 – 1910;

- b) A. Wacker, C. G. Yan, G. Kaltenpoth, A. Ginsberg, A. M. Arif, R. D. Ernst, H. Pritzkow, W. Siebert, *J. Organomet. Chem.* **2002**, *641*, 195 – 202; c) A. Weiss, H. Pritzkow, W. Siebert, *Eur. J. Inorg. Chem.* **2002**, 1607 – 1614; d) A. Wacker, H. Pritzkow, W. Siebert, *Eur. J. Inorg. Chem.* **1998**, 843 – 849; e) U. Kernbach, M. Ramm, P. Luger, W. P. Fehlhammer, *Angew. Chem., Int. Ed. Engl.*, **1996**, *35*, 310 – 312; f) A. J. Arduengo III, D. Tapu, W. J. Marshall, *Angew. Chem. Int. Ed.* **2005**, *44*, 7240 – 7244.
- [6] a) T. Deforth, M. Kaschke, H. Stock, H. Pritzkow, W. Siebert, *Z. Naturforsch.* **1997**, *B52*, 823 – 830. b) M. Enders, B. Gangnus, R. Hettrich, Z. Magos-Martin, M. Stephan, H. Pritzkow, W. Siebert, U. Zenneck, *Chem. Ber.* **1993**, *126*, 2197 – 2203.
- [7] Crystallographic data. **1**: C₃₀H₄₅B₂F₃N₂O₃S, 193 K, 0.31 × 0.26 × 0.21 mm, monoclinic, space group *P2₁/c*, *a* = 19.2460(17), *b* = 10.6805(10), *c* = 15.7888(14) Å, $\alpha = \gamma = 90^\circ$, $\beta = 90.9307(15)^\circ$, *V* = 3245.01(5) Å³, *Z* = 4, $\rho_{\text{calcd}} = 1.212 \text{ g cm}^{-3}$, $2.30 \leq \theta \leq 25.86^\circ$, *R*₁ = 0.0507 (*I* > 2σ(*I*)), *wR*₂ = 0.1330 (all data); residual electron density: 0.280/-0.276 e Å⁻³. **2**: C₂₉H₄₄B₂N₂, 193 K, 0.65 × 0.57 × 0.23 mm, orthorhombic, space group *Pbca*, *a* = 15.6004(13), *b* = 21.0211(18), *c* = 17.1099(15) Å, $\alpha = \gamma = \beta = 90^\circ$, *V* = 5611.0(8) Å³, *Z* = 8, $\rho_{\text{calcd}} = 1.047 \text{ g cm}^{-3}$, $2.27 \leq \theta \leq 26.36^\circ$, *R*₁ = 0.0472 (*I* > 2σ(*I*)), *wR*₂ = 0.1415 (all data); residual electron density: 0.287/-0.224 e Å⁻³. **3**: C₃₇H₅₉B₂LiN₂O₂, 173 K, 0.16 × 0.12 × 0.08 mm, monoclinic, space group *C2/c*, *a* = 12.227(4), *b* = 18.309(6), *c* = 16.603(3) Å, $\alpha = \gamma = 90^\circ$, $\beta = 97.74(2)^\circ$, *V* = 3683.0(18) Å³, *Z* = 4, $\rho_{\text{calcd}} = 1.068 \text{ g cm}^{-3}$, $3.30 \leq \theta \leq 27.50^\circ$, *R*₁ = 0.070 (*I* > 2σ(*I*)), *wR*₂ = 0.209 (all data); residual electron density: 0.53/-0.37 e Å⁻³. MoK α radiation ($\lambda = 0.71073 \text{ \AA}$). Non-hydrogen atoms were refined anisotropically. Hydrogen atoms were included at geometrically calculated positions and refined using the riding model. CCDC-602682 - 602684 contain the supplementary crystallographic data for compounds **1** – **3**, respectively. These data can be obtained free of charge from The Cambridge Crystallographic Data Centre via www.ccdc.cam.ac.uk/data_request/cif.
- [8] a) I. V. Stankevich, E. G. Gal'pern, R. Seto, *Russ. J. Phys. Chem.* **2003**, *77*, 768 – 773; b) R. J. Doerksen, A. J. Thakkar, *J. Phys. Chem. A* **1998**, *102*, 4679 – 4686; c) T. Kar, D. E. Elmore, S. Scheiner, *J. Mol. Struct.* **1997**, *392*, 65 – 74; d) S. T. Massey, R. W. Zoellner, *Inorg. Chem.* **1991**, *30*, 1063 – 1066; e) R. Hoffmann, *J. Chem. Phys.* **1964**, *40*, 2474 – 2480.
- [9] a) R. Boese, A. H. Maulitz, P. Stellberg, *Chem. Ber.* **1994**, *127*, 1887 – 1889; b) C. A. Jaska, K. Temple, A. J. Lough, I. Manners, *J. Am. Chem. Soc.* **2003**, *125*, 9424 – 9434; c) T. Jäschke, M. Jansen, *Z. Anorg. Allg. Chem.* **2004**, *630*, 239 – 243.
- [10] P. Schreyer, P. Paetzold, R. Boese, *Chem. Ber.* **1988**, *121*, 195 – 205; C. J. Carmalt, W. Clegg, A. H. Cowley, F. J. Lawlor, T. B. Marder, N. C. Norman, C. R. Rice, O. J. Sandoval, A. J. Scott, *Polyhedron* **1997**, *16*, 2325 – 2328.
- [11] a) P. P. Power, *Appl. Organomet. Chem.* **2005**, *19*, 488 – 493; b) P. P. Power, *Chem. Commun.* **2003**, 2091 – 2101; c) J. A. C. Clyburne, N. McMullen, *Coord. Chem. Rev.* **2000**, *210*, 73 – 99; d) B. Twamley, S. T. Haubrich, P. P. Power, *Adv. Organomet. Chem.* **1999**, *44*, 1 – 65.
- [12] a) G. Boche, C. Hilf, K. Harms, M. Marsch, J. C. W. Lohrenz, *Angew. Chem. Int. Ed. Engl.* **1995**, *34*, 487 – 489; b) A. Wacker, H. Pritzkow, W. Siebert, *Eur. J. Inorg. Chem.* **1998**, 843 – 449; c) A. J. Arduengo, III, M. Tamm, J. C. Calabrese, F. Davidson, W. J. Marshall, *Chem. Lett.* **1999**, 1021 – 1022; d) R. Fränkel, C. Birg, U. Kernbach, T. Habereeder, H. Nöth, W. P. Fehlhammer, *Angew. Chem. Int. Ed.* **2001**, *40*, 1907 – 1910; e) P. L. Arnold, S. A. Mungur, A. J. Blake, C. Wilson, *Angew. Chem. Int. Ed.* **2003**, *42*, 5981 – 5984; f) S. A. Mungur, S. T. Liddle, C. Wilson, M. J. Sarsfield, P. L. Arnold, *Chem. Commun.* **2004**, 2738 – 2739; g) P. L. Arnold, M. Rodden, K. M. Davis, A. C. Scarisbrick, A. J. Blake, C. Wilson, *Chem. Commun.* **2004**, 1612 – 1613.
- [13] a) T. Maetzke, D. Seebach, *Helv. Chim. Acta* **1989**, *72*, 624 – 631; b) G. W. Rabe, R. D. Sommer, A. L. Rheingold, *Organometallics* **2000**, *19*, 5537 – 5540; c) G. W. Rabe, R. D. Sommer, A. L. Rheingold, *Organometallics* **2000**, *19*, 5537 – 5540; d) N. J. Hardman, B. Twamley, M. Stender, R. Baldwin, S. Hino, B. Schiemenz, S. M. Kauzlarich, P. P. Power, *J. Organomet. Chem.* **2002**, *643* – *644*, 461 – 467; e) G. E. Wegner, R. J. F. Berger, A. Schier, H. Schmidbaur, *Z. Naturforsch.* **2000**, *55B*, 995 – 999; f) R. C. Crittendon, B. C. Beck, J. Su, X.-W. Li, G. H. Robinson, *Organometallics* **1999**, *18*, 156 – 160. f) G. S. Girolami, M. E. Riehl, K. S. Suslick, S. R. Wilson, *Organometallics* **1992**, *11*, 3907 – 3910; g) M. M. Olmstead, P. P. Power, *J. Organomet. Chem.* **1991**, *408*, 1 – 6.
- [14] a) P. L. Arnold, M. Rodden, C. Wilson, *Chem. Commun.* **2005**, 1743 – 1745; g) P. L. Arnold, S. T. Liddle, *Organometallics* **2006**, *25*, 1485 – 1491.
- [15] For an account of ELF analysis in quantifying the σ -donating properties of NHCs, see: G. Frison, A. Sevin, *J. Phys. Chem. A* **1999**, *103*, 10998 – 11003.
- [16] a) S. Dapprich, G. Frenking, *J. Phys. Chem.* **1995**, *99*, 9352 – 9362; b) C. Boehme, G. Frenking, *Organometallics* **1998**, *17*, 5801 – 5809; c) G. Frenking, M. SolT, S. F. Vyboishchikov, *J. Organomet. Chem.* **2005**, *690*, 6178 – 6204.