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Conformations and Energetics of Sulfur and Selenium Diimides

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Abstract

The geometries and energetics of different conformations of sulfur and selenium diimides $E(NR)_2$ ($E = S, Se$; $R = H, Me, ^tBu, C_6H_3Me_2-2,6, SiMe_3$) have been studied by using various *ab initio* and DFT molecular orbital techniques. The *syn,syn* conformation is found to be most stable for parent $E(NH)_2$ but in general the preferred molecular conformation for substituted chalcogen diimides is *syn,anti*. In the case of $E(NH)_2$ the present calculations further confirm that *syn,syn* and *syn,anti* conformations lie energetically close to each other. From the three different theoretical methods used, B3PW91/6-31G* proved to be the most suitable method for predicting the geometries of chalcogen diimides. The optimized geometrical parameters are in a good agreement with all available experimental data. While qualitative energy ordering of the different conformations is independent of the level of theory, the quantitative energy differences are dependent on the method used. The performance and reliability of higher-level *ab initio* calculations and DFT methods using large basis sets were tested and compared with experimental information where available. All of the higher-level *ab initio* methods give very similar results but the use of large basis sets with the B3PW91 method does not increase the reliability of the results. The combination of CCSD(T)/cc-pVDZ with the B3PW91/6-31G* optimized geometries is found to be the method of choice to study energetic properties of chalcogen diimides.

Introduction

The structure and isomerism of sulfur diimides $S(NR)_2$ have attracted considerable interest in view of their utility as ligands in transition metal complexes¹⁻³ and as reagents in organic syntheses.⁴ $S(NSiMe_3)_2$, in particular, is a convenient source for the NSN fragment in a variety of cyclic and acyclic main-group compounds.⁵

Three different conformations are possible for chalcogen diimides, as shown in Figure 1. Variable temperature NMR spectroscopic studies have indicated that the *syn,anti* conformation of $S(NR)_2$ is most stable in solution in the case of small organic groups ($R = Me, ^tBu$).⁷ With bulkier organic substituents, small amounts of the *anti,anti* conformation was inferred to be in equilibrium with the *syn,anti* conformation.⁸⁻¹⁰ Recently, however, the existence of *syn,syn* conformations has been reported in the solutions of $S(NC_6H_2Br_3-2,4,6)_2$, $S(NC_6H_3Me_2-2,6)_2$, and $S(NC_6F_5)_2$.^{11,12} Structural determinations of a variety of $S(NR)_2$ molecules by X-ray diffraction in the solid state and electron diffraction in the gas phase have shown the presence of *syn,anti* and *syn,syn* conformations depending on the organic substituent R .¹¹⁻²³ There are a few semiempirical,^{11,12,24,25} *ab initio*,^{26,27} and DFT²⁸ calculations on sulfur diimides that have established that while *syn,anti* and *syn,syn* conformations lie close to each other in energy, the *anti,anti* conformation lies at a higher relative energy.

In contrast to stable sulfur diimides, selenium diimides $Se(NR)_2$ are thermally unstable.²⁹⁻³² The only known X-ray structure is that of *N,N'*-diadamantyl selenium diimide, which expectedly shows a *syn,anti* conformation in the solid state.³³ It has also been inferred from the

¹H and ¹³C NMR spectra that Se(N^tBu)₂ exists in solution in the *syn,anti* conformation.³¹ The analogous tellurium diimides have been observed to undergo a facile [2+2] cycloaddition with the formation of thermally stable dimers RNTe(μ -NR)₂TeNR.³⁴⁻³⁶ A similar [2+2] cycloaddition has also been observed between Se(N^tBu)₂ and ^tBuNSeO.³⁷ Sandblom *et al.*²⁸ have reported a DFT study of the conformations and dimerization energies of E(NMe)₂ (E = S, Se, Te).³⁸ These calculations have verified the facile [2+2] cycloaddition of Te(NMe)₂. The cycloaddition of methyl sulfur diimide was found to be strongly endothermic and that of methyl selenium diimide was concluded to be approximately thermochemically neutral.

The present *ab initio* and DFT study was undertaken in order to compare the geometries and energetics of a number of sulfur and selenium diimides at different levels of theory. The scope is to establish the level of reliability in the prediction of geometries and relative energies of different conformations of the monomeric diimides and establish the relevant structural trends. Comparison to experimental information is carried out where possible.

Computational Methods

All calculations were carried out with the *Gaussian* 98 program.³⁹ Geometries of all different conformations of E(NR)₂ (E = S, Se; R = H, Me, ^tBu, C₆H₃Me₂-2,6, SiMe₃) were fully optimized at the RHF,⁴⁰ MP2,⁴¹ and B3PW91^{42,43} levels of theory. The standard 6-31G* basis set was used as implemented in *Gaussian* 98. The hybrid B3PW91 functional was chosen, since it has been shown to perform well for molecules containing third-row elements.⁴⁴ A standard pruned (75,302) grid was used for B3PW91 optimizations but a denser pruned (99, 590) grid was chosen

for the single point energy calculations in order to minimize inaccuracies arising from the use of numerical integration. The fundamental frequencies were calculated at the RHF and DFT levels of theory in order to assess the nature of stationary points and estimate the zero-point energy (ZPE) corrections.

Two different high-level methods were used to calculate the relative energies of different conformations more accurately. Single point calculations for all optimized geometries were performed at the coupled-cluster⁴⁵ level of theory using Dunning's correlation consistent cc-pVDZ basis set as implemented in *Gaussian* 98. In addition, single point energies were also calculated with the B3PW91 method using basis sets of increasing size on B3PW91/6-31G* optimized geometries. Two Pople-type basis sets (6-31G* and 6-311G**) and two correlation consistent basis sets (cc-pVDZ and cc-pVTZ) were used as internally available in *Gaussian* 98.

Results and Discussion

Geometries: The optimized geometries of sulfur diimides are given in Table 1 and those of selenium diimides in Table 2. In the case of each sulfur diimide studied in this work, the calculated values can be compared with molecular parameters that have been determined by X-ray or electron diffraction (see Table 1). The comparison of calculated and experimental geometries indicates that the bond parameters can be predicted reasonably well at all levels of theory.⁴⁶ The RHF/6-31G* calculations underestimate the S=N bond lengths by 0.01-0.03 Å, whereas at the MP2/6-31G* level they are overestimated by approximately 0.06 Å. The S=N bond lengths yielded by the B3PW91/6-31G* calculations seem to be in best overall agreement

with the experimental values. The experimental H-N, C-N, and Si-N bond lengths are accurately reproduced at every level of theory. Calculated RMS deviations between the experimental and calculated bond lengths are 0.022, 0.041, and 0.018 Å at the RHF/6-31G*, MP2/6-31G*, and B3PW91/6-31iG* levels of theory, respectively.

With the exception of S(NH₂)₂ and S(NC₆H₃Me₂-2,6)₂ that appear as *syn,syn* conformers, the experimental molecular structures of all other sulfur diimides shown in Table 1 exhibit *syn,anti* conformations.⁴⁷ It can also be noted that the two NSN fragments in S(NSNSiMe₃)₂⁴⁹ and Se(NSNSiMe₃)₂⁶ also lie in *syn,anti* conformations. As seen in Table 1, in most cases the S=N bond of the *syn*-S=N-R fragment is shorter than that corresponding to the *anti*-S=N-R orientation. This can be rationalized in terms of the bonding description discussed by Sandblom *et al.*²⁸ While the S=N bond lengths generally seem to be independent of the identity of the organic group, those in S(NSiMe₃)₂ are shorter in every conformation. This can be explained by charge delocalization due to the hyperconjugative interaction between the nitrogen 2p lone pair and the SiMe₃ fragment. Results from the preliminary ELF^{50,51} calculations for S(N^tBu)₂ and S(NSiMe₃)₂ seem to support this conclusion, as they show that the electron population in lone pairs is much smaller in S(NSiMe₃)₂ than in S(N^tBu)₂.

Of the selenium diimides, the X-ray structure is known only for *N,N'*-diadamantyl selenium diimide Se(NC₁₀H₁₅)₂,³³ which shows a *syn,anti* conformation consistent with most sulfur diimides. The experimental Se=N bond lengths are 1.679(8) Å (*syn*-Se=N-R) and 1.732(7) Å (*anti*-Se=N-R) displaying a similar trend as *syn,anti* sulfur diimides. Other estimates for the Se=N double bond length can be obtained from the X-ray structures of ^tBuNSe(μ-N^tBu)₂SO₂ and

$t\text{BuNSe}(\mu\text{-N}^t\text{Bu})_2\text{SeO}$.³⁷ The exocyclic Se=N bonds in these two compounds show lengths of 1.665(2) and 1.687(4) Å, respectively. The optimized *syn*-Se=N bond lengths for the different *syn,anti* conformations of $\text{Se}(\text{NR})_2$ lie in the range 1.633-1.650, 1.708-1.743, and 1.678-1.702 Å at RHF/6-31G*, MP2/6-31G*, and B3PW91/6-31G* levels of theory, respectively (see Table 2). The corresponding predictions for the *anti*-Se=N bond lengths are 1.646-1.663, 1.732-1.757, and 1.699-1.722 Å. The calculated Se=N bonds in the *syn*-Se=N-R fragments are again predicted to be systematically shorter than those in the *anti*-Se=N-R fragments thus reproducing the experimental trend. Consistent with the bonding arrangement in sulfur diimides, both Se=N bonds of $\text{Se}(\text{NSiMe}_3)_2$ are shorter than the corresponding bonds in any other selenium diimide.

Although the calculations reproduce the bond lengths relatively accurately, the calculated bond angles show a larger deviation from the experimental values. It can be seen from Table 1 that the NSN bond angles are in good agreement with experimental values (deviations *ca.* $\pm 1^\circ$) at each level of theory, but the RHF/6-31G* and MP2/6-31G* optimized SNA ($\text{A} = \text{H}, \text{C}, \text{Si}$) angles deviate from the experimental values. The RHF/6-31G* calculations overestimate the SNA angles, whereas MP2/6-31G* calculations lead to a small underestimation. By contrast, the B3PW91/6-31G* calculations seem to provide a good prediction also for the experimental SNA bond angles. The RMS deviations between the experimental and calculated bond angles are 3.7, 3.4 and 2.3° at the RHF/6-31G*, MP2/6-31G*, and B3PW91/6-31G* levels of theory, respectively. The optimized NSeN bond angles span a range of 107.2-132.1, 102.8-139.5, and 104.0-136.7° and the SeNA bond angles 107.1-153.5°, 103.1-143.3°, and 104.1-148.8° at RHF/6-31G*, MP2/6-31G*, and B3PW91/6-31G* levels, respectively. The experimental NSeN angle in $\text{Se}(\text{NC}_{10}\text{H}_{15})_2$ is 113.0(6)°, and the two SeNC bond angles are 125.0(6)° and 117.6(5)°.³³

Although the bond lengths in the different conformers of the molecules with the same composition do not show significant variation, both the NEN ($E = S$ or Se) and ENA bond angles are larger in the *syn,syn* conformations than in the *anti,anti* conformations for a given group R (see Figure 2). The NEN angles in the *syn,anti* conformation expectedly fall between the values for the *syn,syn* and *anti,anti* conformations. The *syn*-ENA angle and the *anti*-ENA angle in the *syn,anti* conformation mimic the respective angles in the two symmetric conformations. It can clearly be seen from Figure 2 that the bond angles expectedly get larger, as the R-group becomes bulkier. It can also be noted in Tables 1 and 2 that the ENA angles in selenium diimides are smaller than those in the corresponding sulfur diimides. This indicates that the electronic factors in the bonding arrangement in sulfur and selenium diimides are relatively similar and the molecular geometry depends largely on the steric factors.

The present calculations seem to yield a better agreement between the calculated and experimental bond parameters than the previous theoretical approaches. Semi-empirical CNDO and MNDO calculations^{11,12,24,25} should only be considered as qualitative and the early *ab initio* calculations^{26,27} have suffered from the use of small, inflexible minimal STO-3G* basis sets that limit the accuracy in the predictions of both molecular geometry and energetics. Sandblom *et al.*²⁸ have recently reported a DFT study of the conformations and dimerization of $E(NMe)_2$ ($E = S$, Se , Te). These calculations agree well with the present B3PW91/6-31G* results in the case of *N,N'*-dimethyl sulfur diimide with somewhat more pronounced differences in the case of *N,N'*-dimethyl selenium diimide. The DFT prediction of Sandblom *et al.*²⁸ yields Se=N bond lengths that are *ca.* 0.070 Å longer and bond angles that are 2-4° smaller than the corresponding

B3PW91/6-31G* values. It seems that generally the B3PW91/6-31G* geometries are in better agreement with experiment than those yielded by the earlier DFT calculations.³⁸

Energetics: The total energies and ZPE corrections calculated for sulfur and selenium diimides at RHF/6-31G*, MP2/6-31G*, and B3PW91/6-31G* levels of theory are presented in Supporting Information (Table S1). The relative energies of the different conformations are depicted in Figure 3.

It can be seen from Figure 3 that, with the exception of E(NH)₂ for which the *syn,syn* conformation is most stable, the *syn,anti* conformation is predicted to be most stable for substituted sulfur and selenium diimides regardless of the level of theory.⁵² In the case of parent E(NH)₂, our calculations further confirm the earlier deductions that *syn,syn* and *syn,anti* conformations of the parent molecule lie close to each other in energy. These results are expectedly in agreement with the previous experimental¹¹⁻²³ and computational^{11,12,24-28} findings. The qualitative energy ordering of the different conformations is found to be independent of the level of theory. Results in Figure 3 also show that the *anti,anti* conformer is more unstable than could be expected on the basis of steric factors alone. Sandblom *et al.*²⁸ have recently suggested that the antibonding interaction between the in-plane nitrogen lone pair and the p orbital on sulfur destabilizes the sterically favorable *anti,anti* conformation.

Quantitative differences in the relative energies can, however, be observed between the calculations carried out at different levels of theory. MP2/6-31G* generally destabilizes the *anti,anti* conformations and stabilizes the *syn,syn* conformations by 5-10 kJ mol⁻¹ compared to

RHF/6-31G* calculations. The B3PW91/6-31G* results fall between those from RHF/6-31G* and MP2/6-31G* calculations. The B3PW91/6-31G* energy difference between the *syn,syn* and *syn,anti* conformations resembles that extracted from the MP2/6-31G* calculations, while the energy difference between the *syn,anti* and *anti,anti* conformations lies closer to that from the RHF/6-31G* calculations than from MP2/6-31G* results.

Since no accurate experimental energetics data are available for chalcogen diimides, the reliability of the results lies solely on the grounds of theory. This was evaluated by high-level CCSD/cc-pVDZ and CCSD(T)/cc-pVDZ calculations. It was not, however, possible to carry out full geometry optimizations at coupled cluster levels of theory. Therefore single point calculations utilizing optimized geometries that were calculated at RHF/6-31G*, MP2/6-31G*, and B3PW91/6-31G* levels of theory (see Tables 1 and 2) were carried out for all sulfur and selenium diimides. The total energies yielded by these calculations are presented in Supporting Information (Table S2). Relative energies are presented at Table 3. The CCSD(T)/cc-pVDZ//B3PW91/6-31G* relative energies are shown in Figure 4. The CCSD(T)/cc-pVDZ single point calculations were chosen to utilize B3PW91/6-31G* optimized geometries, since they were in closest agreement with experimental geometries, where available. Furthermore, in most cases, the B3PW91/6-31G* geometries yielded lowest total energies at the CCSD/cc-pVDZ level and could therefore be considered as best geometries at coupled cluster levels of theory. The energetics of the three conformations of the different chalcogen diimides will therefore be discussed below at CCSD(T)/cc-pVDZ//B3PW91/6-31G* level of theory.

It can be seen from Figures 3 and 4 that the terminal R-group in E(NR)₂ seems to have a greater influence on the relative energies of the conformations than the identity of the chalcogen atom. The relative energies for molecules with same substituents but different chalcogen atoms deviate only in the case of *tert*-butyl diimides, whereas the relative energies for different substituents vary substantially. For example, *syn,syn*-S(NMe)₂ lies 12.2 kJ mol⁻¹ and *syn,syn*-S(N^tBu)₂ lies 46.8 kJ mol⁻¹ above their respective *syn,anti* conformations, and the *syn,syn* conformations of Se(NMe)₂ and Se(N^tBu)₂ lie at 8.3 and 29.9 kJ mol⁻¹ above the most stable *syn,anti*-conformations, respectively (see Figure 4). In the cases of S(NSiMe₃)₂ and Se(NSiMe₃)₂ both the *syn,syn* and *anti,anti* conformations lie at almost identical energies of 23.4-25.4 kJ mol⁻¹ above the *syn,anti* conformation.

The CCSD(T)/cc-pVDZ//B3PW91/6-31G* results also indicate that the atomization energies are systematically *ca.* 90 kJ mol⁻¹ smaller for selenium diimides than for corresponding sulfur diimides.⁵³ This is consistent with the experimentally observed thermal instability of selenium diimides.²⁹⁻³²

The comparison of the CCSD(T)/cc-pVDZ//B3PW91/6-31G* results with existing computational work indicates that the previous *ab initio* calculations^{26,27} have predicted that the *syn,syn* conformations lie lower and the *anti,anti* conformations lie higher in energy with respect to the *syn,anti* conformations than what is predicted by the present calculations. The present energy ordering of the three conformations of S(NMe)₂ also differs from that reported by Sandblom *et al.*²⁸ This divergence can partially be explained by the different methods used in the calculations.^{38,54} We also note, however, that the geometry optimizations of the *syn,syn*

conformations by Sandblom *et al.*²⁸ are carried out by constraining the molecule in C_{2v} symmetry. The present RHF/6-31G* and B3PW91/6-31G* level optimizations indicate that the C_{2v} structure is in fact a first order transition state with respect to the rotation of the methyl groups. Local minimum for the *syn,syn* conformation was found with C_2 symmetry, as verified by frequency calculations.

To test the conventional *ab initio* methods against the chosen B3PW91 functional and to evaluate the reliability of DFT results, B3PW91 single point energies were calculated with more flexible basis sets on B3PW91/6-31G* geometries. The calculated relative energies of different chalcogen diimides are shown in Table 4. These values can be compared to the CCSD(T)/cc-pVDZ//B3PW91/6-31G* relative energies that represent the highest level of reliability in the present *ab initio* calculations (see Table 3). It can clearly be seen from Table 4 that as the size of the basis set is increased, the B3PW91 relative energies show variation, the mutual differences between the basis sets often being above 10 kJ mol⁻¹. In the case of the most flexible basis set, cc-pVTZ, the B3PW91 calculations in fact yield relative energies that are furthest away from those obtained by the CCSD(T)/cc-pVDZ//B3PW91/6-31G* level calculations. The smallest Pople-type basis sets seem to reproduce best the high level *ab initio* relative energies in the case of all sulfur and selenium diimides.

By contrast, it is well-known that the relative energies calculated using *ab initio* methods converge and become more accurate, when the level of sophistication is increased (See Tables S1 and S2 in Supporting Information). In most cases, the relative energies calculated at MP2/cc-pVDZ level of theory can already be considered reasonably reliable.

These findings seem to contradict the one-dimensionality of the DFT methods *i.e.* the increase in the size of a basis set should result in a better description of the energy.⁵⁶ It has been stated earlier that "...it is rather plausible that the many honors density functional theory has earned in this field in the past are due to massive error compensation effects arising from the use of a small basis sets".⁵⁷ Large basis sets should therefore not be kept as a definite measure of the quality of DFT results. This has not been explored to its full extent in computational chemistry. It can be very tempting in individual situations to select a basis set that will best reproduce the observed property and hold the method superior, though the agreement between calculated and experimental property might only be fortuitous. Current hybrid density functional methods are fast and valuable means to study structures and chemical properties of molecules. However, as there seems to be no systematic way to improve the accuracy of the DFT calculations, the comparison of the calculated parameters with experimental or sufficiently high-level *ab initio* data is necessary to evaluate the level of reliability of the results.

Conclusion

The geometries and energetics of the different conformations of a number of sulfur and selenium diimides E(NR)₂ (E = S, Se; R = H, Me, ¹Bu, C₆H₃Me₂-2,6, SiMe₃) have been studied by using *ab initio* and DFT molecular orbital techniques. With the exception of parent E(NH)₂, for which the *syn,syn* conformation is most stable, the *syn,anti* conformation is predicted to be most stable for substituted sulfur and selenium diimides regardless of the level of theory. In the case of

parent E(NH)₂, the present calculations further confirm that *syn,syn* and *syn,anti* conformations lie close to each other.

The geometry optimizations were carried out at RHF/6-31G*, MP2/6-31G*, and B3PW91/6-31G* levels of theory. B3PW91/6-31G* proved to be the most suitable method for reproducing the experimental geometries of chalcogen diimides. It has a speed advantage compared to *ab initio* methods and provides results of at least the same accuracy as computationally much more demanding MP2/6-31G*.

The total and relative energies of the different conformations of S(NR)₂ and Se(NR)₂ molecules have been calculated at RHF/6-31G*, MP2/6-31G*, and B3PW91/6-31G* levels of theory and compared with those yielded by higher-level CCSD(T) and B3PW91 calculations. The qualitative energy ordering of the different conformations is found to be independent of the level of theory. By contrast, the quantitative energy differences are found to be dependent on the method. Whereas *ab initio* methods yield increasingly accurate relative energies with increasing level of theory, B3PW91 results show divergence, as the size of the basis set is increased. The B3PW91 involving small Pople-type basis sets seem to reproduce best the energetics of chalcogen diimides from high-level *ab initio* calculations. Since the B3PW91/6-31G* calculations predict well the optimum geometries, reliable relative energies of the chalcogen diimide conformers can conveniently be computed at CCSD(T)/cc-pVDZ//B3PW91/6-31G* level of theory.

It is well established experimentally that, whereas sulfur diimides $S(NR)_2$ exist as monomeric species in solution, the corresponding tellurium diimides $Te(NR)_2$ undergo a facile [2+2] cycloaddition reaction to form dimeric $RNTe(\mu-NR)_2TeNR$.³⁴⁻³⁶ It is not quite clear whether selenium diimides are preferentially monomeric or dimeric. The NMR spectroscopic evidence of selenium diimides in solution³¹ as well as the X-ray structure of *N,N'*-diadamantyl selenium diimide³³ indicate a preference for monomeric structures, but [2+2] cycloaddition reactions have been inferred for the dimerization of tBuNSeO to $OSe(\mu-^tBuN)_2SeO$ ³³ and for the reaction of tBuNSeO and $Se(N^tBu)_2$ that affords $^tBuNSe(N^tBu)_2SeO$.³⁷ The DFT calculations of Sandblom *et al.*²⁸ have indicated that the cycloaddition reactions of sulfur diimides are strongly endothermic, those of selenium diimides are approximately thermally neutral, and those involving tellurium diimides are strongly exothermic. Since the present calculations demonstrate that the relative energies of various chalcogen diimide monomers are dependent on the method and level of theory, it is conceivable that the energetics of cycloaddition is also affected by the computational approach. Calculations to explore the cycloaddition process at different levels of theory are currently in progress.

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Supporting Information Available: Total energies of sulfur and selenium diimides calculated at different levels of theory (Tables S1 and S2). This material is available free of charge via the Internet at <http://pubs.acs.org>.

References:

- (1) Vrieze, K.; van Koten, G. *Recl. Trav. Chim. Pays-Bas* **1980**, *99*, 145.
- (2) Chivers, T.; Hilts, R.W. *Coord. Chem. Rev.* **1994**, *137*, 201.
- (3) Hill, A. *Organomet. Chem. Rev.* **1994**, *36*, 159.
- (4) Bussas, R.; Kresze, G.; Münsterer, H.; Schwöbel, A. *Sulphur Rep.* **1983**, *2*, 215.
- (5) For recent reviews, see Ref. 6.
- (6) Konu, J.; Maaninen, A.; Paananen, K.; Ingman, P.; Laitinen, R.S.; Chivers, T.; Valkonen, J. *Inorg. Chem.* **2002**, *41*, 1430, and references therein.
- (7) Grunwell, J.R.; Hoyng, C.F.; Rieck, J.A. *Tetrahedron Lett.* **1973**, *26*, 2421.
- (8) Kuyper, J.; Vrieze, K. *J. Organomet. Chem.* **1974**, *74*, 289.
- (9) Kuyper, J.; Vrieze, K. *J. Organomet. Chem.* **1975**, *86*, 127.
- (10) Sicinska, W.; Stefaniak, L.; Witanowski, M.; Webb, G.A. *J. Mol. Struct.* **1987**, *158*, 57.
- (11) Bagryanskaya, I.Yu.; Gatilov, Y.; Shakirov, M.M.; Zibarev, A.V. *Mendeleev. Commun.* **1994**, *136*.
- (12) Bagryanskaya, I.Yu.; Gatilov, Y.; Shakirov, M.M.; Zibarev, A.V. *Mendeleev. Commun.* **1994**, *167*.
- (13) Leandri, G.; Busetti, V.; Valle, G.; Mammi, M. *Chem. Commun.* **1970**, 413.
- (14) Leitch, J.; Nyburg, S.C.; Armitage, D.A.; Clark, M.J. *J. Cryst. Mol. Struct.* **1973**, *3*, 337.
- (15) Olsen, F.P.; Barrick, J.C. *Inorg. Chem.* **1973**, *12*, 1353.
- (16) Kuyper, J.; Isselmann, P.H.; Mijlhoff, F.C.; Spelbos, A.; Renes, G. *J. Mol. Struct.* **1975**, *29*, 247.
- (17) Busetti, V. *Acta Crystallogr.* **1982**, *B38*, 665.
- (18) Suenram, R.D.; Lovas, F.J.; Stevens, W.J. *J. Mol. Spect.* **1985**, *112*, 482.

- (19) Anderson, D.G.; Robertson, H.E.; Rankin, D.W.H.; Woollins, J.D. *J. Chem. Soc., Dalton Trans.* **1989**, 859.
- (20) Herberhold, M.; Gerstmann, S.; Wrackmeyer, B.; Borrmann, H. *J. Chem. Soc., Dalton Trans.* **1994**, 5, 633.
- (21) Herberhold, M.; Gerstmann, S.; Milius, W.; Wrackmeyer, B.; Borrmann, H. *Phosphorus, Sulfur Silicon Relat. Elem.* **1996**, 112, 261.
- (22) Lork, E.; Mews, R.; Shakirov, M.M.; Watson, P.G.; Zibarev, A.V. *Eur. J. Inorg. Chem.* **2001**, 2123.
- (23) Zibarev, A.V.; Gatalov, Y.V.; Bagryanskaya, I.Yu.; Konchenko, S.N. *Polyhedron* **1992**, 11, 2787.
- (24) Grunwell, J. R.; Danison W.C. *Tetrahedron* **1971**, 27, 5315.
- (25) Issa, Y.; Hossain, F. B. S.; Davood, N. S. *J. Mol. Struct.* **1996**, 364, 79.
- (26) Raghavachari, K.; Haddon, R.C. *J. Phys. Chem.* **1983**, 87, 1308.
- (27) Millefiori, S.; Millefiori, A.; Granozzi, G. *Inorg. Chim. Acta* **1984**, 90, 105.
- (28) Sandblom, N.; Ziegler, T.; Chivers, T. *Inorg. Chem.* **1998**, 37, 354.
- (29) Herberhold, M.; Jellen, W. *Z. Naturforsch., Teil B* **1986**, 41b, 144.
- (30) Sharpless, K. B.; Hori, T.; Truesdale, L. K.; Dietrich, C. O. *J. Am. Chem. Soc.* **1976**, 98, 269.
- (31) Wrackmeyer, B.; Distler, B.; Gerstmann, S.; Herberhold, M. *Z. Naturforsch.* **1993**, 48b, 1307.
- (32) Fockenberg, F.; Haas, A. *Z. Naturforsch. Sect. B* **1986**, 41b, 413.
- (33) Maaninen, T.; Laitinen, R.; Chivers, T. *Chem. Commun.* **2002**, 1812.
- (34) Chivers, T.; Gao, X.; Parvez, M. *J. Chem. Soc., Chem. Commun.* **1994**, 2149.

- (35) Chivers, T.; Gao, X.; Parvez, M. *J. Am. Chem. Soc.* **1995**, *117*, 2359.
- (36) Chivers, T.; Gao, X.; Parvez, M. *Inorg. Chem.* **1996**, *35*, 9.
- (37) Maaninen, T.; Chivers, T.; Laitinen, R.S.; Schatte, G.; Nissinen, M. *Inorg. Chem.* **2000**, *39*, 5341.
- (38) These calculations were based on DFT within the local density approximation. In addition, Becke's nonlocal exchange correction and Perdew's inhomogeneous gradient corrections for correlation were used. A double-zeta STO basis sets were used for the ns and np shells of the main group elements. The basis sets were augmented by a single 3d STO polarization function (2p STO for hydrogen). For full details of the calculation, see Ref. 28 and references therein.
- (39) Frisch, M. J.; Trucks; G. W.; Schlegel, H. B.; Scuseria, G. E.; Robb, M. A.; Cheeseman, J. R.; Zakrzewski, V. G.; Montgomery, J. A Jr.; Stratmann, R. E.; Burant, J. C.; Dapprich, S.; Millam, J. M.; Daniels, A. D.; Kudin, K. N.; Strain, M. C.; Farkas, O.; Tomasi, J.; Barone, V.; Cossi, M.; Cammi, R.; Mennucci, B.; Pomelli, C.; Adamo, C.; Clifford, S.; Ochterski, J.; Petersson, G. A.; Ayala, P. Y.; Cui, Q.; Morokuma, K.; Salvador, P.; Dannenberg, J. J.; Malick, D. K.; Rabuck, A. D.; Raghavachari, K.; Foresman, J. B.; Cioslowski, J.; Ortiz, J. V.; Baboul, A. G.; Stefanov, B. B.; Liu, G.; Liashenko, A.; Piskorz, P.; Komaromi, I.; Gomperts, R.; Martin, R. L.; Fox, D. J.; Keith, D. J.; Al-Laham, M. A.; Peng, C. Y.; Nanayakkara, A.; Challacombe, M.; Gill, P. M. W.; Johnson, B.; Chen, B.; Wong, M. W.; Andres, J. L.; Gonzalez, C.; Head-Gordon, M.; Replogle, E. S.; Pople, J. A., *Gaussian 98* (Revision A.11), Gaussian, Inc., Pittsburgh PA, 2001.
- (40) Roothaan, C. C. J. *Rev. Mod. Phys.* **1951**, *23*, 69.
- (41) Møller, C.; Plesset, M. S. *Phys. Rev.* **1934**, *46*, 618.

- (42) Becke, A. D. *J. Chem. Phys.* **1993**, *98*, 5648.
- (43) Perdew, J. P.; Wang, Y. *Phys. Rev. B* **1992**, *45*, 13244.
- (44) Redfern, P. C.; Blaudeau, J. P.; Curtiss, L. A. *J. Phys. Chem. A* **1997**, *101*, 8701.
- (45) Bartlett, R. J. *J. Phys. Chem.* **1989**, *93*, 1697, and references therein.
- (46) It should be noted, however, that the two phenyl rings of *anti,anti* 2,6-dimethylphenylsulfur diimide are predicted to be coplanar with the NSN fragment at B3PW91/6-31G* level of theory and almost perpendicular to the NSN plane at MP2/6-31G* level of theory. While there is no experimental information for this conformation, the known molecular structures of several sulfur diimides containing an aromatic substituent bound to the nitrogen atoms of the NSN fragment demonstrate that the plane of the ring can indeed assume several different orientations from coplanarity to being perpendicular to the plane of the NSN fragment,^{11-15,17} therefore implying low energy barriers to internal rotation.
- (47) There have been conflicting reports about the conformation of S(NSiMe₃)₂. The assignment of IR and Raman spectra could be made in terms of a distorted *anti,anti* conformation.⁴⁸ The electron diffraction data of a gaseous sample were interpreted in terms of a *syn,syn* conformation,¹⁹ while the solid state structure could unambiguously be determined to show a *syn,anti* conformer by X-ray crystallography.²⁰
- (48) Herbrechtsmeier, A.; Schnebel, F.-M.; Glemser, O. *J. Mol. Struct.* **1978**, *50*, 43.
- (49) Bagryanskaya, I.Yu.; Gatilov, Y.V.; Shakirov, M.M.; Zibarev, A.V., *J. Struct. Chem.* **1996**, *37*, 318.
- (50) Noury, S.; Krokidis, X.; Fuster F.; Silvi, B. TopMoD program package, Universite Pierre et Marie Curie, 1997.

- (51) Malcolm, N. O. J.; Gillespie, R. J.; Popelier, P. L. A. *J. Chem. Soc., Dalton Trans.* **2002**, 17, 3333.
- (52) The second exception is S(NC₆H₃Me₂-2,6)₂ for which the *syn,syn* conformation is also most stable in agreement with the experimental observation.¹¹ The *syn,anti* conformation lies 2.8 kJ mol⁻¹ and the *anti,anti* conformation 18.9 kJ mol⁻¹ above the *syn,syn* conformation at B3PW91/6-31G* level of theory (see Table S1 in Supporting Information for total energies and ZPE corrections).
- (53) CCSD(T)/cc-pVDZ total energies (a.u.) for free H, C, N, Si, S and Se atoms are -0.49928, -37.76038, -54.47850, -288.91509, -397.60055 and -2399.89854, respectively.
- (54) It has recently been concluded that the hybrid functionals generally perform better than LDA and GGA approaches (see Ref. 55, p. 124, and references therein).
- (55) Koch, W.; Holthausen, M.C. *A Chemist's Guide to Density Functional Theory*, VCH: Weinheim 2001.
- (56) Jensen, F. *Introduction to Computational Chemistry*, John Wiley & Sons: Chichester, 1999.
- (57) See Ref. 55, p. 246.

Table 1. The RHF/6-31G*, MP2/6-31G*, and B3PW91/6-31G* Optimized Geometries of Sulfur Diimides S(NR)₂ (R = H, Me, ^tBu, C₆H₃Me₂-2,6, SiMe₃; A = H, C, Si)

Molecule	Method	r(SN ₁)	r(SN ₂)	r(N ₁ A)	r(N ₂ A)	∠NSN	∠SN ₁ A	∠SN ₂ A
S(NH)₂								
(s,s)	RHF	1.507		1.006		122.2	117.8	
	MP2	1.569		1.030		124.2	115.2	
	B3PW91	1.549		1.025		124.1	116.2	
	Exptl. ^a	1.533(7)		1.025(15)		124.0(9)	115.4(8)	
(a,a)	RHF	1.507		1.004		112.7	108.8	
	MP2	1.573		1.028		108.8	105.6	
	B3PW91	1.552		1.024		110.0	106.1	
(s,a)	RHF	1.501	1.513	1.007	1.004	116.2	115.2	110.1
	MP2	1.560	1.578	1.032	1.029	114.7	111.9	107.3
	B3PW91	1.541	1.558	1.027	1.025	115.5	113.1	107.8
S(NMe)₂								
(s,s)	RHF	1.502		1.447		125.2	130.7	
	MP2	1.576		1.464		127.5	126.8	
	B3PW91	1.549		1.448		127.2	129.2	
(a,a)	RHF	1.505		1.453		113.2	116.8	
	MP2	1.581		1.471		109.6	112.5	
	B3PW91	1.555		1.456		110.5	114.2	
(s,a)	RHF	1.501	1.513	1.449	1.453	115.9	122.0	117.6
	MP2	1.573	1.592	1.463	1.471	113.5	116.9	113.5
	B3PW91	1.547	1.564	1.449	1.456	114.7	119.6	115.2
	Exptl. ^b	1.532(10)	1.532(10)	1.464(18)	1.464(18)	113.6(16)	124.3(12)	116.5(12)

Table 1. Continued

Molecule	Method	<i>r</i> (SN ₁)	<i>r</i> (SN ₂)	<i>r</i> (N ₁ A)	<i>r</i> (N ₂ A)	∠NSN	∠SN ₁ A	∠SN ₂ A
S(N^tBu)₂								
(s,s)	RHF	1.492		1.458		131.5	147.3	
	MP2	1.584		1.464		137.2	141.0	
	B3PW91	1.544		1.459		134.9	143.7	
(a,a)	RHF	1.504		1.470		112.0	121.9	
	MP2	1.586		1.482		107.7	117.0	
	B3PW91	1.555		1.473		109.1	118.7	
(s,a)	RHF	1.498	1.510	1.469	1.472	118.3	130.6	122.3
	MP2	1.573	1.598	1.476	1.483	116.7	125.5	117.7
	B3PW91	1.542	1.565	1.472	1.476	117.6	128.1	119.1
	Exptl. ^c	1.530(12)	1.545(11)	1.480(15)	1.492(12)	117.41(3)	127.95(5)	118.22(4)
S(NSiMe₃)₂								
(s,s)	RHF	1.482		1.739		123.3	155.8	
	MP2	1.549		1.768		126.4	147.4	
	B3PW91	1.525		1.752		125.1	152.0	
(a,a)	RHF	1.489		1.746		115.6	132.7	
	MP2	1.556		1.776		112.6	123.4	
	B3PW91	1.534		1.761		113.3	127.2	
(s,a)	RHF	1.488	1.499	1.765	1.750	118.2	134.9	131.6
	MP2	1.569	1.592	1.795	1.779	116.9	126.5	124.1
	B3PW91	1.529	1.545	1.778	1.765	117.9	132.2	126.6
	Exptl. ^d	1.5155(6)	1.5230(5)	1.7585(6)	1.7593(5)	117.77(3)	129.78(4)	123.29(3)

Table 1. Continued

Molecule	Method	<i>r</i> (SN ₁)	<i>r</i> (SN ₂)	<i>r</i> (N ₁ A)	<i>r</i> (N ₂ A)	∠NSN	∠SN ₁ A	∠SN ₂ A
S(NC₆H₃Me₂)₂								
(s,s)	RHF	1.499		1.409		127.4	138.0	
	MP2	1.602		1.408		124.1	124.4	
	B3PW91	1.554		1.397		128.9	136.3	
	Exptl. ^e	1.526(2)		1.423(3)		127.2(2)	131.8(2)	
(a,a)	RHF	1.507		1.427		111.8	117.4	
	MP2	1.591		1.436		108.7	110.8	
	B3PW91	1.573		1.379		104.3	128.6	
(s,a)	RHF	1.499	1.512	1.416	1.422	117.1	129.6	119.8
	MP2	1.580	1.598	1.416	1.433	115.2	123.7	112.9
	B3PW91	1.550	1.568	1.399	1.401	115.4	130.3	123.6

^a Ref. 18. ^b Ref. 16. ^c Ref. 21. ^d Ref. 20. ^e Ref. 11.

Table 2. The RHF/6-31G*, MP2/6-31G*, and B3PW91/6-31G* Optimized Geometries of Selenium Diimides $\text{Se}(\text{NR})_2$ ($\text{R} = \text{H}, \text{Me}, {}^t\text{Bu}, \text{SiMe}_3$; A = H, C, Si).

Molecule	Method	$r(\text{SeN}_1)$	$r(\text{SeN}_2)$	$r(\text{N}_1\text{A})$	$r(\text{N}_2\text{A})$	$\angle \text{NSEn}$	$\angle \text{SeN}_1\text{A}$	$\angle \text{SeN}_2\text{A}$
Se(NH)₂								
(s,s)	RHF	1.655		1.007		119.2	114.0	
	MP2	1.729		1.034		121.8	111.1	
	B3PW91	1.702		1.027		120.9	111.5	
(a,a)	RHF	1.655		1.006		108.2	107.1	
	MP2	1.733		1.034		104.3	103.1	
	B3PW91	1.704		1.027		105.3	104.1	
(s,a)	RHF	1.649	1.660	1.008	1.006	112.1	111.3	108.6
	MP2	1.721	1.737	1.037	1.034	110.9	107.7	104.9
	B3PW91	1.694	1.710	1.029	1.027	111.0	108.5	105.8
Se(NMe)₂								
(s,s)	RHF	1.651		1.446		123.4	128.8	
	MP2	1.741		1.465		126.4	124.8	
	B3PW91	1.705		1.447		125.6	126.6	
(a,a)	RHF	1.653		1.451		108.8	116.5	
	MP2	1.744		1.475		105.6	111.6	
	B3PW91	1.708		1.455		106.0	113.7	
(s,a)	RHF	1.650	1.660	1.449	1.451	111.8	119.8	117.6
	MP2	1.738	1.757	1.466	1.474	108.9	113.9	112.9
	B3PW91	1.702	1.717	1.449	1.455	109.9	116.8	115.0

Table 2. Continued

Molecule	Method	<i>r</i> (SeN ₁)	<i>r</i> (SeN ₂)	<i>r</i> (N ₁ A)	<i>r</i> (N ₂ A)	∠NSeN	∠SeN ₁ A	∠SeN ₂ A
Se(N^tBu)₂								
(s,s)	RHF	1.646		1.454		132.1	143.4	
	MP2	1.760		1.459		139.5	136.7	
	B3PW91	1.708		1.453		136.7	139.1	
(a,a)	RHF	1.653		1.467		107.2	120.5	
	MP2	1.753		1.485		102.8	114.4	
	B3PW91	1.710		1.471		104.0	116.7	
(s,a)	RHF	1.645	1.663	1.466	1.476	115.4	128.3	118.6
	MP2	1.743	1.757	1.474	1.491	114.8	122.8	114.3
	B3PW91	1.696	1.722	1.467	1.482	114.8	125.4	115.0
Se(NSiMe₃)₂								
(s,s)	RHF	1.638		1.730		121.5	153.5	
	MP2	1.711		1.766		127.3	143.3	
	B3PW91	1.676		1.744		124.3	148.8	
(a,a)	RHF	1.636		1.741		110.7	128.3	
	MP2	1.715		1.715		108.4	118.3	
	B3PW91	1.685		1.757		108.3	122.5	
(s,a)	RHF	1.633	1.646	1.757	1.745	114.8	131.6	126.4
	MP2	1.708	1.732	1.791	1.783	114.5	122.9	117.0
	B3PW91	1.678	1.699	1.771	1.763	114.5	128.0	119.9
Se(NC₁₀H₁₅)₂								
(s,a)	Exptl. ^a	1.679(8)	1.732(7)	1.496(11)	1.463(11)	113.0(3)	125.0(6)	117.6(5)

^a Ref. 33.

Table 3. Relative Energies (kJ mol⁻¹) of Sulfur and Selenium Diimide Conformations at CCSD/cc-pVDZ and CCSD(T)/cc-pVDZ Levels of Theory Utilizing Geometries Optimized at Different Levels of Theory.^a

Molecule	CCSD/ cc-pVDZ// RHF/ 6-31G*	CCSD/cc- pVDZ// MP2/ 6-31G*	CCSD/cc- pVDZ// B3PW91/ 6-31G*	CCSD(T)/cc- -pVDZ// B3PW91/ 6-31G*	Molecule	CCSD/cc- pVDZ// RHF/ 6-31G*	CCSD/cc- pVDZ// MP2/ 6-31G*	CCSD/cc- pVDZ// B3PW91/ 6-31G*	CCSD(T)/cc- -pVDZ// B3PW91/ 6-31G*
S(NH)₂					Se(NH)₂				
(s,s)	0.0	0.0	0.0	0.0	(s,s)	0.0	0.0	0.0	0.0
(a,a)	26.9	23.3	23.5	23.3	(a,a)	25.8	23.5	24.0	23.0
(s,a)	4.9	3.9	4.2	4.7	(s,a)	6.6	5.9	6.5	6.9
S(NMe)₂					Se(NMe)₂				
(s,s)	12.0	12.4	12.9	12.2	(s,s)	7.3	8.8	9.6	8.3
(a,a)	35.3	33.7	33.5	34.2	(a,a)	33.1	31.4	31.7	31.8
(s,a)	0.0	0.0	0.0	0.0	(s,a)	0.0	0.0	0.0	0.0
S(N^tBu)₂					Se(N^tBu)₂				
(s,s)	51.9	49.2	48.7	46.8	(s,s)	35.4	38.1	33.7	32.2
(a,a)	31.7	30.5	30.5	33.4	(a,a)	30.8	30.2	30.1	29.9
(s,a)	0.0	0.0	0.0	0.0	(s,a)	0.0	0.0	0.0	0.0
S(NSiMe₃)₂					Se(NSiMe₃)₂				
(s,s)	27.2	26.1	26.0	25.4	(s,s)	25.5	22.2	23.7	23.4
(a,a)	26.0	24.7	23.7	25.2	(a,a)	23.8	23.0	24.1	25.2
(s,a)	0.0	0.0	0.0	0.0	(s,a)	0.0	0.0	0.0	0.0

^a Unscaled zero-point energy corrections for the relative energies has been made at B3PW91/6-31G* level of theory.

Table 4. The Dependence of the B3PW91 Relative Energies on the Basis Sets. The Geometry Optimization and the ZPE Correction Have Been Carried Out at B3PW91/6-31G* Level of Theory.

Molecule	6-31G*	6-311G**	cc-pVDZ	cc-pVTZ	Molecule	6-31G*	6-311G**	cc-pVDZ	cc-pVTZ
S(NH)₂					Se(NH)₂				
(s,s)	0.0	0.0	0.0	0.0	(s,s)	0.0	0.0	0.0	0.0
(a,a)	25.0	25.9	22.3	14.9	(a,a)	20.7	21.7	23.6	15.7
(s,a)	4.7	5.2	4.1	1.5	(s,a)	4.0	5.6	6.8	3.4
S(NMe)₂					Se(NMe)₂				
(s,s)	5.8	4.7	10.9	9.4	(s,s)	8.5	6.1	6.9	8.5
(a,a)	30.5	34.5	30.3	25.6	(a,a)	27.7	29.1	28.9	24.7
(s,a)	0.0	0.0	0.0	0.0	(s,a)	0.0	0.0	0.0	0.0
S(N^tBu)₂					Se(N^tBu)₂				
(s,s)	34.4	35.0	41.9	36.9	(s,s)	40.2	23.1	24.9	24.8
(a,a)	24.4	27.6	24.2	19.4	(a,a)	24.1	27.6	21.4	16.8
(s,a)	0.0	0.0	0.0	0.0	(s,a)	0.0	0.0	0.0	0.0
S(NSiMe₃)₂					Se(NSiMe₃)₂				
(s,s)	14.2	12.9	26.1	18.1	(s,s)	23.3	13.3	20.7	17.0
(a,a)	17.2	20.8	19.3	11.1	(a,a)	15.5	15.1	16.7	10.6
(s,a)	0.0	0.0	0.0	0.0	(s,a)	0.0	0.0	0.0	0.0

Figure Captions

Figure 1. The *syn,syn*, *syn,anti*, and *anti,anti* conformations of chalcogen diimides E(NR)₂.

Figure 2. The dependence of NEN and ENA bond angles (E = S, Se; A = H, C, Si) on the organic substituent R in (a) S(NR)₂ and (b) Se(NR)₂.

Figure 3. The RHF/6-31G*, MP2/6-31G*, and B3PW91/6-31G* relative energies of the different conformations of (a) S(NR)₂ and (b) Se(NR)₂.

Figure 4. The CCSD(T)/cc-pVDZ//B3PW91/6-31G* relative energies of the different conformations of S(NR)₂ and Se(NR)₂.

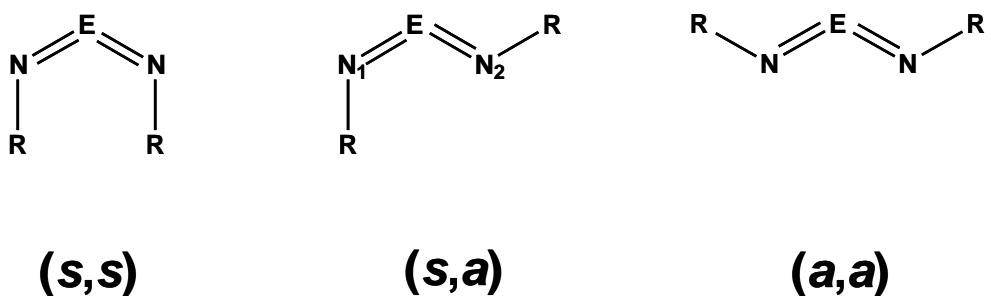


Figure 1

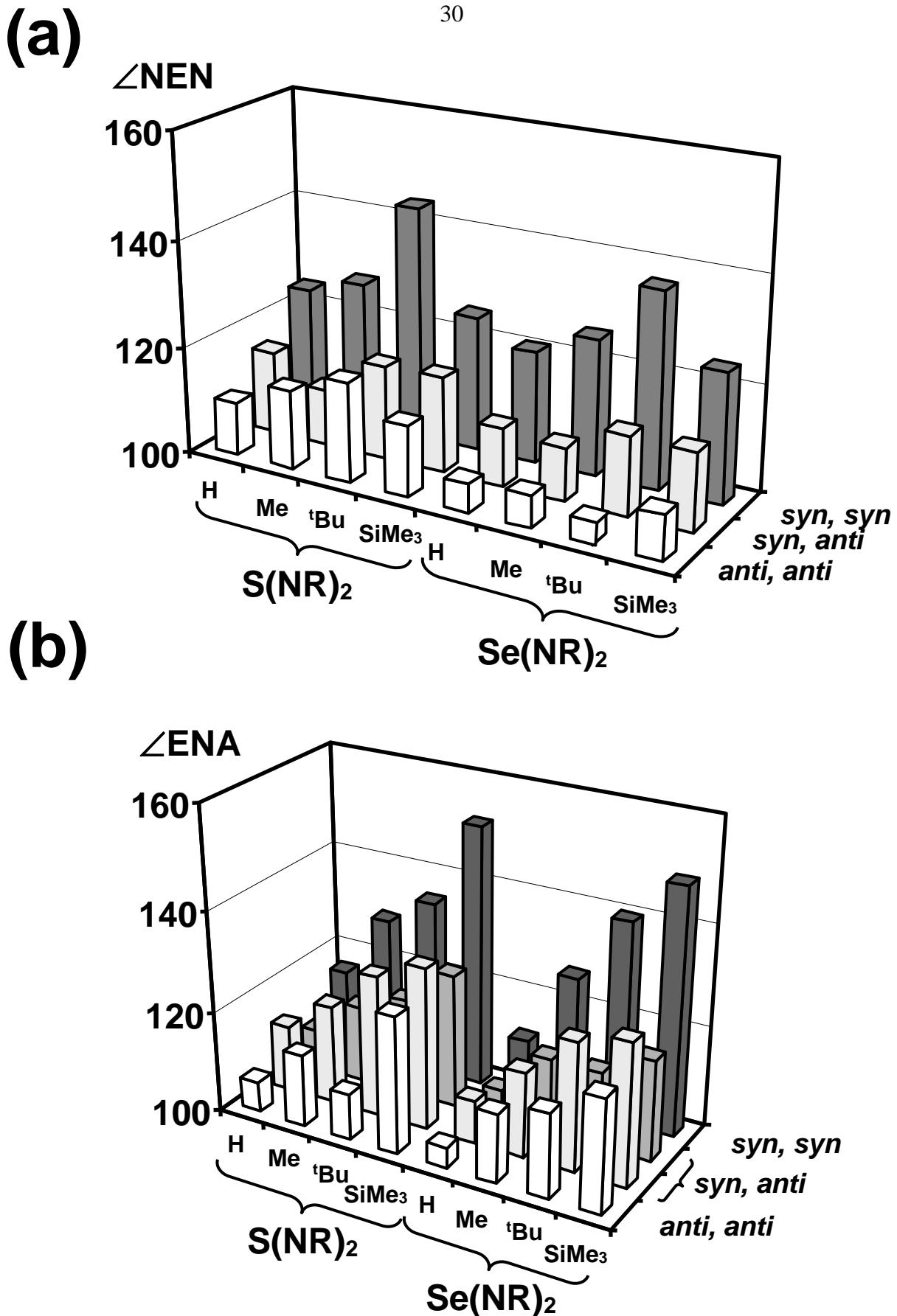


Figure 2

(a)

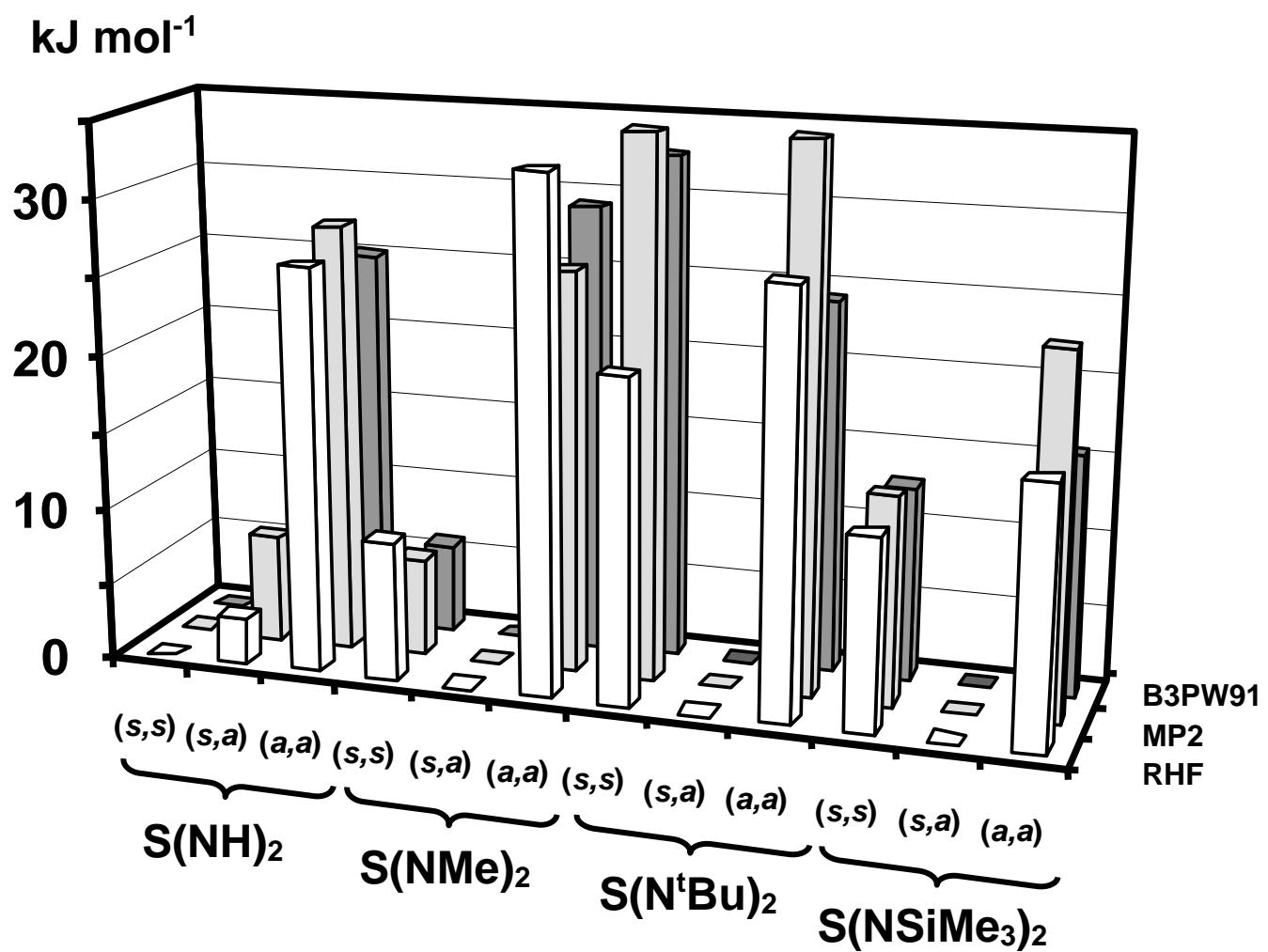


Figure 3(a)

(b)

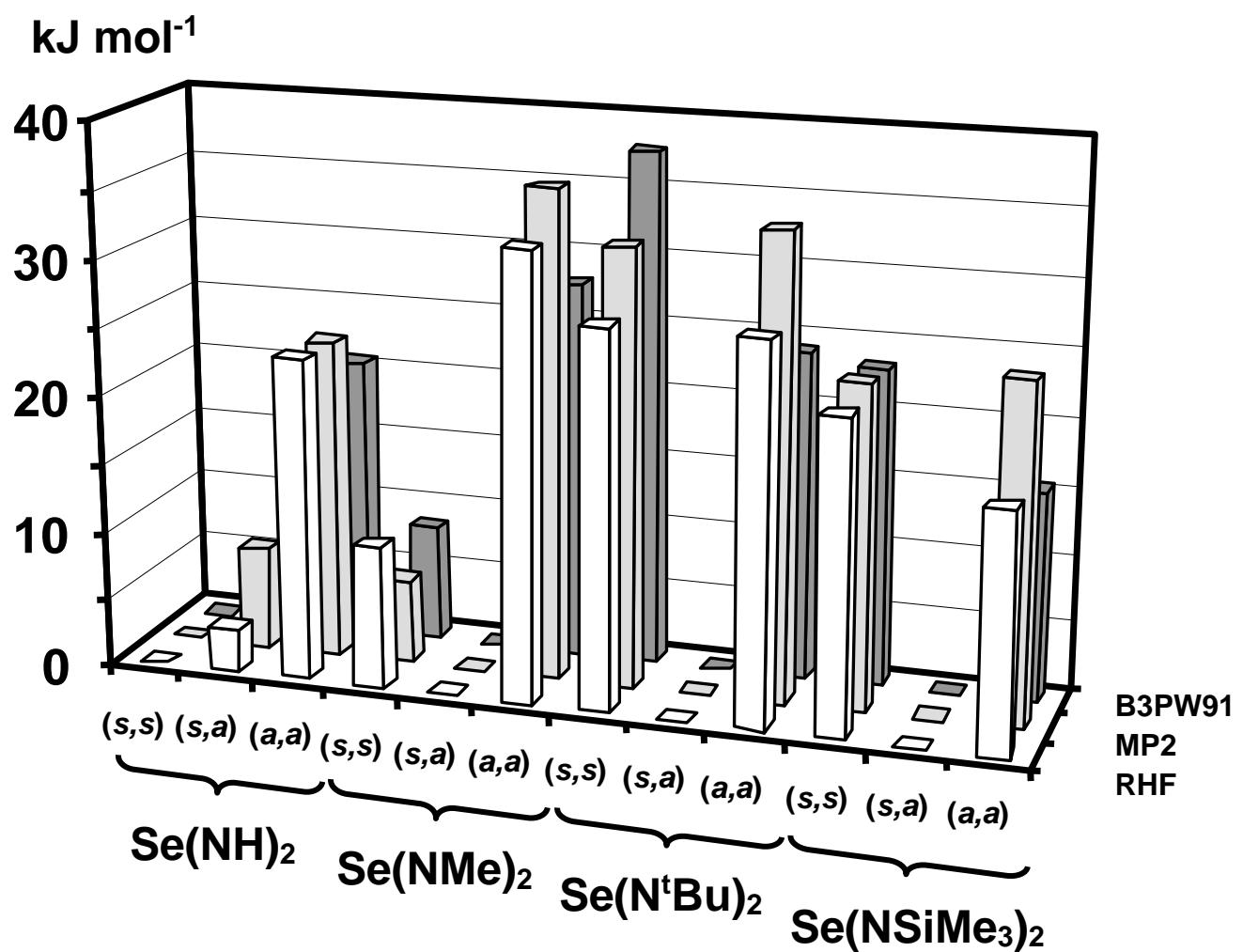


Figure 3(b)

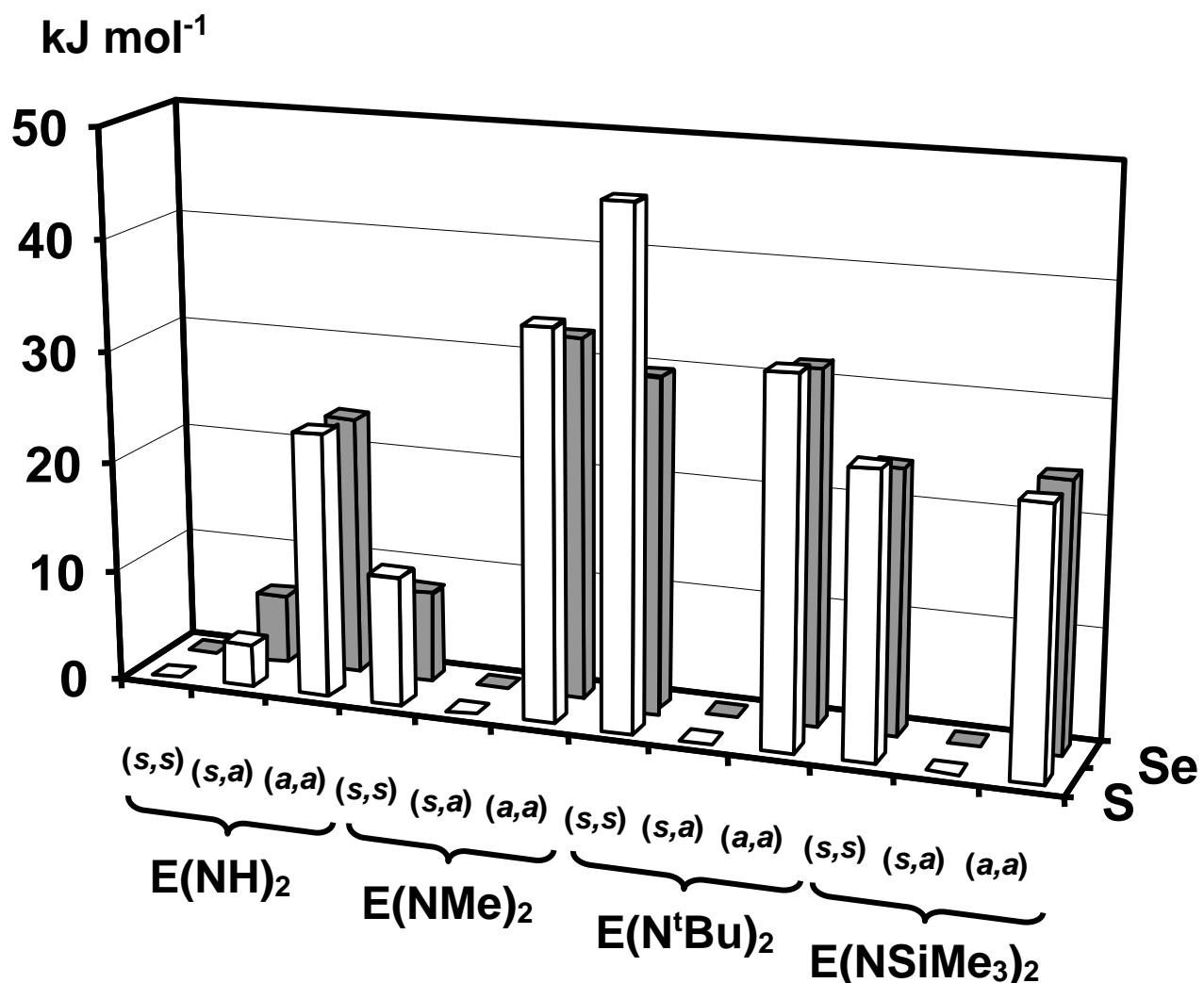


Figure 4