

This is a self-archived version of an original article. This version may differ from the original in pagination and typographic details.

Author(s): Jayapaul, Jabadurai; Komulainen, Sanna; Zhivonitko, Vladimir V.; Mareš, Jiří; Giri, Chandan; Rissanen, Kari; Lantto, Perttu; Telkki, Ville-Veikko; Schröder, Leif

Title: Hyper-CEST NMR of metal organic polyhedral cages reveals hidden diastereomers with diverse guest exchange kinetics

Year: 2022

Version: Published version

Copyright: © The Author(s) 2022










Rights: CC BY 4.0

Rights url: <https://creativecommons.org/licenses/by/4.0/>

Please cite the original version:

Jayapaul, J., Komulainen, S., Zhivonitko, V. V., Mareš, J., Giri, C., Rissanen, K., Lantto, P., Telkki, V.-V., & Schröder, L. (2022). Hyper-CEST NMR of metal organic polyhedral cages reveals hidden diastereomers with diverse guest exchange kinetics. *Nature Communications*, 13, Article 1708. <https://doi.org/10.1038/s41467-022-29249-w>

Hyper-CEST NMR of metal organic polyhedral cages reveals hidden diastereomers with diverse guest exchange kinetics

Jabadurai Jayapaul ^{1,2}, Sanna Komulainen ³, Vladimir V. Zhivonitko ³, Jiří Mareš ^{3,4}, Chandan Giri ⁵, Kari Rissanen ⁵, Perttu Lantto ^{3✉}, Ville-Veikko Telkki ³ & Leif Schröder ^{1,2✉}

Guest capture and release are important properties of self-assembling nanostructures. Over time, a significant fraction of guests might engage in short-lived states with different symmetry and stereoselectivity and transit frequently between multiple environments, thereby escaping common spectroscopy techniques. Here, we investigate the cavity of an iron-based metal organic polyhedron (Fe-MOP) using spin-hyperpolarized ^{129}Xe Chemical Exchange Saturation Transfer (hyper-CEST) NMR. We report strong signals unknown from previous studies that persist under different perturbations. On-the-fly delivery of hyperpolarized gas yields CEST signatures that reflect different Xe exchange kinetics from multiple environments. Dilute pools with $\sim 10^4$ -fold lower spin numbers than reported for directly detected hyperpolarized nuclei are readily detected due to efficient guest turnover. The system is further probed by instantaneous and medium timescale perturbations. Computational modeling indicates that these signals originate likely from Xe bound to three Fe-MOP diastereomers (T , C_3 , S_4). The symmetry thus induces steric effects with aperture size changes that tunes selective spin manipulation as it is employed in CEST MRI agents and, potentially, impacts other processes occurring on the millisecond time scale.

¹Molecular Imaging, Department of Structural Biology, Leibniz-Forschungsinstitut für Molekulare Pharmakologie (FMP), 13125 Berlin, Germany. ²Division of Translational Molecular Imaging, Deutsches Krebsforschungszentrum (DKFZ), 69120 Heidelberg, Germany. ³NMR Research Unit, University of Oulu, 90014 Oulu, Finland. ⁴Research Unit of Medical Imaging, Physics and Technology (MIPT), University of Oulu, 90014 Oulu, Finland. ⁵University of Jyväskylä, Department of Chemistry, 40014 Jyväskylä, Finland. ✉email: perttu.lantto@oulu.fi; leif.schroeder@dkfz.de

Self-assembled metal organic cages are supramolecular nanostructures that are widely used as hosts for various guests with applications as synthetic ion channels¹, sensors for organic analytes^{2,3}, in molecular recognition^{4,5}, drug delivery^{6,7}, catalysis⁸, gas adsorption and separation⁹, and magnetic resonance imaging (MRI)^{10,11}. Besides the three spatial dimensions, time is an important fourth dimension for the overall parameterization of cavity shape and guest exchange kinetics in cages and metal organic frameworks (MOFs)^{12,13}. Tailored host cavities have a crucial impact on the capacity and selectivity for guests^{14,15}. The time dimension is directly linked to guest turnover and the control of, e.g., magnetic properties of guests or catalysis mediated by supramolecular chemistry¹⁶. The involved stereochemistry is an important parameter as it impacts stereoselective transformations, chiral recognition of reversibly bound guests and steric fine tuning of the host portals. Metal organic polyhedra (MOPs) are one class of such nanostructures and they generally allow formation of hosts with broken symmetry, albeit some conformers can be highly favored over others. A mixture of different isomers is usually obtained upon constructing MOPs from asymmetric ligands. The application potential can be greatly influenced by the type of isomer and thus knowledge about the distribution and the availability of certain isomers is an important aspect. As an example, certain isomers from a mixture can exhibit a significantly lower IC₅₀ value for treatment of different cancer cells than others¹⁷. Additionally, resolved isomers of enantiopure MOPs are applied in asymmetric catalysis and chiral recognition, respectively^{18,19}. Also, a chiral amorphous MOP applied as the stationary phase in gas chromatography (GC) was recently shown to separate different enantiomers and various types of organic compounds²⁰. In general, NMR spectroscopy (¹H and ¹⁵N NMR²¹) is suitable to reveal symmetry differences but requires relatively high concentrations and is unsuitable for dilute, short-lived species. NMR signals can be enhanced by hyperpolarization (e.g., via dynamic nuclear polarization, DNP) in solution state NMR²². However, such DNP at room temperature still requires mandatory spin-labeling of the analyte during synthesis and 10¹⁷ spins in 10⁻⁷ L sample volume. Even this enhancement is insufficient for highly dilute states, particularly when somewhat rapid exchange is involved.

Due to such limitations, certain conformations might have been overlooked in MOPs which can exist as tetrahedral cages²³⁻³⁶ constructed from octahedral complex-forming metal ions. These hosts can temporarily bind different hydrophobic guests³⁷⁻⁴⁰. Some MOP ligands generate considerable fractions of heterochiral (C₃) or achiral (S₄) cages⁴¹ that are guest-responsive and exhibit dynamic interconversion between the diastereomers²¹. However, many initially synthesized M₄L₆ cages (M and L refer to metal and ligand) have been reported to exhibit only homochiral (T) species in solution. An example is the water-soluble tetrahedral [Fe₄L₆]⁴⁻ cage formed by treating 4,4'-diaminobiphenyl-2,2'-disulphonic acid and 2-formylpyridine in water together with Fe(II) and a base⁴². Experimental evidence (e.g., ¹H and ¹⁵N NMR) is available exclusively for its T species (Fig. 1a). Recent studies on a series of related Co-containing MOPs^{43,44} also found no signs for any other versions.

Conformations with broken symmetry but sparse population could be undetectable albeit such conformers in combination with a certain exchange regime would indeed cause a considerable amount of guests to experience a certain type of host cavity over time. Such inclusion complexes might, however, survive long enough to have an impact on how a guest probes different symmetries and thus also allow for stereoselective recognition of reactants that are processed in tailored cages. Apart from guest recognition, the overall host symmetry is linked to the spatial arrangement of the ligands and enables fine tuning of the portals.

This can have great impact on exchange kinetics. A related aspect is the field of MOFs where the presence of metastable intermediate states in such dynamic structures is in certain cases only observable after loading such structures with gaseous guests or by controlling temperature and pressure^{12,45}. The flexibility limits of most reported self-assembling systems is underexplored¹³ and understanding the consequences of fundamental properties like broken symmetry that may impact guest exchange mechanisms in flexible frameworks is key to the successful development of applications. Sensitive tools to monitor all cavity conformations and their guest uptake and release are therefore of high importance for understanding and designing these intriguing systems.

We thus explored ultra-sensitive NMR saturation transfer with ¹²⁹Xe to search for additional conformations of a yet exclusively assumed homochiral host. Xe as a monoatomic, spherical guest serves as a neutral “spy” to explore multiple symmetry environments. Hyperpolarized (HP) ¹²⁹Xe can be easily dispersed into solution and saturated magnetization is readily replaced by on-the-fly delivery of fresh HP gas. The experiment can be repeated within a few seconds to allow for stepwise acquisition of well-resolved chemical shift information and/or signal averaging. Even though this type of spectroscopy unlike multi-dimensional NMR does not directly reveal structural information on the cavity seen by the guest atom, differences in the Xe exchange timescales become visible in the line shape of the saturation responses. Exchange kinetics are highly relevant in host-guest systems with non-covalent nature where Xe can serve as a model guest to explore cavity access. The noble gas is a powerful hyperpolarized NMR reporter because it can be more or less directly applied to bulk material and sense exchange kinetics in host sites without the need for cumbersome isotope labeling of the cavity-shaping atoms but rather provides a reversibly bound spin label itself to probe the existence of multiple conformations.

We investigated the water-soluble tetrahedral [Fe₄L₆]⁴⁻ cage by performing chemical exchange saturation transfer (CEST) NMR experiments using HP ¹²⁹Xe^{46,47} since Xe forms inclusion complexes with this Fe-MOP⁴⁸. Combining the sensitivity enhancement of HP nuclei⁴⁹ with CEST, i.e., hyper-CEST, yields a ca. 10⁵ - 10⁶ fold signal enhancement over conventional NMR⁵⁰⁻⁵⁴. Recently, the ¹²⁹Xe hyper-CEST technique was applied for demonstrating a significant enhancement of the NMR signal of dissolved gas using a water soluble MOF⁵⁵. In another recent study, the capability of ¹²⁹Xe hyper-CEST NMR was demonstrated for differentiating a complex mixture containing MOFs as nano-cages in which the latter displayed similar chemical compositions. This was possible by tuning the MOF structure and by encoding multiple false colors for observed Xe chemical shifts⁵⁶. Saturation transfer (ST) techniques are a powerful tool that find an increasing number of applications. The magnetization in the bound pool is saturated by utilizing a continuous wave rf pulse applied over a time that is significantly longer than the average residence time of transiently bound spins. During this saturation period, chemical exchange causes a signal loss in bulk pool by affecting hundreds to thousands of spins per exchange site, thereby leading to accumulation of destroyed magnetization in the detection (bulk) pool⁵¹. ST techniques improve NMR sensitivity, retain the spectral dimension⁵⁷, and can be “tuned” in sensitivity to different time scales of exchange processes. Furthermore, Xe significantly extends the detection range towards exchange rates of 10⁴ Hz due to its large chemical shift range⁵⁸. This enables detecting so far elusive host conformations of stereochemistry interest.

In this work, the Xe@Fe-MOP system is investigated by ¹²⁹Xe NMR under different perturbation conditions (e.g., temperature cycling, portal blocking, and guest competition). We observe multiple signals of HP ¹²⁹Xe interacting with Fe-MOP and

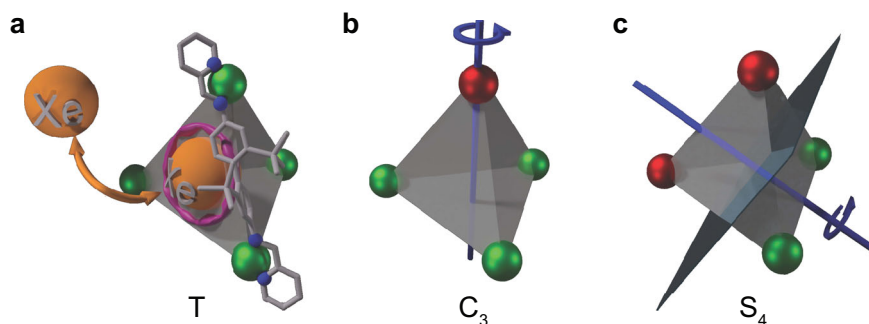


Fig. 1 Schematic illustration of Fe-MOP and its three diastereomers (atoms not to scale). Only one of the ligands oriented along the tetrahedral edge complexing two Fe^{2+} cations is shown for clarity purpose. Depending on how the participating nitrogen atoms (shown in blue) arrange around the metal centers (shiny spheres), the portal size and shape is slightly deformed (illustrated by the pink blurry ring). This impacts Xe exchange kinetics. Schematics of different Fe-MOP diastereomers and their symmetry axes/planes for the *T* (a), C_3 (b), and S_4 (c) diastereomers. a The *T* isomer has the ligands coordinated to homochiral metal centers (only green colored Fe^{2+} metal nodes) via the *anti*-linkages. b, c Conversely the ligands coordinate to the heterochiral metal centers (mix of red and green colored Fe^{2+} metal nodes) in C_3 and S_4 schemes via the *syn*-linkages.

validate normal cage integrity for guest access. The exchange pathways that Xe undergoes between different chemical environments of Fe-MOP and solution are checked further by inversion recovery (IR) and 2D CEST (short saturation) measurements. The multiple responses in the CEST spectra are interpreted using complementary theoretical methods and first principles modeling as computational methods are considered an integral part to understand such systems¹³.

Results and discussion

Thermal stability of supramolecular Fe-MOP. Tetrahedral Fe-MOPs ($[\text{Fe}_4\text{L}_6]^{4-}[(\text{CH}_3)_4\text{N}]_4^+$) were self-assembled under basic conditions at 50 °C (Supplementary Fig. 1 and Section 3)⁴². The product was structurally characterized using $^1\text{H-NMR}$ and ESI-MS (Supplementary Figs. 2 and 3). Fe-MOP (100 μM) encapsulation of Xe was investigated by Xe hyper-CEST NMR at temperatures ranging between $T = 25 \dots 50$ °C (see Supplementary Section 3.2, for additional tests between 4 and 20 °C). Beside the direct saturation signal of free Xe in water at 0 ppm, the CEST spectra displayed three responses at 13, 24, and 30 ppm, respectively, at 25 °C (Fig. 2; it is common practice in CEST experiments that the chemical shifts are referenced with respect to the signal that is observed, i.e., in this case, the signal of free Xe in solution). This is in contrast to only one peak corresponding to the 13 ppm signal of bound Xe observed using direct ^{129}Xe NMR detection⁴⁸. Also, the Co-containing versions $[\text{Co}_n\text{Fe}_{(4-n)}\text{L}_6]^{(4-)}$ ($n = 1-3$) have been reported to display only one signal each for encapsulated Xe⁴⁴. With increasing temperature, the intensities of the narrow signals at 13 and 24 ppm increased while the broad one at 30 ppm disappeared into the baseline upon reaching 35 °C. When reducing T from 50 °C back to 25 °C, the data matched the initial result (Fig. 2 and Supplementary Fig. 4 together with Supplementary Video 1). We thus conclude that the 30 ppm signal represents already relatively fast exchange at room temperature (RT), which yields inefficient CEST conditions at higher temperatures. The two stronger signals, however, represent intermediate exchange at RT, which yields even better saturation transfer at higher exchange rates. Overall, the temperature cycling revealed the structural robustness of the self-assembled Fe-MOP via reproducible hyper-CEST signatures.

Examining the origin of multiple CEST peaks. Next, the reason for observing multiple CEST peaks was checked. No indication was found that the oxidation state of the Fe(II) centers changed while the iron was participating in coordination with the ligands

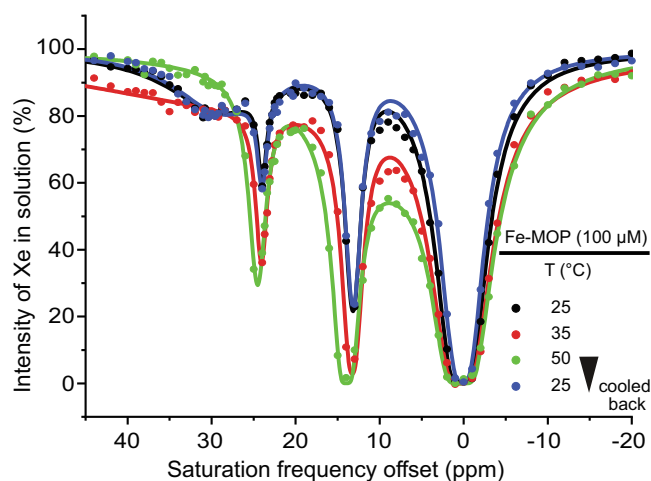


Fig. 2 Thermal stability of Fe-MOP studied at variable temperature using ^{129}Xe -hyper-CEST. Fe-MOP (100 μM) at 25 °C showed 3 responses at 13, 24, and 30 ppm. Stepwise increase in temperature causes two signals (13, 24 ppm) to increase while the third peak (30 ppm) disappeared at $T \geq 35$ °C. Cooling the Fe-MOP solution back to 25 °C restored the initial result.

(see Supplementary Fig. 6). Likewise, the presence of residual amounts of acetone used during Fe-MOP recrystallization cannot be the origin of additional CEST signals; in fact, large excess of $(\text{CH}_3)_2\text{CO}$ causes reduced intensity, see Supplementary Fig. 6. We also checked if storage of Fe-MOP in solution might have led to deterioration. However, CEST spectra of freshly mixed Fe^{2+} cations and ligands revealed instantaneously the previously observed signals. The response at 0 ppm appears broadened and a somewhat matching observation applies to the signal at 30 ppm (Supplementary Fig. 7). We interpret this as a presumably less efficient initial locking of the gas reflected by accelerated exchange until a chemical equilibrium is reached.

Another option for explaining multiple CEST signals is the contribution of different optically resolved (D ($\Delta\Delta\Delta\Delta$) and L ($\Lambda\Lambda\Lambda\Lambda$)) Fe-MOP isomers. The protocol for resolving the racemic (*rac*) mixture¹⁸ is described in Section 6 of the Supplementary information. The CEST spectra indicate that both D- and L-isomers contribute with the same type of hyper-CEST responses (Supplementary Fig. 9).

Next, we altered the access of Xe to the cages by a blocking strategy acting outside the cavity. The access through the four

faces of Fe-MOP was impeded using excess guanidinium hydrochloride (Gu · HCl) which is known to alter the incoming guest kinetics for cyclohexane⁵⁹. Here, Fe-MOP (100 μ M) was gradually blocked by increasing the Gu · HCl concentration. The CEST spectra remained unaffected for [Gu · HCl] < 1 mM. At 1 mM Gu · HCl, a slight left shift of CEST peaks was observed (somewhat anticipated due to higher salt concentration). A 10 mM concentration of Gu · HCl finally resulted in more left shifted peaks while presumably slowing down the Xe exchange for the signal at 30 ppm, which becomes better resolved (Fig. 3). Xe is supposedly small enough to retain sufficient cavity access. Even high concentrations of Gu · HCl can only decelerate the exchange, primarily affecting signals that reflect already enhanced exchange under normal conditions.

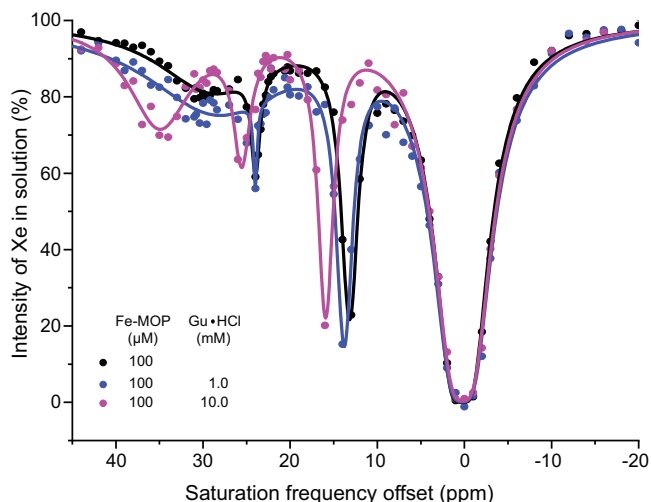


Fig. 3 Access of Xe to Fe-MOP cages under variable blocking of the faces by Gu · HCl. The CEST peaks of Fe-MOP (100 μ M) shifted slightly upon treatment with 1 mM Gu · HCl. 10 mM Gu · HCl resulted in more spectral shift due to higher salt concentration and slowed down the Xe exchange kinetics for the 30 ppm signal.

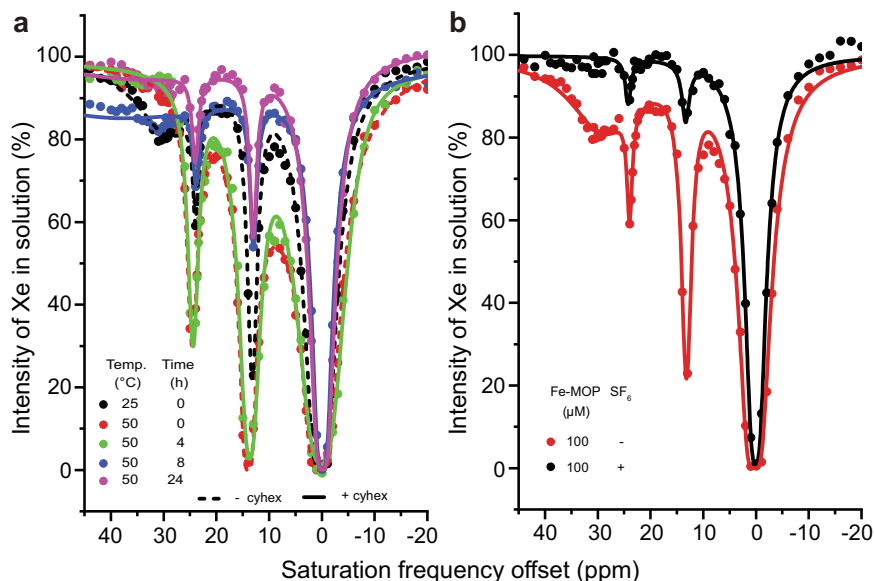


Fig. 4 Challenging the capture of Xe in Fe-MOP by competing larger guests. **a** The reference spectrum of Fe-MOP (100 μ M) plus cyclohexane at 25 °C revealed 3 CEST peaks. At 50 °C, the CEST peak at 30 ppm vanished into the baseline and intensities of 2 CEST peaks (13, 24 ppm) were approaching the baseline over 24 h indicating more cyclohexane encapsulation than Xe. **b** The larger competing guest, SF₆, caused even stronger decrease in the observed CEST responses.

Challenging the Fe-MOP cavity with larger competing guests.

Assuming sufficient purity and integrity of the hosts, we challenged the interaction of Xe with Fe-MOP directly inside the cavity through competition by adding guests like cyclohexane and SF₆. Fe-MOP (100 μ M) was mixed with cyclohexane (150-fold excess) and a reference CEST spectrum for no cyclohexane encapsulation was acquired at 25 °C (Fig. 4a). The Fe-MOP cavity was then probed by Xe in the presence of cyclohexane at 50 °C over 24 h. The slowly progressing cyclohexane encapsulation was imminent after 8 h as it led to reduced CEST peak intensities. Subsequently, Xe was removed by degassing and the sample mixture containing cyclohexane was maintained at 50 °C for 24 h. Then, re-measuring the sample at 50 °C with Xe displayed ca. 56% and 45% reduction of the 13 and 24 ppm CEST peak intensities compared to Fe-MOP alone (Fig. 4a; the 30 ppm signal is anyway hardly identifiable at 50 °C). Similarly, the Fe-MOP cavity was challenged by encapsulating an even larger and stronger binding guest, SF₆⁶⁰. Measuring the SF₆ c Fe-MOP (100 μ M) at 25 °C after 3 days of introducing the SF₆ atmosphere at 50 °C showed ca. 62 % and 31 % reduction in the intensities of the 13 and 24 ppm CEST peaks compared to untreated Fe-MOP. The fast exchanging CEST peak at 30 ppm disappeared into the baseline after the SF₆ treatment even when measuring the sample at 25 °C (Fig. 4b). Altogether, Xe binding remains possible owing to its smaller size and incomplete competing guest encapsulation (e.g., ~75% for SF₆⁶⁰). Here, the CEST technique reveals otherwise missed residual encapsulation. These results illustrate that the newly discovered signals must originate from cavity environments and that not only Xe has access to these.

Exchange pathways between different CEST pools.

We next revisited direct HP Xe NMR spectroscopy at fairly high Fe-MOP concentration (10.3 mM) and could show a previously unknown second and third Xe@Fe-MOP peak at 24 and 30 ppm, respectively (Supplementary Fig. 11a, b). This enabled us to estimate the populations of the suspected additional species (Supplementary Fig. 11c). The relative populations of the 13, 0 ppm (solution pool), 24 and 30 ppm pools at 10.3 mM cage concentration estimated from the direct spectrum are 1, 0.32, 0.010, and 0.0017,

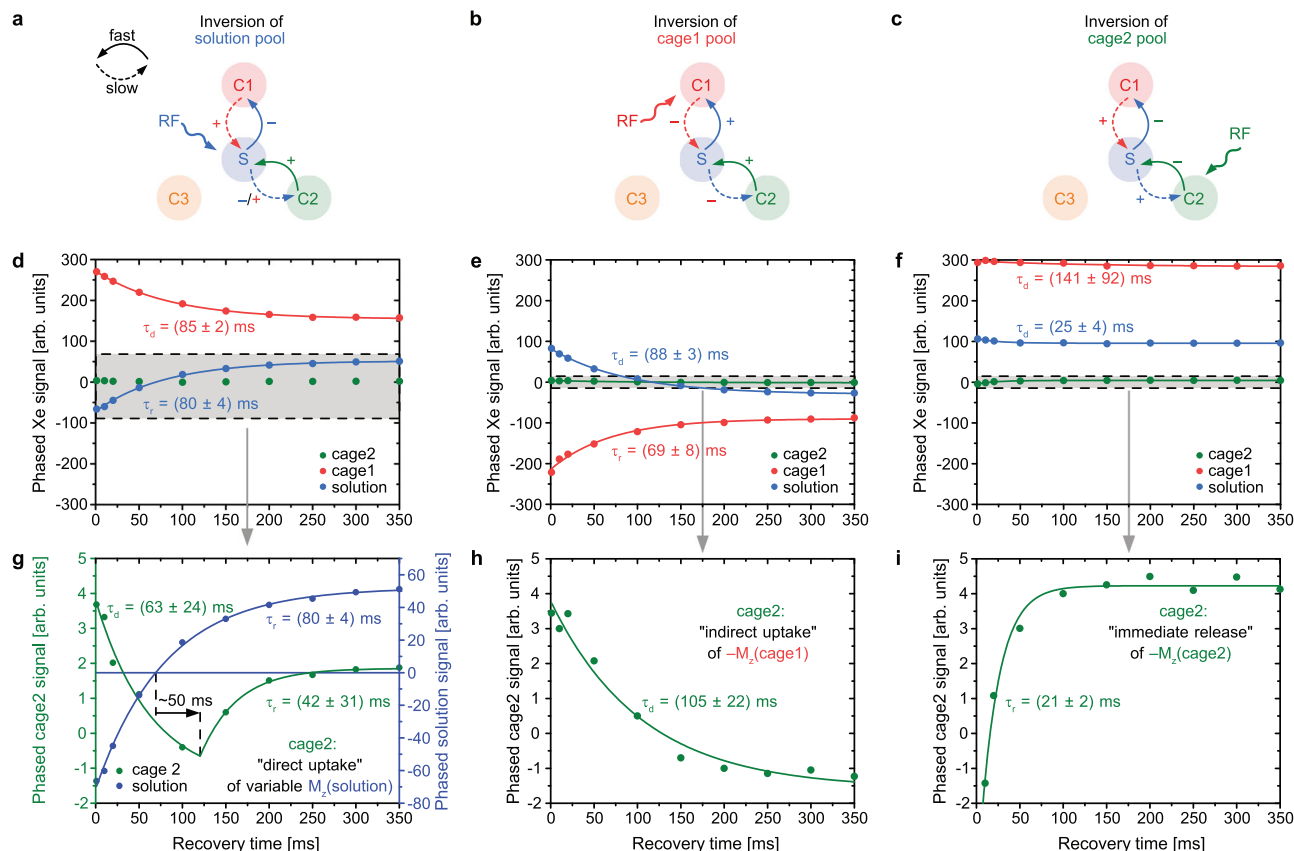


Fig. 5 Selective inversion recovery studies of Fe-MOP (10.3 mM) at 25 °C. **a–c** Pool network and detectable exchange routes. In each pairwise on/off exchange there is a fast rate constant from the smaller into the larger pool and a slow release constant vice versa. The directly inverted pool is indicated by the RF label. +/– signs indicate the net impact of enabling a signal recovery (growth) or a signal decay (as reaction to uptake of inverted magnetization). **d–f** Time evolution of the magnetizations. All can be fit to mono-exponential laws. The time constants τ_d (decay) and τ_r (recovery) give a qualitative impression for ordering direct (faster) and indirect (slower) responses. Equilibration between C1 and S happens on a medium time scale. The directly inverted pool always reacts fastest. **f** C1 reacts rather slow to C2 inversion while S sees the rather fast release from C2. **g–i** Zoom into the time curves for C2 illustrating three different regimes. **g** The perturbation from S propagates with medium speed into C2 because it simultaneously equilibrates with C1. C2 changes its behavior ca. 50 ms after the directly connected magnetization in S changes sign (piecewise fit for C2 with variable transition point). This reflects the direct uptake of the variable S magnetization. **h** Perturbation of C1 propagates only indirectly into C2, as reflected by a slow uptake of negative magnetization. **i** Directly inverted C2 signal is immediately released with rather fast exchange into S. Its uptake in S is partially compensated by positive magnetization from C1.

respectively (see Supplementary Fig. 11c). Despite their low population, these are acting as rather efficient CEST pools and are indicative for relative fast exchange rates. The strongest peak represents ca. 5.6 mM of occupied MOPs (see Supplementary Information), which is ca. 55 % of the overall host concentration.

In addition, we carried out inversion recovery (IR) experiments at conditions where the solution peak and the two most prominent cage signals can be observed, while one of these three signals was selectively inverted (Fig. 5 and Supplementary Fig. 12 together with Supplementary Video 2). The two bound Xe signals at 13 and 24 ppm were arbitrarily assigned as Cage1 and Cage2, respectively. Regarding this initial subset, the data supports a three site exchange model (Cage1 \leftrightarrow Solution \leftrightarrow Cage2), in which the two more intense cage signals were directly connected with the solution pool, but not directly connected with each other (Section 10.1, Supplementary information).

The IR measurements at high concentrations sense the immediate flux of magnetization after a selective, instantaneous perturbation of one pool. They showed that the system equilibrates within ca. 350 ms (Supplementary Movie 2). CEST measurements with detecting pools other than the solution pool provides complementary insights at increased sensitivity for intermediate timescale fluxes between the pools. This must be

detected before saturated Xe accumulates in the solution pool and then spreads throughout the network. We thus investigated CEST spectra at 25 and 10 °C with global readout (Supplementary Fig. 13). A short saturation ($t_{\text{sat}} = 250$ ms) was used for measuring the exchange with limited sensitivity enhancement. Saturating the solution peak at 25 °C yields a strong signal for Cage2 and a weak response for Cage1. This is because the small Cage2 pool is immediately saturated from the bulk pool whereas Cage1 represents a large enough pool that is not completely exchanged within 250 ms. The solution peak showed a (rather weak) response to Cage2 saturation and a decent response for Cage1 saturation. However, Cage2 saturation yields an over-proportional strong CEST response for the solution pool. This again reflects the rather efficient exchange from Cage2 into the solution pool. As Cage1 does practically not respond to Cage2 saturation, we conclude that the flux via the solution pool has not yet reached Cage1 within 250 ms. For the reverse direction, data suggests that despite the indirect flux of saturation from Cage1 via the solution pool into Cage2, this is indeed detectable after 250 ms and comparable to direct flux from the solution pool to Cage2 (Supplementary Fig. 14).

We then cooled the Fe-MOP solution to 10 °C for decelerated exchange from Cage3 signal and employed a higher saturation

strength ($B_1 = 5 \mu\text{T}$) for a length (t_{sat}) of 250 ms. Only the solution peak reacted to Cage3 saturation. This suggests that either it is very inefficient pathway or too fast flux is observed from Cage3 into Cage1 or Cage2 (Supplementary Fig. 14b, 15). Overall, the 2D CEST experiment enabled us to establish the possible CEST network between highly and sparsely populated CEST pools and the respective Fe-MOP states.

CEST spectra simulation and quantification. After excluding Fe-MOP structural degradation and hyper-CEST artifacts (CEST responses were observed without significant changes even after several months and were eventually also observed with thermally polarized Xe; see Supplementary Fig. 16) or structural imbalance induced by Xe bubbling (Supplementary Fig. 10) as explanation, we further analyzed the multiple CEST peaks. We hypothesize that they originate from Xe interaction with three different Fe-MOP diastereomers (T , C_3 , S_4). Literature strongly suggests that only T is observable in solution using conventional detection methods^{42,61}. Hyper-CEST, however, made it possible to detect such “hidden states” with low populations, although the ^1H signals of the minor diastereomers were covered by the stronger signals of the dominating T diastereomer (Supplementary Section 14). It is important to note that CEST extends the range that NMR can cover on the time scale of guest exchange kinetics and thus contributes to the need for methods that include time as an important fourth dimension for implementing novel applications with MOFs/MOPs¹³. Even though we did not vary the saturation conditions for the spectra, adjusting the saturation time and power makes CEST with xenon a versatile tool: it covers processes that are too fast for conventional NMR detection as well as those kinetics that accumulate only over 10s of seconds in the bulk pool as the HP ^{129}Xe NMR signal has a long life time to store information that is amplified through saturation transfer from a dilute pool. The observation that the CEST signals appear more or less instantaneously (Supplementary Fig. 7) indicates that the sub-populations of locked MOPs form in parallel and that the synthesis does not allow straightforward separation (or inter-conversion) of them for further isolated investigation. Experimentally obtained z-spectra can be described in the frame of the proposed four site exchange model (Supplementary Fig. 18) in combination with a simulation of the CEST experiment at 25 °C. Simulations were done by numerical integration of Bloch-McConnell type equations with chemical exchange terms and provided a z-spectrum which fitted excellent to the experimental data (Supplementary Fig. 17).

The simulation gave the following populations of sites for a Fe-MOP concentration of 100 μM : $P(\text{F}) = 96\% > P(\text{T}) = 3.8\% > P(\text{C}_3) = 0.14\% > P(\text{S}_4) = 0.041\%$. The 96% fraction of free Xe (F) corresponds to 369 μM dissolved Xe at 25 °C, thus, the readily detectable CEST signal at 30 ppm represents only 160 nM of Xe@ S_4 . In fact, the pool size of ^{129}Xe experiencing the S_4 cavity is then limited to 42 nM. This translates to only 3.8×10^{13} detectable spins for the S_4 pool in an 1.5 mL sample, thus demonstrating the enormous sensitivity of the applied technique (10^4 -fold less spins than in DNP NMR²²). The relative populations of the encapsulated sites are slightly different from those detected from direct spectra under different conditions where the Fe-MOP concentrations exceeds the free Xe concentration. The quantitative analysis showed significant differences in the Xe release rates, ranging from $4 \times 10^3 \text{ s}^{-1}$ for the sparsely populated S_4 conformer to 2.7 s^{-1} for the well known T -cage. These numbers agree with the qualitative changes in the spectra observed upon heating the sample. This “exchange tuning” aspect that the symmetry brings into the parameter space of MOPs extends the options for designing future hyper-CEST MRI

reporters that can also be selectively addressed through sufficient spectral dispersion.

Computational approach and findings. Quantum chemical calculations of ^{129}Xe NMR chemical shifts were performed (see *Computational methods* section) for Xe atoms placed inside of the three individual diastereomers. The otherwise unknown molecular structures of C_3 and S_4 diastereomers were derived from the X-ray structure⁴² of the T diastereomer. The chirality of one or two metal centers was manually changed to create C_3 or S_4 diastereomer, respectively. The conformational space of each isomer was sampled by generating high number of conformers (see *Methods* section below and Supplementary Section 13). The lowest free energy (G) values for conformations of the C_3 , and S_4 are $\Delta G = 0.6$ and 43.0 kJ/mol, respectively, with respect to the lowest energy conformation of T diastereomer. The thorough scrutiny of the full conformational space indicate that the lowest G conformations are the plausible structure candidates. This is supported further by large free energy differences to the next lowest conformations of each diastereomer ($\Delta G = +4 \dots +9$ kJ/mol) as well as high energy barriers between them.

Despite the density functional theory (DFT) computed ^{129}Xe chemical shifts for a few low energy structures of the conformational search are overestimated (see Section 13.3 in Supplementary material), there is an obvious trend of $\delta(\text{T}) < \delta(\text{C}_3) < \delta(\text{S}_4)$. The energetically most plausible lowest G structures also follow this trend and, therefore, their structures were studied further. DFT geometry optimization increases their volumes (118.4, 105.7, and 96.0 \AA^3 for T , C_3 , and S_4 , respectively; Xe van der Waals volume⁶²: 44 \AA^3). The cages seem to relax around the Xe atom during the optimization in liquid environment, thereby leading to a decrease in cavity volume compared to X-ray data.⁴² Larger cavities by DFT optimization results in ^{129}Xe chemical shifts within the experimental chemical shift range (see Table 5 in Supplementary material). The best estimates of $\delta^{129}\text{Xe}$ were obtained by statistical thermal averaging over Xe motion at room temperature. For the lowest G conformation of T diastereomer, the obtained $\delta(\text{T}) = 14.1$ ppm (solvent shift w.r.t gas shift is 188 ppm) matches well with the experimental value of 201 ppm. Values for other diastereomers, $\delta(\text{C}_3) = 66.7$ ppm and $\delta(\text{S}_4) = 114.5$ ppm, follow the above-mentioned chemical shift trend. Obviously, chemical shift differences are much larger than those obtained from the CEST spectra. However, currently neglected finite temperature effects due to motion of the cage atoms as well as solvent molecules may well decrease the chemical shift differences between the three diastereomers.

Therefore, the computational modeling suggests that the lowest CEST signal (13 ppm with respect to Xe in water) arises from the well-studied T diastereomer, while the 24 and 30 ppm signals may originate from C_3 and S_4 diastereomers, respectively.

For S_4 diastereomer, there are nine conformations with large variation in the volume while under similar energy criteria, our sampling approach exhibited only two and three conformations for T and C_3 cavities, respectively. This indicates that the structure of S_4 diastereomer is more flexible thus allowing larger openings for accommodation of different guests with varied sizes by its cavity. This finding potentially indicates easier and faster in-out exchange of Xe and may explain the specific features of the 30 ppm CEST signal: Firstly, the broader signal (see Fig. 1) is a demonstration of the faster exchange process. Secondly, the signal disappeared above 35 °C, as a consequence of the transition from intermediate to fast exchange regime. Thirdly, the chemical shift of the signal responds more to the temperature cycling than other signals (see Fig. 1 and Supplementary Fig. 4), and the CEST signal

becomes significantly narrow at low temperatures, denoting the transition between exchange regimes.

Theoretically, the additional CEST signals might arise also from cages loaded with several Xe atoms (see *Computational methods* section). However, loading the much larger solid-state X-ray cavity of the *T* diastereomer⁴² with two Xe atoms increased the calculated chemical shift of about 400 ppm compared to loading single Xe atom (accommodating three and four Xe atoms resulted in even larger chemical shifts). Therefore, we can exclude this explanation. Similarly, we were able to exclude the possibility that the additional signals could arise from simultaneous encapsulation of Xe with N₂ or He, which were present in the hyperpolarized gas mixture; computational estimates resulted in much larger chemical shift changes than the experimentally observed values and signatures of additional signals were observed even in the thermally polarized experiments without N₂ and He. Finally, we expect that, due to its hydrophobicity, the cage does not encapsulate water molecules together with Xe; loading of the *T* diastereomer with Xe and H₂O resulted also computed chemical shifts, which are much higher than the experimentally observed shifts.

In conclusion, we obtained experimental evidence through Xe NMR for two previously unknown bound Xe signals (24 and 30 ppm) for a self-assembling Fe-MOP. Xe even retained accessibility to the cavity upon impeding the portals by excess Gu · HCl, albeit with slower exchange. Competition with larger and stronger binders such as cyclohexane and SF₆ considerably reduces (~30–60 %) the transient binding of Xe but also demonstrates that other guests have access to the additionally identified cavity environments. The Xe exchange pathways between different CEST pools were established by IR and 2D CEST experiments and the exchange was found to occur presumably all via the solution pool. This indicates that Xe interaction with different structural forms (diastereomers) of Fe-MOP might be the reason for observing multiple, consistent and reproducible CEST peaks. The hypothesis was supported through extensive first principles modeling showing that structural differences of diastereomers explain both the order of observed Xe chemical shifts and the variety in exchange kinetics. This study demonstrates that the molecular symmetry is linked to different guest exchange regimes with release of bound Xe ranging between 10³ s⁻¹ and 10⁰ s⁻¹. The insights thus address two of the central questions in this field^{12,13}: (a) what are the guest exchange time scales occurring in supramolecular nanostructures and (b) what tools are available to observe them? These results illustrate the value of hyperpolarized guests that can be delivered repeatedly within a few seconds to reveal hidden host conformations in detailed solution spectra of bulk material without sophisticated sample preparation. MOP systems thus provide intriguing encapsulation, spin capture, and exchange properties that can be generated by performing selective chemistry either at the metal center (e.g., by metal substitution/exchange) or via the ligand structure (e.g., by post-synthetic ligand exchange, tethering different linker types), respectively. Future studies shall investigate the CEST imaging capabilities of different MOP symmetries and if the time scale that was sufficient for selective RF manipulation inside the formerly hidden conformations would also allow chemical reactions to happen with stereoselective features.

Methods

Materials used. 4,4'-diaminobiphenyl-2,2'-disulfonic acid hydrate (70 wt.% in H₂O) was purchased from Santa Cruz Biotechnology, Inc. Heidelberg, Germany. Iron(II) sulfate heptahydrate (≥ 99%), 2-formylpyridine (99%), tetramethylammonium hydroxide pentahydrate (≥ 97%), cyclohexane (≥ 99.8%), ((S)-BINOL) i.e., S-(−)-1,1'-Bi(2-naphthol), 99 % and Pluronic®L-81 were procured

from Sigma Aldrich Chemie GmbH, Steinheim, Germany. Guanidinium hydrochloride (Gu · HCl, ≥ 99.5%), was obtained from Carl Roth GmbH & Co.KG, Karlsruhe, Germany. Sulfur hexafluoride (SF₆) gas was kindly provided by Prof. Thomas Braun, Humboldt-Universität zu Berlin, Germany. MilliQ water (18.2 MΩ·cm) was used for all sample preparations. All chemicals and solvents were used as received without any further purification.

General methods. ¹H NMR data of Fe-MOP were recorded in D₂O. A potential acetone impurity encapsulated by Fe-MOP was removed by repeating the lyophilization cycles and complete removal of acetone was checked by ¹H NMR. The electron spray ionization mass spectrometry (ESI-MS) of Fe-MOP was recorded in 1:1 H₂O:MeOH mixture using Orbitrap Fusion (Thermo Scientific) with Xcalibur set up. The ESI-MS was performed by a direct injection (4 μL/min) of the sample and measuring in the negative mode using a Standard-Source Spray (3 kV) without using any gas. The circular dichroism (CD) measurements of Fe-MOP and its enantiomers in H₂O were carried out using Jasco J-715 spectrophotometer (Jasco, Maryland, U.S.A.).

Synthesis of supramolecular Fe₄L₆ (Fe-MOP). The supramolecular iron-based metal organic polyhedra (Fe₄L₆:Fe-MOP) was synthesized (Supplementary Fig. 1) as reported in an earlier publication⁴². Briefly, 4,4'-diaminobiphenyl-2,2'-disulfonic acid (2.03 mmol), 2-formylpyridine (4.06 mmol), tetramethylammonium hydroxide pentahydrate (4.06 mmol) and iron (II) sulfate heptahydrate (1.35 mmol) were dissolved in 25 ml degassed water. After mixing, the reaction mixture appeared as a dark purple solution. Reaction was carried out for 20 h at 50 °C under argon atmosphere. The final product was isolated as dark purple crystals by slow vapor diffusion of acetone into the aqueous solution. The structure was validated by ¹H NMR (Supplementary Fig. 2) and ESI-MS (Supplementary Fig. 3) in agreement to earlier reports⁴².

Xenon hyperpolarization and delivery. The NMR experiments were performed with a 9.4 T Bruker AV400 NMR spectrometer (Bruker Biospin, Ettlingen, Germany) equipped with gradient coils for MR imaging. The excitation and detection of the NMR signal was achieved using a 10 mm inner diameter double-resonance (¹²⁹Xe and ¹H) probe. A custom-designed continuous-flow polarizer was utilized to produce hyperpolarized (HP) Xe by spin-exchange optical pumping^{63,64}. A gas mixture of 2% Xe (26.4 % ¹²⁹Xe natural abundance), 10% N₂ and 88% He was used unless otherwise noted. The applied total pressure was 4.5 bar abs., thus the dissolved Xe concentration is ca. 369 μM under these conditions. The achieved ¹²⁹Xe spin polarization was approximately 25%. The gas mix was directly bubbled into 1.5 mL of the sample placed inside a 10 mm thick-walled NMR tube by using a spectrometer-triggered bubble dispenser (3.5 bar overpressure) via five fused-silica capillaries (Polymicro Technologies, Molex Incorporated, Caudebec les Elbeuf, France). The bubbling was controlled by using spectrometer-triggered gas-flow regulators (mass flow controllers, Omega Newport, Deckenpfronn, Germany).

Hyper-CEST spectroscopy. For each data point in the Chemical Exchange Saturation Transfer (CEST) spectrum, the gas mixture was directly bubbled for 12 s into the sample solution (flow rate 0.1 standard liter per minute (SLM)) followed by a 3 s delay to allow the bubbles to collapse. Hyper-CEST spectra were acquired using a continuous wave (CW)-saturation pulse (length $t_{\text{sat}} = 15$ s, amplitude $B_1 = 5.14$ μT). The (CW)-saturation pulses were applied with 1 ppm increments within 45 to −26 ppm spectral frequency offsets. The spectral acquisition was implemented by referencing the observed chemical shifts to the free Xe in solution peak at 0 ppm. Z-spectra were fitted to the exponential of a sum of Lorentzians using OriginPro 2018 (OriginLab, Northampton, MA). Unless otherwise indicated, the spectra of all samples were acquired at 25 °C.

Computational methods. Conformational searches for each diastereomers (*T*, *C*₃, *S*₄) were carried out for all three stereoisomers (see Supplementary material). First, high number of structures inaccessible from a local minimum by low energy motion were generated by rotating the benzylsulfonate units (5 angles for every unit resulting in 0.24 × 10⁹ rotamers). Then structures with van der Waals (vdW) clashes and other unrealistic conformers were filtered out and obtained plausible structures were minimally solvated with 36 water molecules - one for each oxygen atom in twelve SO₃⁻ groups. The overall charge on each conformer was compensated with four NH₄⁺ counterions on each face of the [Fe₄L₆]¹⁴⁻ complex and the Xe atom was placed inside its cavity. The structures were then optimized using the semiempirical extended tight binding GFN2-xTB method⁶⁵ via the XTB code (see Code availability statement). In addition to the explicit counterions and water solvent molecules, an implicit GBSA water solvent model was used in the geometry optimization. The aforementioned optimized structures were selected for further calculations only if Xe remained inside the cavity and SO₃⁻ groups were well exposed to the solvent. To obtain free energy estimates of maximum of twenty lowest energy conformers of each diastereomer, they were subjected to vibrational analysis calculation with the XTB code.

Density functional theory (DFT) calculations for the lowest free energy (*G*) conformers of each diastereomer (see more details in Supplementary material and

Supplementary Table 5) were carried out with the Turbomole code⁶⁶. Geometries of the three complexes including four NH_4^+ counterions were computed at the composite B97-3c method⁶⁷ with implicit COSMO solvation model of water⁶⁸.

At the obtained equilibrium geometries, ^{129}Xe NMR nuclear shielding calculations were carried out at scalar-relativistic local exact two-component (X2C) level⁶⁹. We used BHandHLYP^{70,71} DFT functional with D4 dispersion correction⁷² and x2c-TZVPall-s(Xe)/x2c-SVPall(other) basis sets together with finite nucleus model⁷³. BHandHLYP has been shown to provide best estimates for Xe chemical shifts in several studies of molecular cavities^{48,74–77} and molecules^{78–80}. It was observed in those studies that, while the inclusion of SR effects advances the non-relativistic (NR) calculations for Xe shifts, quite small and invariant spin-orbit (SO) relativistic contribution can be neglected. Therefore, SR method like X2C provides high quality results with moderate computational requirements.

While full molecular dynamics treatment of current extensive systems with counterions and large enough explicit solvation is currently not feasible, we focus on treatment of the most important effect, i.e., Xe thermal motion inside the rigid cage. We use the same approach as in previous studies^{48,75,76}, in which canonical Monte Carlo (MC-NVT) simulations with *in house* code were carried out on numerical potential energy surface (PES). PES and Xe NMR shift hyper-surface used for averaging were computed at the same level as the equilibrium geometry calculation above (see simulation details and thermal averages in Supplementary Table 5).

The loading effect on Xe chemical shift was studied by placing one to four Xe atoms inside the Xe-cage potential energy surface of $T\text{-Fe}_4\text{L}_6$ ⁴⁸. The thermal averaging over the dynamics of Xe atoms at $T = 300\text{ K}$ was obtained using the modified version of the *in house* MC-NVT code, where both $T\text{-Fe}_4\text{L}_6$ cage contribution to potential energy and NMR chemical shift of Xe were piecewise-linearly interpolated from the non-uniformly sampled 3D surfaces⁴⁸. For Xe-Xe interactions the relativistic potential energy⁶² and the non-relativistic effective binary chemical shift functions by Hanni et al.⁸¹ were used. Therefore, the Xe chemical shift for each Xe atom is a sum of the thermal averages of the Xe-Xe shift from the binary chemical shift and Xe- $T\text{-Fe}_4\text{L}_6$ shift (with respect to free Xe atom).

Data availability

All data generated and analysed during this study are included in this article and its Supplementary Information, and are also available from the authors upon request. The details about synthesis, structural characterization, hyper-CEST spectra under different perturbation conditions, inversion recovery measurements, 2D CEST (short saturation) experiment and computational methods are available as supplementary information.

Code availability

The code used for computational methods, including in-house MC-NVT python code, is also available from the authors upon request. The semiempirical extended tight binding program package xtb can be assessed using the following link <https://github.com/grimme-lab/xtb>. The xTB standalone code, Version 6.2. from the above link was utilized for the semiempirical calculations.

Received: 26 May 2021; Accepted: 3 March 2022;

Published online: 31 March 2022

References

- Hille, B. Ion channels of excitable membranes (Sinauer Associates, Inc., 2001).
- Gu, L. Q., Braha, O., Conlan, S., Cheley, S. & Bayley, H. Stochastic sensing of organic analytes by a pore-forming protein containing a molecular adapter. *Nature* **398**, 686–690 (1999).
- Howorka, S., Cheley, S. & Bayley, H. Sequence-specific detection of individual DNA strands using engineered nanopores. *Nat. Biotechnol.* **19**, 636–639 (2001).
- Kikuchi, T., Sato, S. & Fujita, M. Well-defined DNA nanoparticles templated by self-assembled $\text{M}_{12}\text{L}_{24}$ molecular spheres and binding of complementary oligonucleotides. *J. Am. Chem. Soc.* **132**, 15930–15932 (2010).
- Kamiya, N., Tominaga, M., Sato, S. & Fujita, M. Saccharide-coated $\text{M}_{12}\text{L}_{24}$ molecular spheres that form aggregates by multi-interaction with proteins. *J. Am. Chem. Soc.* **129**, 3816–3817 (2007).
- Hannon, M. J. Supramolecular DNA recognition. *Chem. Soc. Rev.* **36**, 280–295 (2007).
- Therrien, B. Drug delivery by water-soluble organometallic cages. *Top. Curr. Chem.* **319**, 35–55 (2012).
- Smulders, M. M. J. & Nitschke, J. R. Supramolecular control over Diels-Alder reactivity by encapsulation and competitive displacement. *Chem. Sci.* **3**, 785–788 (2012).
- Sudik, A. C. et al. Design, synthesis, structure, and gas (N_2 , Ar, CO_2 , CH_4 , and H_2) sorption properties of porous metal-organic tetrahedral and heterocuboidal polyhedra. *J. Am. Chem. Soc.* **127**, 7110–7118 (2005).
- He, C. et al. A hexanuclear gadolinium-organic octahedron as a sensitive MRI contrast agent for selectively imaging glucosamine in aqueous media. *Chem. Commun.* **48**, 9290–9292 (2012).
- Sokolow, G. E. et al. Metal-organic polyhedron with four Fe(III) centers producing enhanced T_1 magnetic resonance imaging contrast in tumors. *Inorg. Chem.* **61**, 2603–2611 (2022).
- Evans, J. D., Bon, V., Senkova, I., Lee, H.-C. & Kaskel, S. Four-dimensional metal-organic frameworks. *Nat. Commun.* **11**, 2690 (2020).
- Cerasale, D. J., Ward, D. C. & Easun, T. L. MOFs in the time domain. *Nat. Rev. Chem.* **6**, 9–30 (2022).
- Wang, Y. et al. Host-guest chemistry with water-soluble gold nanoparticle supraspheres. *Nat. Nanotechnol.* **12**, 170–176 (2017).
- Bai, S. et al. Permeable metal-organic framework gel membrane enables long-life cycling of rechargeable organic batteries. *Nat. Nanotechnol.* **16**, 77–84 (2021).
- Zhao, H. et al. Reversible trapping and reaction acceleration within dynamically self-assembling nanoflasks. *Nat. Nanotechnol.* **11**, 82–88 (2016).
- Samanta, S. K. & Isaacs, L. Biomedical applications of metal organic polygons and polyhedra (MOPs). *Coord. Chem. Rev.* **410**, 213181 (2020).
- Wan, S., Lin, L.-R., Zeng, L., Lin, Y. & Zhang, H. Efficient optical resolution of water-soluble self-assembled tetrahedral M_4L_6 cages with 1,1'-bi-2-naphthol. *Chem. Commun.* **50**, 15301–15304 (2014).
- Zhang, D., Ronson, T. K., Zou, Y.-Q. & Nitschke, J. R. Metal-organic cages for molecular separations. *Nat. Rev. Chem.* **5**, 168–182 (2021).
- Tang, B. et al. Chiral amorphous metal-organic polyhedra used as the stationary phase for high-resolution gas chromatography separations. *Chirality* **32**, 1178–1185 (2020).
- Clegg, J. K. et al. A stimuli responsive system of self-assembled anion-binding $[\text{Fe}_4\text{L}_6]^{18+}$ cages. *Chem. Sci.* **4**, 68–76 (2013).
- Liu, G. et al. One-thousand-fold enhancement of high field liquid nuclear magnetic resonance signals at room temperature. *Nat. Chem.* **9**, 676–680 (2017).
- Caulder, D. L., Powers, R. E., Parac, T. N. & Raymond, K. N. The self-assembly of a pre-designed tetrahedral M_4L_6 supramolecular cluster. *Angew. Chem. Int. Ed.* **37**, 1840–1843 (1998).
- Parac, T. N., Caulder, D. L. & Raymond, K. N. Selective encapsulation of aqueous cationic guests into a supramolecular tetrahedral $[\text{M}_4\text{L}_6]^{12-}$ anionic host. *J. Am. Chem. Soc.* **120**, 8003–8004 (1998).
- He, C. et al. Self-assembly of tetrahedral M_4L_6 clusters from a new rigid ligand. *J. Chem. Soc., Dalton Trans.* **2**, 134–135 (2002).
- Custelcean, R., Bosano, J., Bonnesen, P., Kertesz, V. & Hay, B. Computer-aided design of a sulfate-encapsulating receptor. *Angew. Chem. Int. Ed.* **48**, 4025–4029 (2009).
- Tidmarsh, I. S. et al. Further investigations into tetrahedral M_4L_6 cage complexes containing guest anions: new structures and NMR spectroscopic studies. *New J. Chem.* **33**, 366–375 (2009).
- Caulder, D. L. & Raymond, K. N. Supermolecules by design. *Acc. Chem. Res.* **32**, 975–982 (1999).
- Bolliger, J. L., Belenguer, A. M. & Nitschke, J. R. Enantiopure water-soluble $[\text{Fe}_4\text{L}_6]$ cages: host-guest chemistry and catalytic activity. *Angew. Chem. Int. Ed.* **52**, 7958–7962 (2013).
- Castilla, A. M., Ramsay, W. J. & Nitschke, J. R. Stereochemistry in subcomponent self-assembly. *Acc. Chem. Res.* **47**, 2063–2073 (2014).
- Jiménez, A. et al. Selective encapsulation and sequential release of guests within a self-sorting mixture of three tetrahedral cages. *Angew. Chem. Int. Ed.* **53**, 4556–4560 (2014).
- Giri, C., Sahoo, P. K., Puttreddy, R., Rissanen, K. & Mal, P. Solvent-free ball-milling subcomponent synthesis of metallosupramolecular complexes. *Chem. Eur. J.* **21**, 6390–6393 (2015).
- Zhang, D., Ronson, T. K. & Nitschke, J. R. Functional capsules via subcomponent self-assembly. *Acc. Chem. Res.* **51**, 2423–2436 (2018).
- Zhang, D. et al. Enantiopure $[\text{Cs}^+/\text{Xe} \text{ Cryptophane}] \text{C Fe}^{\text{II}}_4\text{L}_4$ hierarchical superstructures. *J. Am. Chem. Soc.* **141**, 8339–8345 (2019).
- Kieffer, M., Bilbeisi, R. A., Thoburn, J. D., Clegg, J. K. & Nitschke, J. R. Guest binding drives host redistribution in libraries of $\text{Co}^{\text{II}}_4\text{L}_4$ cages. *Angew. Chem. Int. Ed.* **59**, 11369–11373 (2020).
- Percástegui, E. G., Ronson, T. K. & Nitschke, J. R. Design and applications of water-soluble coordination cages. *Chem. Rev.* **120**, 13480–13544 (2020).
- Sawada, T. & Fujita, M. A single Watson-Crick G · C base pair in water: Aqueous hydrogen bonds in hydrophobic cavities. *J. Am. Chem. Soc.* **132**, 7194–7201 (2010).
- Hastings, C. J., Pluth, M. D., Biros, S. M., Bergman, R. G. & Raymond, K. N. Simultaneously bound guests and chiral recognition: a chiral self-assembled supramolecular host encapsulates hydrophobic guests. *Tetrahedron* **64**, 8362–8367 (2008).
- Biros, S. M., Bergman, R. G. & Raymond, K. N. The hydrophobic effect drives the recognition of hydrocarbons by an anionic metal-ligand cluster. *J. Am. Chem. Soc.* **129**, 12094–12095 (2007).
- Kishi, N., Li, Z., Yoza, K., Akita, M. & Yoshizawa, M. An M_2L_4 molecular capsule with an anthracene shell: Encapsulation of large guests up to 1 nm. *J. Am. Chem. Soc.* **133**, 11438–11441 (2011).

41. Meng, W., Clegg, J. K., Thoburn, J. D. & Nitschke, J. R. Controlling the transmission of stereochemical information through space in terphenyl-edged Fe_4L_6 cages. *J. Am. Chem. Soc.* **133**, 13652–13660 (2011).
42. Mal, P., Schultz, D., Beyeh, K., Rissanen, K. & Nitschke, J. An unlockable-relockable iron cage by subcomponent self-assembly. *Angew. Chem. Int. Ed.* **47**, 8297–8301 (2008).
43. Du, K., Zemerov, S. D., Hurtado Parra, S., Kikkawa, J. M. & Dmochowski, I. J. Paramagnetic organocobalt capsule revealing xenon host-guest chemistry. *Inorg. Chem.* **59**, 13831–13844 (2020).
44. Du, K., Zemerov, S. D., Carroll, P. J. & Dmochowski, I. J. Paramagnetic shifts and guest exchange kinetics in $\text{Co}_n\text{Fe}_{4-n}$ metal-organic capsules. *Inorg. Chem.* **59**, 12758–12767 (2020).
45. Katsoulidis, A. P. et al. Chemical control of structure and guest uptake by a conformationally mobile porous material. *Nature* **565**, 213–217 (2019).
46. Schröder, L. Xenon for NMR biosensing-inert but alert. *Physica Med.* **29**, 3–16 (2013).
47. Schröder, L. Xenon biosensor HyperCEST MRI. In *Hyperpolarized and Inert Gas MRI*, 263–277 (Elsevier, 2017).
48. Roukala, J. et al. Encapsulation of xenon by a self-assembled Fe_4L_6 metallosupramolecular cage. *J. Am. Chem. Soc.* **137**, 2464–2467 (2015).
49. Barskiy, D. A. et al. NMR hyperpolarization techniques of gases. *Chem. Eur. J.* **23**, 725–751 (2017).
50. Schröder, L., Lowery, T. J., Hilty, C., Wemmer, D. E. & Pines, A. Molecular imaging using a targeted magnetic resonance hyperpolarized biosensor. *Science* **314**, 446–449 (2006).
51. Jayapaul, J. & Schröder, L. Molecular sensing with host systems for hyperpolarized ^{129}Xe . *Molecules* **25**, 4627 (2020).
52. Kunth, M., Witte, C., Hennig, A. & Schröder, L. Identification, classification, and signal amplification capabilities of high-turnover gas binding hosts in ultra-sensitive NMR. *Chem. Sci.* **6**, 6069–6075 (2015).
53. Rose, H. M. et al. Development of an antibody-based, modular biosensor for ^{129}Xe NMR molecular imaging of cells at nanomolar concentrations. *Proc. Natl. Acad. Sci.* **111**, 11697–11702 (2014).
54. Kunth, M., Döpfert, J., Witte, C., Rossella, F. & Schröder, L. Optimized use of reversible binding for fast and selective nmr localization of caged xenon. *Angew. Chem. Int. Ed.* **51**, 8217–8220 (2012).
55. Zeng, Q. et al. Hyperpolarized Xe NMR signal advancement by metal-organic framework entrapment in aqueous solution. *Proc. Natl. Acad. Sci.* **117**, 17558–17563 (2020).
56. Yang, Y. et al. Coloring ultrasensitive MRI with tunable metal-organic frameworks. *Chem. Sci.* **12**, 4300–4308 (2021).
57. Goren, E., Avram, L. & Bar-Shir, A. Versatile non-luminescent color palette based on guest exchange dynamics in paramagnetic cavitands. *Nat. Commun.* **12**, 3072 (2021).
58. Kunth, M. & Schröder, L. Binding site exchange kinetics revealed through efficient spin-spin dephasing of hyperpolarized ^{129}Xe . *Chem. Sci.* **12**, 158–169 (2021).
59. Zarra, S., Smulders, M. M. J., Lefebvre, Q., Clegg, J. K. & Nitschke, J. R. Guanidinium binding modulates guest exchange within an M_4L_6 capsule. *Angew. Chem. Int. Ed.* **51**, 6882–6885 (2012).
60. Riddell, I. A., Smulders, M. M. J., Clegg, J. K. & Nitschke, J. R. Encapsulation, storage and controlled release of sulfur hexafluoride from a metal-organic capsule. *Chem. Commun.* **47**, 457–459 (2011).
61. Ronson, T. K., Zarra, S., Black, S. P. & Nitschke, J. R. Metal-organic container molecules through subcomponent self-assembly. *Chem. Commun.* **49**, 2476–2490 (2013).
62. Hanni, M., Lantto, P., Runeberg, N., Jokisaari, J. & Vaara, J. Calculation of binary magnetic properties and potential energy curve in xenon dimer: Second virial coefficient of ^{129}Xe nuclear shielding. *J. Chem. Phys.* **121**, 5908–5919 (2004).
63. Witte, C., Kunth, M., Döpfert, J., Rossella, F. & Schröder, L. Hyperpolarized xenon for NMR and MRI applications. *J. Vis. Exp.* **67**, e4268–e4268 (2012).
64. Alonso de Mezquia, D., Mounir Bou-Ali, M., Madariaga, J. A. & Santamaría, C. Mass effect on the Soret coefficient in n-alkane mixtures. *J. Chem. Phys.* **140**, 084503 (2014).
65. Bannwarth, C., Ehlert, S. & Grimme, S. GFN2-xTB—An accurate and broadly parametrized self-consistent tight-binding quantum chemical method with multipole electrostatics and density-dependent dispersion contributions. *J. Chem. Theory Comput.* **15**, 1652–1671 (2019).
66. TURBOMOLE V7.5.1 2021, a development of University of Karlsruhe and Forschungszentrum Karlsruhe GmbH, 1989–2007, TURBOMOLE GmbH, since 2007; available from <https://www.turbomole.org>.
67. Brandenburg, J. G., Bannwarth, C., Hansen, A. & Grimme, S. B97-3c: A revised low-cost variant of the B97-D density functional Method. *J. Chem. Phys.* **148**, 064104 (2018).
68. Klamt, A. & Schüürmann, G. COSMO: a new approach to dielectric screening in solvents with explicit expressions for the screening energy and its gradient. *J. Chem. Soc., Perkin Trans.* 2799–2805 (1993).
69. Franzke, Y. J. & Weigend, F. NMR shielding tensors and chemical shifts in scalar-relativistic local exact two-component theory. *J. Chem. Theory Comput.* **15**, 1028–1043 (2019).
70. Becke, A. D. A new mixing of Hartree-Fock and local density-functional theories. *J. Chem. Phys.* **98**, 1372–1377 (1993).
71. Lee, C., Yang, W. & Parr, R. G. Development of the Colle-Salvetti correlation-energy formula into a functional of the electron density. *Phys. Rev. B* **37**, 785–789 (1988).
72. Caldeweyher, E. et al. A generally applicable atomic-charge dependent London dispersion correction. *J. Chem. Phys.* **150**, 154122 (2019).
73. Franzke, Y. J., Treß, R., Pazdera, T. M. & Weigend, F. Error-consistent segmented contracted all-electron relativistic basis sets of double- and triple-zeta quality for NMR shielding constants. *Phys. Chem. Chem. Phys.* **21**, 16658–16664 (2019).
74. Straka, M., Lantto, P. & Vaara, J. Toward calculations of the ^{129}Xe chemical shift in $\text{Xe}@C_{60}$ at experimental conditions: Relativity, correlation, and dynamics. *J. Phys. Chem. A* **112**, 2658–2668 (2008).
75. Selent, M. et al. Clathrate structure determination by combining crystal structure prediction with computational and experimental ^{129}Xe NMR spectroscopy. *Chem. Eur. J.* **23**, 5258–5269 (2017).
76. Komulainen, S. et al. Inside information on xenon adsorption in porous organic cages by NMR. *Chem. Sci.* **8**, 5721–5727 (2017).
77. Peuravaara, P. et al. Chemical shift extremum of $^{129}\text{Xe}(\text{aq})$ reveals details of hydrophobic solvation. *Sci. Rep.* **8**, 7023 (2018).
78. Lantto, P. & Vaara, J. ^{129}Xe chemical shift by the perturbational relativistic method: Xenon fluorides. *J. Chem. Phys.* **127**, 084312 (2007).
79. Lantto, P., Standara, S., Riedel, S., Vaara, J. & Straka, M. Exploring new ^{129}Xe chemical shift ranges in HXeY compounds: Hydrogen more relativistic than xenon. *Phys. Chem. Chem. Phys.* **14**, 10944–10952 (2012).
80. Lantto, P., Kangasvieri, S. & Vaara, J. Rovibrational effects on NMR shieldings in a heavy-element system: XeF_2 . *J. Chem. Phys.* **137**, 214309 (2012).
81. Hanni, M., Lantto, P. & Vaara, J. Pairwise additivity in the nuclear magnetic resonance interactions of atomic xenon. *Phys. Chem. Chem. Phys.* **11**, 2485–2496 (2009).

Acknowledgements

We acknowledge support provided by the DFG through Koselleck grant no. 316693477 (SCHR 995/5-1) and by the Dieter Morszeck Stiftung to L.S., University of Jyväskylä, the Academy of Finland (projects 285666 (P.L.), 289649 (V.-V.T.), 294027 (V.-V.T.), 319216 (V.-V.T.), 340099 (V.-V.T.) and 323480 (V.V.Z.), University of Oulu (Kvantum Institute) and Horizon 2020 (H2020/2018-2022)/ERC grant agreement no. 772110 to V.-V.T. Computational resources due to CSC (Espoo, Finland) and the Finnish Grid and Cloud Infrastructure project (persistent identifier urn:nbn:fi:research-infra-2016072533), were used. The authors thank Barth van Rossum, FMP Berlin, for support with rendering of graphical illustrations.

Author contributions

J.J., L.S., and V.-V.T. designed and J.J. and L.S. executed different experiments presented in the manuscript. C.G., K.R., and J.J. synthesized the MOP cages. S.K., J.M., and P.L. carried out conformational search and first principles modeling of Xe NMR shielding. J.J., L.S. analyzed and fitted the hyper-CEST data. V.V.Z. performed the four site exchange CEST spectrum simulation and the IR data fitting using the three site exchange model. J.J. wrote the first version of the manuscript. All authors were involved in results discussion and editing of the manuscript.

Funding

Open Access funding enabled and organized by Projekt DEAL.

Competing interests

The authors declare no competing interests.

Additional information

Supplementary information The online version contains supplementary material available at <https://doi.org/10.1038/s41467-022-29249-w>.

Correspondence and requests for materials should be addressed to Perttu Lantto or Leif Schröder.

Peer review information *Nature Communications* thanks Russell Morris, Rochus Schmid, Qi-Qiang Wang and the other, anonymous, reviewer(s) for their contribution to the peer review of this work.

Reprints and permission information is available at <http://www.nature.com/reprints>

Publisher's note Springer Nature remains neutral with regard to jurisdictional claims in published maps and institutional affiliations.



Open Access This article is licensed under a Creative Commons Attribution 4.0 International License, which permits use, sharing, adaptation, distribution and reproduction in any medium or format, as long as you give appropriate credit to the original author(s) and the source, provide a link to the Creative Commons license, and indicate if changes were made. The images or other third party material in this article are included in the article's Creative Commons license, unless indicated otherwise in a credit line to the material. If material is not included in the article's Creative Commons license and your intended use is not permitted by statutory regulation or exceeds the permitted use, you will need to obtain permission directly from the copyright holder. To view a copy of this license, visit <http://creativecommons.org/licenses/by/4.0/>.

© The Author(s) 2022